



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
General Certificate of Education Ordinary Level

CANDIDATE
NAME

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CENTRE
NUMBER

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COMBINED SCIENCE

5129/02

Paper 2

October/November 2008

2 hours 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

A copy of the Periodic Table is printed on page 20.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

For Examiner's Use

This document consists of **17** printed pages and **3** blank pages.

- 1 The following is a list of gases.

ammonia argon chlorine hydrogen
nitrogen sulphur dioxide

Complete the following sentences using the gases from the list.

Each gas may be used once, more than once, or not at all.

- (a) A gas that causes erosion of buildings is [1]
- (b) A gas that displaces bromine from potassium bromide is [1]
- (c) A gas used in the manufacture of margarine is [1]
- (d) A gas used in the purification of water supplies is [1]
- (e) A gas that dissolves in water to give an alkaline solution is [1]

- 2 A student connects a cell, a lamp, a variable resistor and an ammeter in series.

- (a) In the space below, draw the circuit diagram.

[3]

- (b) For one setting of the variable resistor, the lamp has normal brightness. The student increases the resistance of the variable resistor.

State what, if anything, happens to

- (i) the potential difference across the lamp,
- (ii) the brightness of the lamp. [2]

3 Fig. 3.1 represents the human alimentary canal.

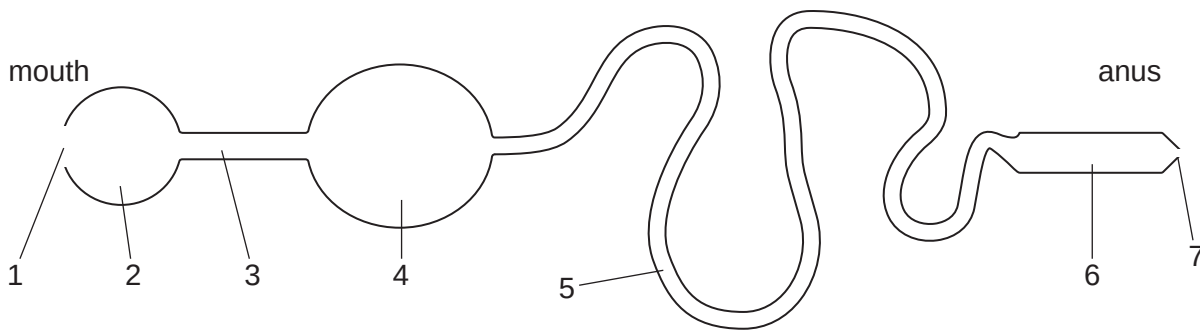


Fig. 3.1

(a) State the number of the region where each of the following processes occurs.

- (i) absorption
- (ii) digestion
- (iii) egestion
- (iv) ingestion [4]

(b) State

- (i) the process that moves food through region 5,
..... [1]

- (ii) the part of a balanced diet that helps this process.
..... [1]

(c) Saliva is secreted into the mouth.
State two functions of saliva.

- 1.
- 2. [2]

4 Air is a mixture of gases.

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(a) Fig. 4.1 shows the percentages of different gases in air.

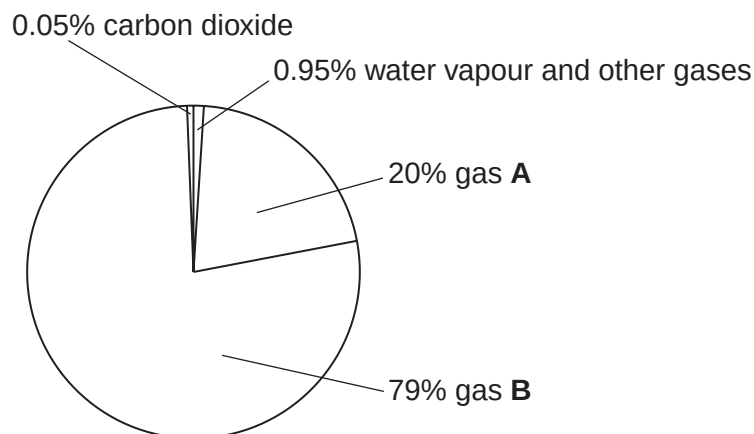


Fig. 4.1

State the name of

gas A,

gas B.

[2]

(b) Suggest why the amount of water vapour present in air is **not** constant.

.....

..... [1]

(c) Name a pollutant gas that may be found in air.
State the source of this gas.

pollutant gas

source

[2]

5 An electric lamp produces heat and light.

(a) Name the form of energy that is converted into heat and light. [1]

(b) A lamp is rated as 100W.

