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5129/21

October/November 2013

2 hours 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

A copy of the Periodic Table is printed on page 24.

Electronic calculators may be used.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

This document consists of **22** printed pages and **2** blank pages.

- 1 Use words from the list to complete the sentences below.

amino acids	bile	egestion	excretion
expired air	fat	glucose	kidneys
liver	lungs	respiration	urine

Each word may be used once, more than once or not at all.

Urea is formed from that the body no longer needs.

The formation of urea takes place in the

Urea is passed out of the body in

This is an example of

[4]

- 2 Fig. 2.1 gives information about some of the elements in Group I of the Periodic Table.

element	symbol	proton number
lithium	Li	3
sodium	Na	11
potassium	K	19
rubidium	Rb	37
caesium	Cs	55

Fig. 2.1

- (a) State the name given to the elements in Group I.

.....[1]

- (b) State the trend shown by the melting points as the proton number increases.

.....[1]

- (c) All the elements in Group I react with water to produce a metal hydroxide and hydrogen.

State the test for hydrogen.

test.....

result.....

.....[2]

- (d) Rubidium reacts with chlorine to produce rubidium chloride.

- (i) Construct an equation for this reaction.

.....[1]

- (ii) State the type of bonding present in rubidium chloride.

.....[1]

For
Examiner's
Use

3 A battery supplies energy to the electric motor of a toy car.

(a) When the car is moving at constant speed, the current in the motor is 0.80 A.

Calculate the charge flowing through the motor in 10 minutes and state the unit.

charge = unit [3]

(b) Complete the following sentence about energy changes.

Some of the electrical energy is converted into sound energy and thermal energy.

The rest is converted into energy. [1]

(c) The toy car moves 4.8 m in 1.5 s.

Calculate its average speed.

speed = m/s [2]

- 4 Fig. 4.1 shows a ray of light incident on one face of a parallel-sided glass block. The angle of incidence is 22° and the angle of refraction is 15° .

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Use

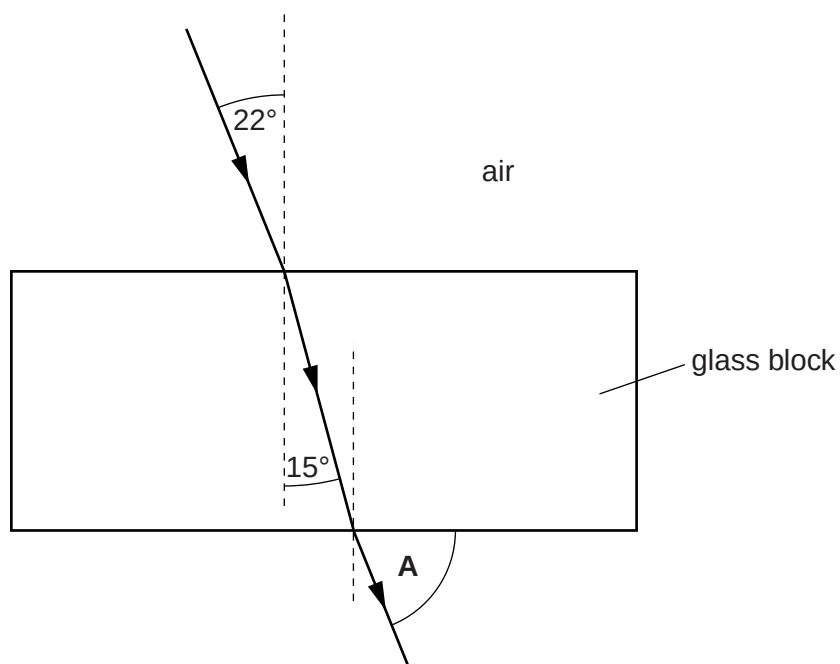


Fig. 4.1

The ray passes through the glass and emerges into air.

- (a) Calculate the refractive index of the glass.

refractive index = [2]

- (b) State the value of the angle **A** shown in Fig. 4.1.

angle = $^\circ$ [1]

- (c) Light travels at different speeds in different materials.

State the speed of light in a vacuum.

speed = m/s [1]

5 Fig. 5.1 shows two food chains **A** and **B**.

The organisms in the food chain are not drawn to the same scale.

