

Centre Number	Candidate Number	Name
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UNIVERSITY OF CAMBRIDGE LOCAL EXAMINATIONS SYNDICATE
General Certificate of Education Ordinary Level

SCIENCE

5124/03, 5126/03

Paper 3 Chemistry

October/November 2006

1 hour 15 minutes

Additional Materials: Answer Booklet/Paper

READ THESE INSTRUCTIONS FIRST

If you have been given an Answer Booklet, follow the instructions on the front cover of the booklet.
Write your Centre number, candidate number and name on all the work you hand in.
Write in dark blue or black pen.
You may use a soft pencil for any diagrams, graphs, tables or rough working.
Do not use staples, paper clips, highlighters, glue or correction fluid.

Section A

Answer **all** questions.

Write your answers in the spaces provided on the question paper.

Section B

Answer any **two** questions.

Write your answers on the lined pages provided and, if necessary, continue on separate answer paper.

A copy of the Periodic Table is printed on page 16.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

FOR EXAMINER'S USE	
Section A	
Section B	
TOTAL	

This document consists of **12** printed pages and **4** lined pages.



Section A

Write your answers in the spaces provided on the question paper.

Answer **all** the questions.

1 (a) Name **three** of the components of clean, dry air.

(i)

(ii)

(iii)

[3]

(b) The air can be polluted by various chemicals.

(i) Give the chemical name for **one** of these pollutants.

.....

(ii) Describe how this pollutant enters the air.

.....

.....

[3]

- 2 A sample of water contains salt as an impurity. The apparatus shown in Fig. 2.1 is used to produce pure water from the sample.

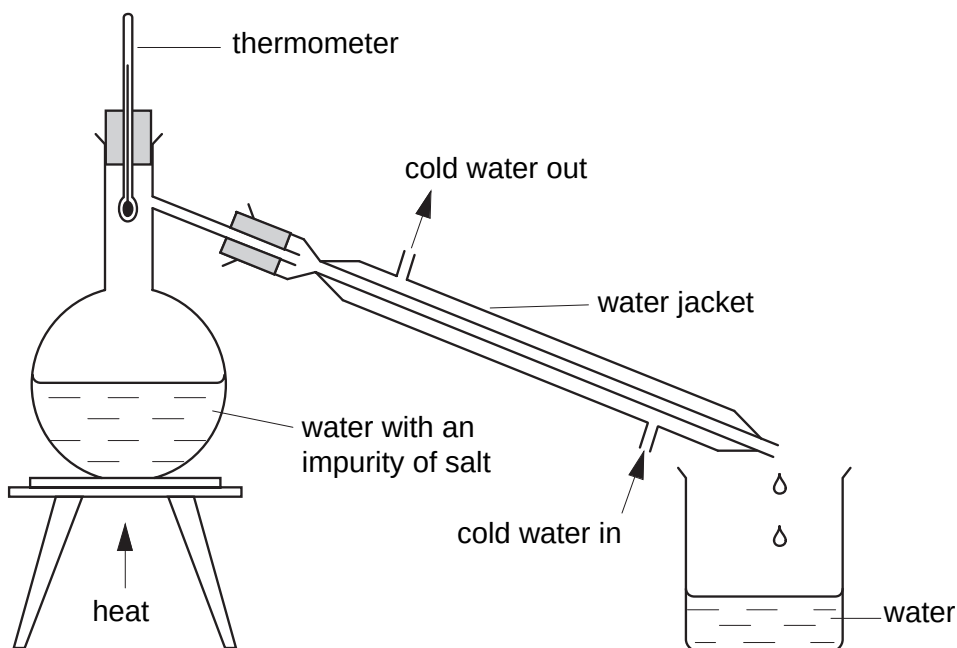


Fig. 2.1

- (a) (i) Name the method of purification.

.....

- (ii) Suggest the purpose of the water jacket.

.....

[2]

- (b) What would be the approximate reading on the thermometer during the purification?

.....

[1]

- (c) Draw a cross (X) on Fig. 2.1 where the salt would be left after purification is complete.

[1]

- 3 (a) Table 3.1 describes **two** plastics. Complete the table. Part of the table has been completed for you as an example.

Table 3.1

name	repeating unit	use	type of polymerisation used in manufacture
poly(ethene)		making clingfilm	
	$\begin{array}{c} \text{O} \qquad \qquad \text{O} \\ \parallel \qquad \parallel \\ -\text{C}-\square-\text{C}-\text{N}-\square-\text{N}- \\ \qquad \qquad \\ \text{H} \qquad \qquad \text{H} \end{array}$		condensation polymerisation

[4]

- (b) The careless disposal of both plastics and iron or steel causes pollution problems. An article made from one of the plastics in Table 3.1 is likely to cause pollution for a longer period of time than a similar article made from iron or steel. Explain why.