(i) Calculate the amount of energy changed into heat and light in the lamp in 5.0 minutes.

amount of energy = J [2]

(ii) The lamp produces 28 500 J of heat in the 5.0 minutes.

Calculate the amount of energy converted into light by the lamp.

amount of energy = J [1]

6 (a) Insulin is produced by the pancreas. It acts in the liver, where it affects the conversion of blood glucose to stored carbohydrate.

(i) What type of substance is insulin?

..... [1]

(ii) How is insulin transported to the liver?

..... [1]

(iii) Which organ destroys the insulin after it has had its effect?

..... [1]

(b) Name the part of the eye that changes when

(i) a bright light suddenly shines on the eye,

.....

(ii) the eye adjusts from viewing a near object to viewing a distant object.

..... [2]

- 7 Alkanes are obtained by the fractional distillation of petroleum (crude oil).
Alkenes are made by cracking alkanes.

(a) State **one** condition used in the cracking of alkanes. [1]

(b) State how the molecular structure of alkenes differs from the molecular structure of alkanes.

..... [1]

(c) Alkenes may be distinguished from alkanes by adding aqueous bromine.
State the colour change when bromine is added to an alkene.

colour changes from to [2]

(d) Poly(ethene) is made from ethene by polymerisation.
Explain the meaning of the term *polymerisation*.

.....
..... [2]

- 8 A student performs an experiment to obtain an extension-load graph for a spring.

(a) State what measurements are made in order to determine the extension of the spring.

.....
.....
..... [1]

(b) The length of the spring with no load is 8.0 cm.
Fig. 8.1 shows the length of the spring for different loads.
Complete Fig. 8.1 by writing in the missing values of extension and load.

load/N	length/cm	extension/cm
1.0	9.2	
	15.2	7.2

Fig. 8.1

[2]

- 9 (a) The cotyledons of a seed contain an enzyme that digests starch.

(i) Name an enzyme that digests starch.

.....

(ii) Name the product of this digestion.

.....

[2]

- (b) A piece of cotyledon was placed on a jelly containing starch.
A second piece was boiled in water and cooled before being placed on the jelly.

After 24 hours, iodine solution was poured onto the jelly.
Iodine and starch react to give a dark blue/black colour.
Fig. 9.1 shows how the dish looked.

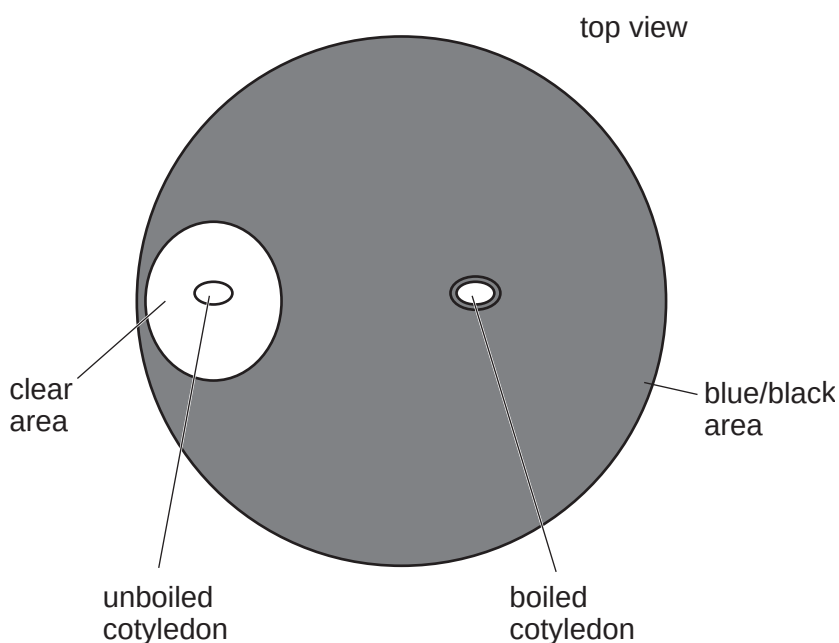


Fig. 9.1

Explain why there is a clear area around the unboiled cotyledon.

.....

.....