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Use

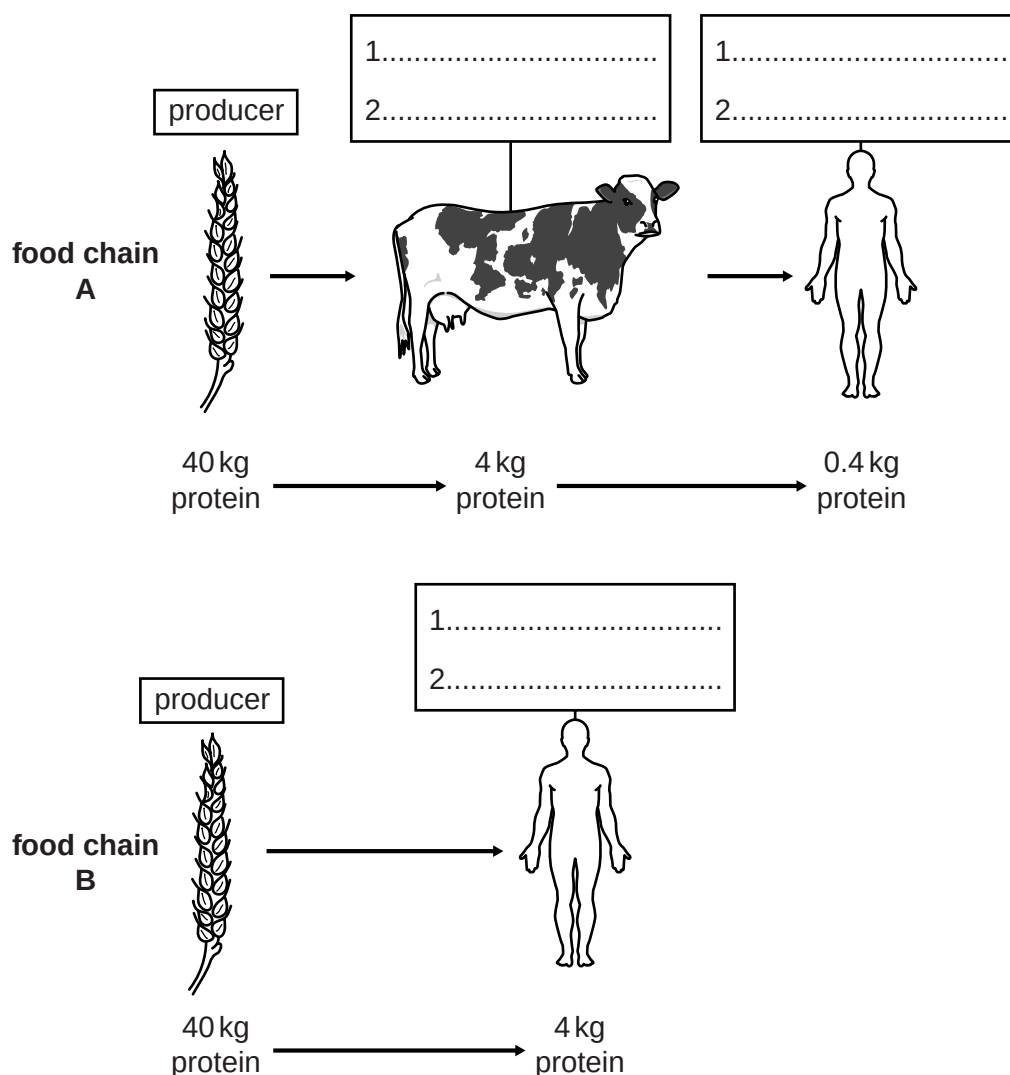


Fig. 5.1

(a) The list below gives roles of organisms in a food chain.

In each box on Fig. 5.1, write two terms **from the list** below that describe the roles of the organisms in each food chain.

Each term may be used once, more than once or not at all.

carnivore decomposer herbivore primary consumer
producer secondary consumer tertiary consumer

[6]

- (b)** Fig. 5.1 shows how an initial mass of protein changes at each stage in each food chain.

Calculate the percentage of the protein in the producer that becomes protein in the human in each food chain.

- (i)** food chain **A**

percentage =%

- (ii)** food chain **B**

percentage =%
[2]

- (c)** Some people think that cereals are a better food source than meat for humans.

State and explain, by referring to part **(b)**, what evidence there is from food chains **A** and **B** to support this view.

.....
.....
.....
.....[2]

- 6 (a) Hydrogen sulfide contains sulfur and hydrogen and has the formula H_2S .

Sulfur is in Group VI of the Periodic Table.

Complete Fig. 6.1 to show the arrangement of the outer shell electrons in a molecule of hydrogen sulfide.

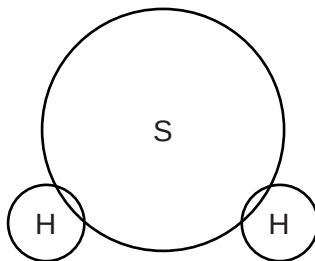
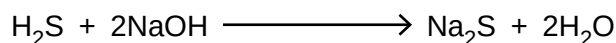


Fig. 6.1

[2]

- (b) Hydrogen sulfide reacts with sodium hydroxide to form sodium sulfide and water.

The equation for the reaction is



The relative molecular mass, M_r , of sodium hydroxide is 40.

[A_r : S, 32; Na, 23; O, 16; H, 1]

Complete the following sentences.

80 g of sodium hydroxide reacts with g of hydrogen sulfide and produces g of sodium sulfide.

8 g of sodium hydroxide produces g of sodium sulfide.

2 g of sodium hydroxide produces g of sodium sulfide.

[4]

7 The following is a list of substances.

ammonium chloride

calcium carbonate

potassium hydroxide

sodium nitrate

sulfur dioxide

For
Examiner's
Use

Use the list to complete the following sentences.

Each substance may be used once, more than once or not at all.

(a) Two substances that react together to produce ammonia are

..... and [2]

(b) The substance that reacts with dilute hydrochloric acid to produce

a salt and water **only** is [1]

(c) The substance that turns Universal Indicator red is

..... [1]

(d) The substance used to reduce acidity in soil is

..... [1]

- 8 A paint gun is used to paint a metal surface, as shown in Fig. 8.1.

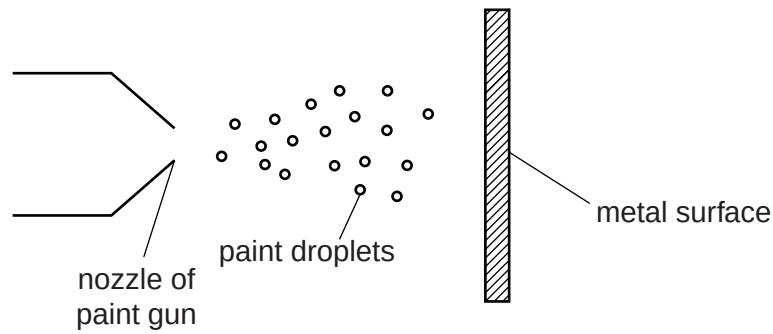


Fig. 8.1

All the paint droplets are given a positive charge as they leave the nozzle.

The metal surface is given a negative charge.

Suggest why

all the droplets are attracted to the metal plate,

.....
 [1]

after leaving the nozzle, the droplets move apart.

.....
 [1]

- 9 Fig. 9.1 shows how the voltage output of a simple a.c. generator changes with time.

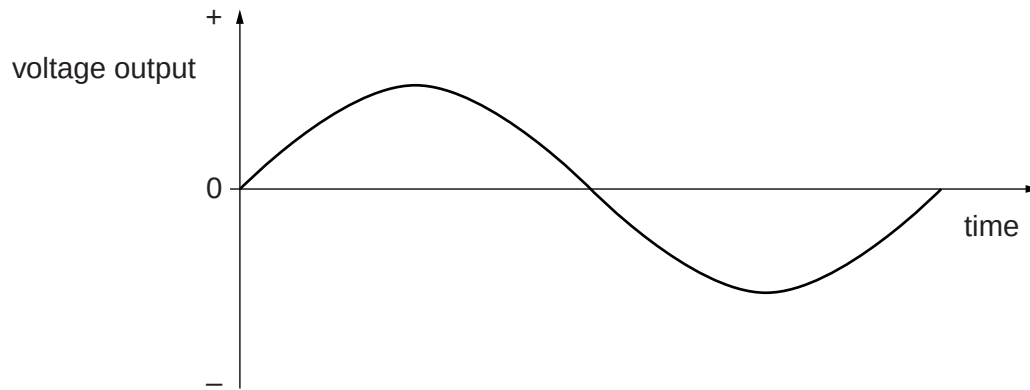


Fig. 9.1

- (a) The coil of the a.c. generator is rotated at twice the original speed.