..... [2]

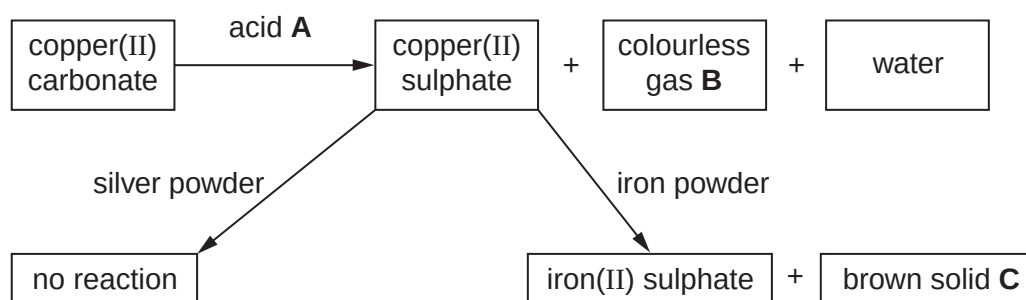
- (c) Explain why there is no clear area around the boiled cotyledon.

.....

..... [2]

10 Study the following reaction scheme.

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(a) Identify the substances **A**, **B** and **C**.

acid **A**

colourless gas **B**

brown solid **C**

[3]

(b) Suggest why the copper(II) sulphate solution reacts with iron but does **not** react with silver.

.....

.....

..... [2]

- 11 Fig. 11.1 shows a magnet moving into a coil of wire. A sensitive ammeter measures the current in the coil.

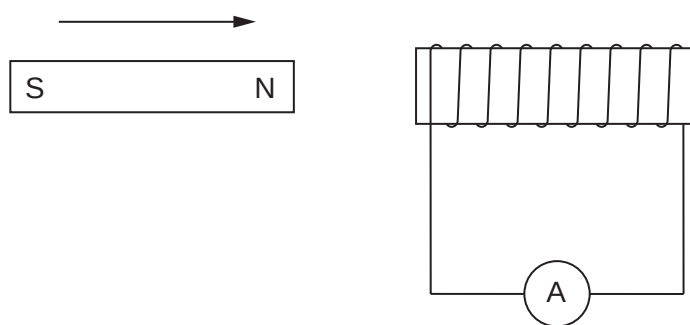


Fig. 11.1

Complete the following sentences using the words below.

You may use each word once, more than once, or not at all.

broken changing constant complete high
induces low produces zero

As the magnet moves into the coil, the magnetic field inside the coil
 an e.m.f. in the circuit. Because the circuit is,
 a current is produced. The current is zero whenever the speed of the magnet is

[4]

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12 Fig. 12.1 represents part of the carbon cycle.

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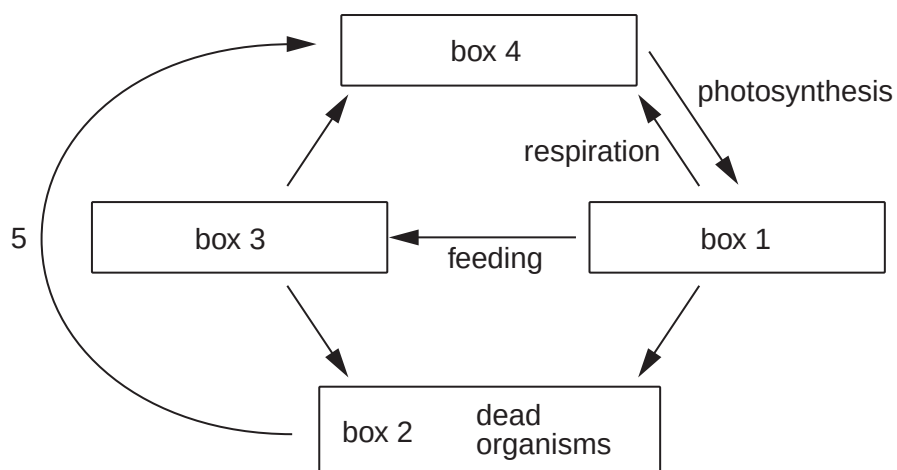


Fig. 12.1

(a) (i) What words should be written in the boxes labelled 1, 3 and 4?

box 1

box 3

box 4

[3]

(ii) State the process that is represented by the arrow labelled 5.

..... [1]

(b) Name the substance in box 1 that contains carbon and is formed by photosynthesis.

..... [1]

- 13** Fig. 13.1 shows a force of 20 N being used to move a block of mass 5.0 kg across a horizontal frictionless surface.