State the difference, if any, this would make to

- (i) the maximum of the voltage output,

..... [1]

- (ii) the frequency of the voltage output.

..... [1]

- (b) The output of the generator is connected to a transformer.

State and explain why a transformer will **not** work if it is connected to a d.c. supply, rather than an a.c. supply.

.....

 [2]

10 Fig. 10.1 shows an outline of the carbon cycle.

The processes in the cycle are represented by letters.

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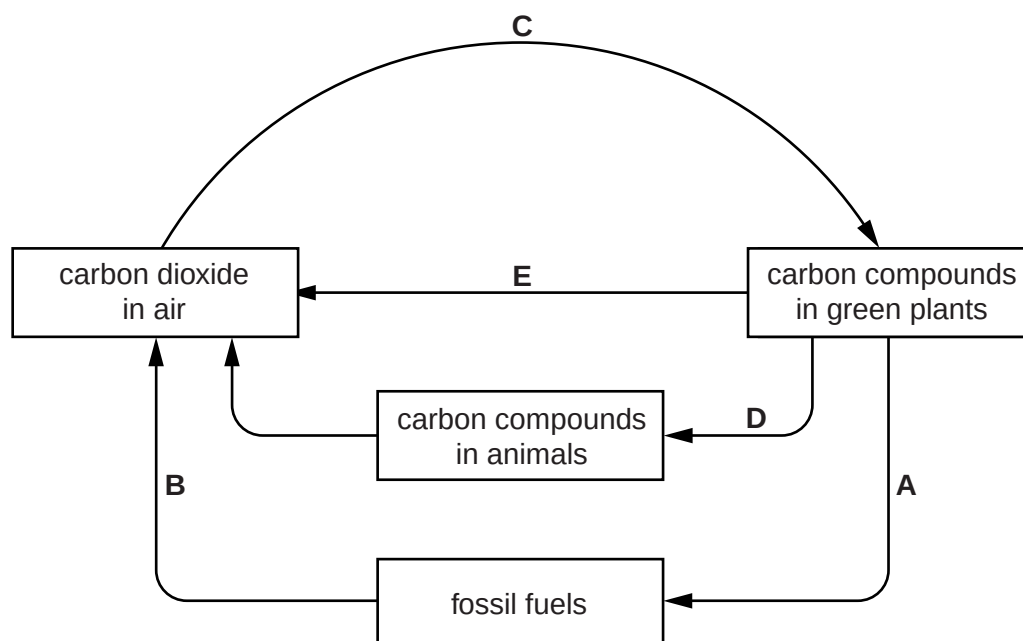


Fig. 10.1

(a) Complete Table 10.1 by naming the process represented by each letter.

An example has been done for you.

Table 10.1

letter	name of process
A	fossilisation
B	
C	
D	
E	

[4]

(b) (i) Complete the word equation for process E.

..... + → + water [3]

(ii) Put a tick in the box by the correct period to show when process E takes place during a 24-hour period.

night and day	
day-time only	
night-time only	

[1]

11 An element X exists as two isotopes, ^{69}X and ^{71}X .

(a) Complete Table. 11.1 for a neutral atom of each isotope.

Table. 11.1

isotope	number of protons	number of neutrons	number of electrons
^{69}X			31
^{71}X	31	40	

[3]

(b) Explain why the chemical properties of the two isotopes are the same.

.....

 [1]

(c) Use the Periodic Table to identify element X.

..... [1]

- 12 A student carries out an experiment using an elastic band to measure its extension for different loads.

For
Examiner's
Use

- (a) In the space below, draw a **labelled** diagram of the apparatus that may be used to obtain an extension-load graph for this elastic band.

[2]

Fig. 12.1 shows an extension-load graph for this elastic band.

Some points have been plotted on the graph.