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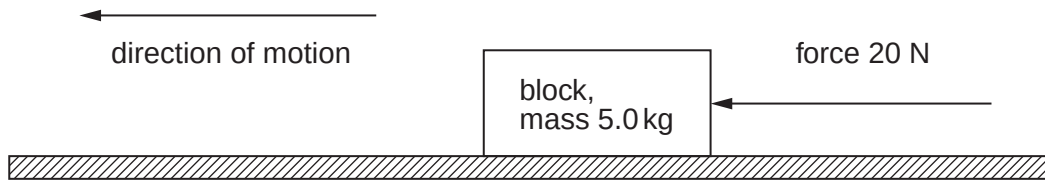


Fig. 13.1

- (a)** Calculate the acceleration of the block.

acceleration = [3]

- (b)** On Earth, the gravitational field strength $g = 10 \text{ N/kg}$.

Calculate the weight of the block.

weight = N [1]

- (c)** The force of 20 N moves the block a distance of 40 cm.

Calculate the work done by the force.

work done = [2]

- 14** A rocket explodes 99 m above the ground. Both light and sound are produced at the same time.

The speed of sound is 330 m/s.

- (a) (i)** Explain why the explosion of the rocket is seen by a boy on the ground before he hears it exploding.

.....
 [1]

- (ii)** Calculate the shortest time that the sound takes to reach the ground.

time = s [2]

- (b)** The light emitted by the exploding rocket has many different frequencies.

State what is meant by the *frequency* of a wave.

.....
 [1]

- (c)** Visible light is one component of the electromagnetic spectrum.

Name a component

- (i)** with a longer wavelength than visible light,

- (ii)** that is emitted from some radioactive nuclei.

[2]

15 Fig. 15.1 shows a cross-section through a flower.

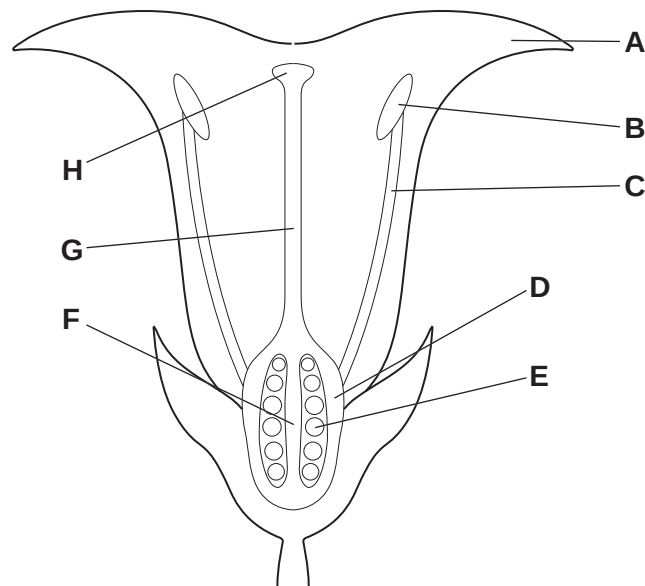


Fig. 15.1

- (a) (i) Give the letters of the structures that contain cells with nuclei which fuse during sexual reproduction?

..... and [2]

- (ii) Name the structure that is formed by the fusion of the nuclei of these two cells.

..... [1]

- (b) Into what do each of the structures labelled **D** and **E** develop?

D develops into

E develops into [2]

16 ^{16}O and ^{18}O are isotopes of the element oxygen.

(a) State how the two isotopes are different.

.....
 [1]

(b) Explain why the two isotopes have the same chemical properties.

.....
 [2]

(c) State the electronic structure of ^{16}O [1]

(d) State the relationship between the electronic arrangement of an oxygen atom and the group that oxygen belongs to in the Periodic Table.

.....
 [1]

17 Fig. 17.1 shows a root hair cell.

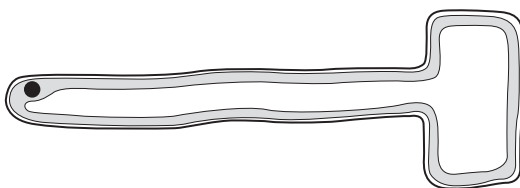


Fig. 17.1

(a) State two substances that are absorbed by a root hair cell.

1.
 2. [2]

(b) State two features of a root hair cell that make it suitable for absorption.

1.

 2.
 [2]

- 18 (a) Fig. 18.1 shows a ray of light from a pin that is incident on a plane mirror at point A.