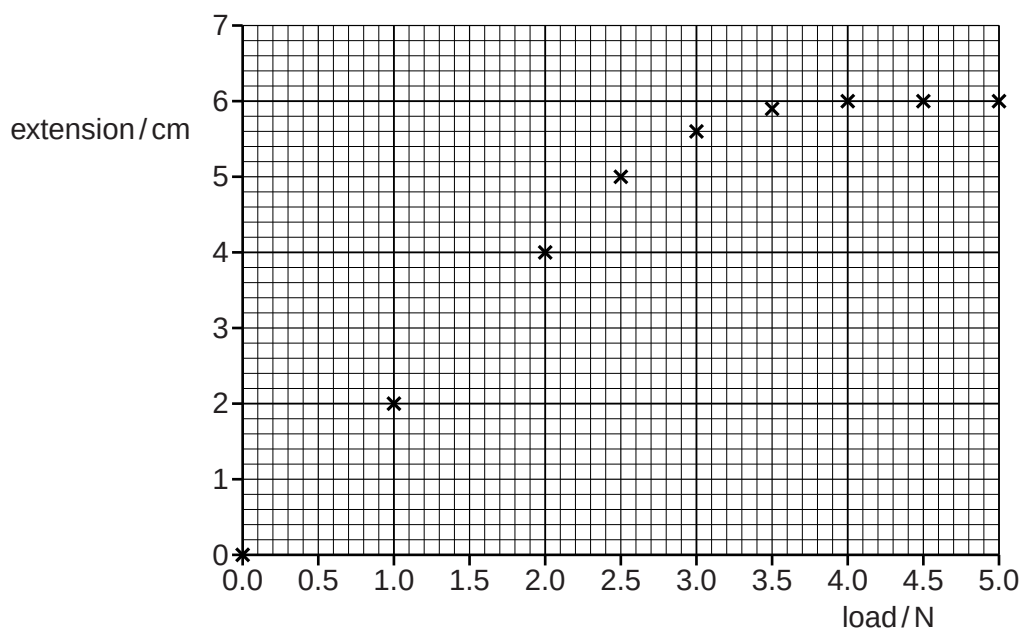


Fig. 12.1

- (b) On Fig. 12.1, draw a line of best fit for the points.

[2]

- (c) With no load on the elastic band, its length is 8.0 cm.

Use Fig. 12.1 to determine the length of the elastic band for a load of 2.0 N.

length = cm [1]

- 13 Fig. 13.1 shows a series circuit containing a 1.5V cell and three resistors.

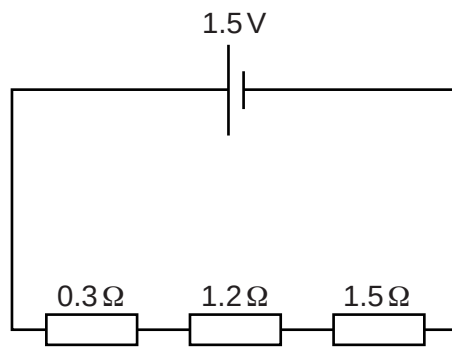


Fig. 13.1

The resistors have values of 0.3Ω , 1.2Ω and 1.5Ω .

- (a) Calculate the combined resistance of the resistors.

resistance = Ω [1]

- (b) Calculate the current in the circuit.

current = A [2]

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- 14** A student carries out an experiment using a metal ball and a metal ring, as shown in Fig. 14.1.