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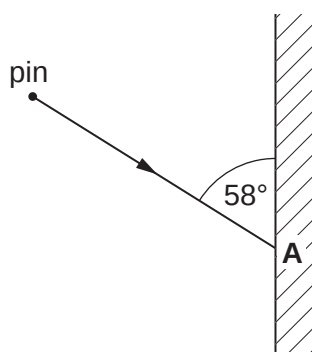


Fig. 18.1

- (i) On Fig. 18.1, draw the normal to the mirror at A. [1]
- (ii) On Fig. 18.1, mark the position of the image of the pin as seen in the mirror. [2]
- (iii) Calculate the angle of incidence.

angle of incidence =° [1]

- (b) Fig. 18.2 shows a ray of light entering a parallel-sided glass block.



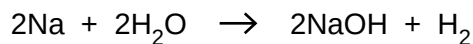
Fig. 18.2

Complete Fig. 18.2 to show the path of the ray inside the block and after leaving the block. [2]

- 19 (a) Define *relative atomic mass*, A_r .

.....
 [2]

- (b) Sodium reacts with water to produce sodium hydroxide solution and hydrogen.
 The equation for the reaction is



- (i) Calculate the relative molecular mass of sodium hydroxide.
 (A_r : H, 1; Na, 23; O, 16.)

..... [1]

- (ii) Calculate the mass of water required to react with 4.6 g of sodium.
 (A_r : Na, 23.)

mass = g [2]

- (c) State a test to show that sodium hydroxide solution is alkaline.

test

result [2]

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DATA SHEET

The Periodic Table of the Elements

Group																																			
I	II											III	IV	V	VI	VII	0																		
		1 H Hydrogen																4 He Helium																	
3 7 Li Lithium	4 9 Be Beryllium											5 11 B Boron	6 12 C Carbon	7 14 N Nitrogen	8 16 O Oxygen	9 19 F Fluorine	10 20 Ne Neon																		
11 23 Na Sodium	12 24 Mg Magnesium											13 27 Al Aluminium	14 28 Si Silicon	15 31 P Phosphorus	16 32 S Sulphur	17 35.5 Cl Chlorine	18 40 Ar Argon																		
19 39 K Potassium	20 40 Ca Calcium	21 45 Sc Scandium	22 48 Ti Titanium	23 51 V Vanadium	24 52 Cr Chromium	25 55 Mn Manganese	26 56 Fe Iron	27 59 Co Cobalt	28 59 Ni Nickel	29 64 Cu Copper	30 65 Zn Zinc	31 70 Ga Gallium	32 73 Ge Germanium	33 75 As Arsenic	34 79 Se Selenium	35 80 Br Bromine	36 84 Kr Krypton																		
37 85 Rb Rubidium	38 88 Sr Strontium	39 89 Y Yttrium	40 91 Zr Zirconium	41 93 Nb Niobium	42 96 Mo Molybdenum	43 98 Tc Technetium	44 101 Ru Ruthenium	45 103 Rh Rhodium	46 106 Pd Palladium	47 108 Ag Silver	48 112 Cd Cadmium	49 115 In Indium	50 119 Sn Tin	51 122 Sb Antimony	52 128 Te Tellurium	53 127 I Iodine	54 131 Xe Xenon																		
55 133 Cs Caesium	56 137 Ba Barium	57 139 La Lanthanum	72 178 Hf Hafnium	73 181 Ta Tantalum	74 184 W Tungsten	75 186 Re Rhenium	76 190 Os Osmium	77 192 Ir Iridium	78 195 Pt Platinum	79 197 Au Gold	80 201 Hg Mercury	81 204 Tl Thallium	82 207 Pb Lead	83 209 Bi Bismuth	84 210 Po Polonium	85 210 At Astatine	86 210 Rn Radon																		
87 226 Fr Francium	88 226 Ra Radium	89 227 Ac Actinium																																	
58-71 Lanthanoid series 79-103 Actinoid series																																			
		a		X		a = relative atomic mass X = atomic symbol b = proton (atomic) number				140 Ce Cerium		141 Pr Praseodymium		144 Nd Neodymium		150 Sm Samarium		152 Eu Europium		157 Gd Gadolinium		159 Tb Terbium		162 Dy Dysprosium		165 Ho Holmium		167 Er Erbium		169 Tm Thulium		173 Yb Ytterbium		175 Lu Lutetium	
										232 Th Thorium		238 Pa Protactinium		238 U Uranium		238 Np Neptunium		238 Am Americium		238 Cm Curium		238 Bk Berkelium		238 Cf Californium		238 Es Einsteinium		238 Fm Fermium		238 Md Mendelevium		238 No Nobelium		238 Lr Lawrencium	

*58-71 Lanthanoid series
†90-103 Actinoid series

Key

a	X	b
a = relative atomic mass	X = atomic symbol	b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).