For
Examiner's
Use

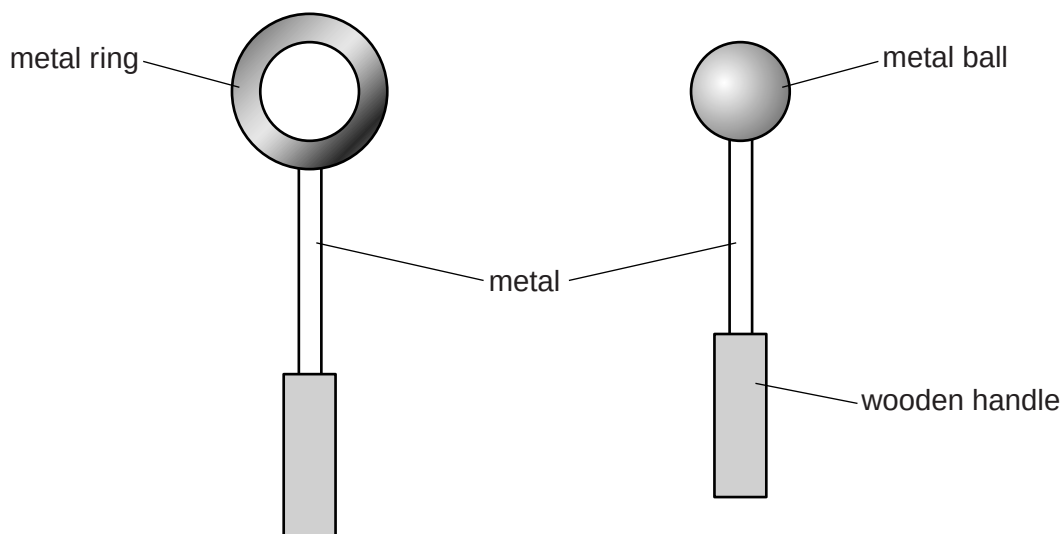


Fig. 14.1

The metal ball will just pass through the metal ring when the ball and ring are at room temperature.

- (a)** Explain why the metal ball will no longer pass through the metal ring when

- (i)** the metal ball is heated and the ring remains at room temperature,

..... [1]

- (ii)** the metal ring is cooled and the ball remains at room temperature.

..... [1]

- (b)** Suggest why the handles are made of wood.

.....
 [1]

15 Fig. 15.1 shows that ethanol may be made from glucose or from ethene.

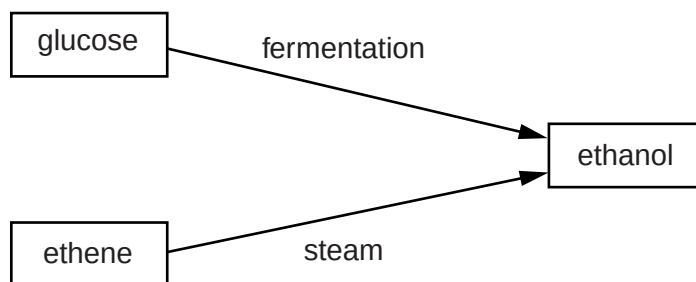


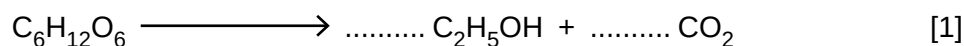
Fig. 15.1

(a) (i) Describe the conditions for the fermentation of glucose to form ethanol.

.....

 [3]

(ii) Balance the equation for the fermentation reaction.



(b) State the type of reaction ethene undergoes when it reacts with steam.

..... [1]

(c) Ethene is an unsaturated hydrocarbon.

Ethane is a saturated hydrocarbon.

(i) State how the structure of ethene differs from the structure of ethane.

..... [1]

(ii) State how the colour of aqueous bromine changes when it is added to ethene.

..... [1]

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16 Fig. 16.1 shows a method of lifting water from a river.

The bucket is raised from the river when a person pushes down on the end of the lever.

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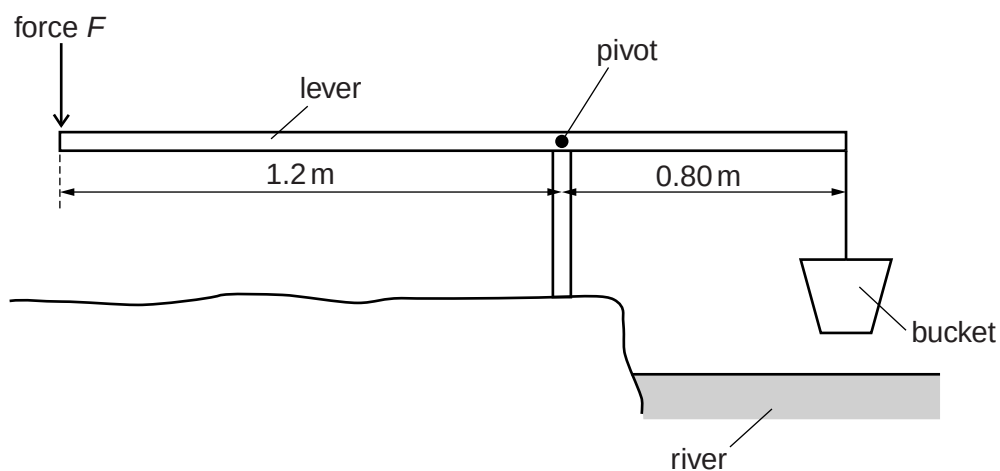


Fig. 16.1

The bucket and its contents weigh 60 N. The bucket is suspended 0.80 m from the pivot.

A man pushes down on the lever with a vertical force F , at a point a distance of 1.2 m from the pivot.

(a) Calculate the force that the man exerts on the end of the lever to keep it horizontal.

force = N [2]

(b) When lifting the bucket and water, the man does 150 J of work in 1.25 s.

Calculate the useful power developed by the man in lifting the bucket and water.

power = W [2]

- 17 (a) Describe the similarities between aerobic respiration and combustion.

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.....

.....

.....

.....

..... [3]

- (b) Name one air pollutant and state the effect of this pollutant on the environment.

pollutant

effect on the environment

..... [2]

18 Fig. 18.1 shows a diagram of the alimentary canal.

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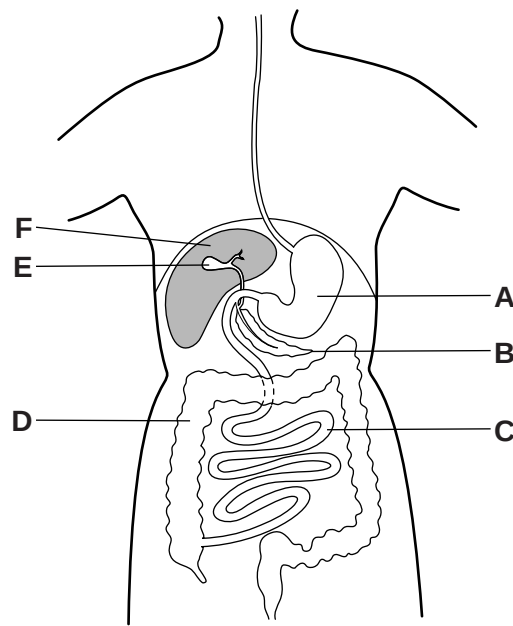


Fig. 18.1

(a) Use the letters on Fig. 18.1 to identify

- (i) the ileum,
- (ii) the stomach,
- (iii) the pancreas.

[3]

(b) State a function for each of the following parts of the digestive system.

stomach

.....

pancreas.....

.....

ileum

.....

colon

.....

[4]

(c) (i) On Fig. 18.1, mark with the letter **X** the organ where bile is made. [1]

(ii) Explain how bile makes digestion more efficient.

.....

.....

.....

.....

.....[3]

19 A nucleus of barium (Ba) contains 56 protons and 81 neutrons.
It is represented by ${}^A_Z\text{Ba}$.

State the value of

(a) A, [1]

(b) Z. [1]

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DATA SHEET

The Periodic Table of the Elements

Group																	
I	II											III	IV	V	VI	VII	0
		1H Hydrogen															
7Li Lithium 3	9Be Beryllium 4																
23Na Sodium 11	24Mg Magnesium 12																
39K Potassium 19	40Ca Calcium 20	45Sc Scandium 21	48Ti Titanium 22	51V Vanadium 23	52Cr Chromium 24	55Mn Manganese 25	56Fe Iron 26	59Co Cobalt 27	59Ni Nickel 28	64Cu Copper 29	65Zn Zinc 30	70Ga Gallium 31	73Ge Germanium 32	75As Arsenic 33	79Se Selenium 34	80Br Bromine 35	84Kr Krypton 36
85Rb Rubidium 37	88Sr Strontium 38	89Y Yttrium 39	91Zr Zirconium 40	93Nb Niobium 41	96Mo Molybdenum 42	99Tc Technetium 43	101Ru Ruthenium 44	103Rh Rhodium 45	106Pd Palladium 46	108Ag Silver 47	112Cd Cadmium 48	115In Indium 49	119Sn Tin 50	122Sb Antimony 51	128Te Tellurium 52	127I Iodine 53	131Xe Xenon 54
133Cs Caesium 55	137Ba Barium 56	139La Lanthanum 57	178Hf Hafnium 72	181Ta Tantalum 73	184W Tungsten 74	186Re Rhenium 75	190Os Osmium 76	192Ir Iridium 77	195Pt Platinum 78	197Au Gold 79	201Hg Mercury 80	204Tl Thallium 81	207Pb Lead 82	209Bi Bismuth 83	209Po Polonium 84	210At Astatine 85	222Rn Radon 86
223Fr Francium 87	226Ra Radium 88	227Ac Actinium 89															
* 58–71 Lanthanoid series † 90–103 Actinoid series																	
aXb		a = relative atomic mass X = atomic symbol b = atomic (proton) number															
Key																	

The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.).