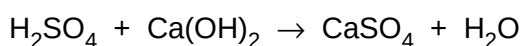
.....

.....

..... [2]

- 4 A spillage of 19.6 tonnes of sulphuric acid results from an accident to a road tanker. Slaked lime is used to neutralise the acid.

- (a) The **unbalanced** chemical equation for the neutralisation is as follows.



Balance this equation.

[1]

- (b) Calculate the relative molecular mass of slaked lime, $\text{Ca}(\text{OH})_2$.

[Relative atomic masses: A_r : H, 1; O, 16; Ca, 40.]

..... [1]

(c) Use the balanced chemical equation to determine

- (i) the mass of slaked lime needed to neutralise the 19.6 tonnes of spilt acid,
[Relative atomic masses: A_r : H, 1; O, 16; S, 32; Ca, 40.]

mass = tonnes

- (ii) the mass of calcium sulphate formed during the neutralisation of the spilt acid.
[Relative atomic masses: A_r : H, 1; O, 16; S, 32; Ca, 40.]

mass = tonnes
[3]

(d) What test could be used to show that the acid has all been neutralised?

.....
..... [1]

(e) The calcium sulphate formed in this neutralisation is insoluble in water. Suggest why this is important.

.....
..... [1]

- 5 Fig. 5.1 shows part of the Periodic Table of the elements. Use information from Fig. 5.1 to answer the questions that follow. The elements are represented by their chemical symbols.

Group																	
I	II											III	IV	V	VI	VII	0
								H									He
Li	Be											B	C	N	O	F	Ne
Na	Mg											Al	Si	P	S	Cl	Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr

Fig. 5.1

- (a) Give the symbol for

(i) a halogen,

(ii) an alkali metal. [2]

- (b) Oxygen, sulphur and selenium are in Group VI. At room temperature oxygen is a gas and sulphur is a solid. Predict whether selenium is, at room temperature, a gas or a liquid or a solid.

..... [1]

- (c) The trend in reactivity in Group VI is similar to that in Group VII. Suggest which is the most reactive element in Group VI.

..... [1]

- (d) Write the formula for a compound that is formed when

(i) an element from Group I reacts with an element from Group VI,

.....

(ii) an element from Group II reacts with an element from Group VI.

..... [2]

6 Fig. 6.1 shows details of four alcohols.

alcohol	molecular formula	molecular mass	boiling point / °C
methanol	CH ₃ OH	32	65
ethanol	C ₂ H ₅ OH	46	79
propan-1-ol	C ₃ H ₇ OH	60	97
butan-1-ol	C ₄ H ₉ OH	74	117

Fig. 6.1

(a) The four alcohols in Fig. 6.1 are members of the same homologous series. The next in this series of alcohols is pentan-1-ol.

(i) Predict the approximate boiling point of pentan-1-ol.

.....

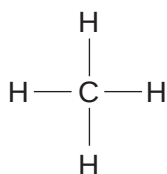
(ii) Determine the relative molecular mass of pentan-1-ol.

[Relative atomic masses: A_r: H, 1; C, 12; O, 16.]

.....

[2]

(b) The structural formula for methane is drawn as



Draw the structural formula for ethanol, C₂H₅OH.

[2]

(c) Ethanol, $\text{C}_2\text{H}_5\text{OH}$, is burnt as a fuel.

(i) Name **two** of the products of burning ethanol in excess oxygen.

.....

.....

(ii) Write a chemical equation for this burning of ethanol. State symbols are **not** required.

.....
[4]

7 Four metals are represented by the letters **A**, **B**, **C** and **D**: these are **not** chemical symbols. Their reactions with cold water and dilute hydrochloric acid are summarised in Fig. 7.1.

metal	reaction with cold water	reaction with dilute hydrochloric acid
A	none	none
B	none	slow
C	fast	fast
D	slow	fast

Fig. 7.1

(a) Place the metals **A**, **B**, **C** and **D** in order of reactivity.

most reactive

.....

.....

least reactive

[1]

(b) Which of the metals **A**, **B**, **C** or **D** could be

(i) sodium,

(ii) copper?

[2]

(c) Suggest which of the metals **A**, **B**, **C** and **D** would probably be the easiest to extract from its naturally occurring ore.

.....
[1]

- 8 Hydrogen peroxide solution decomposes in the presence of a catalyst, producing oxygen gas.

The rate of this reaction can be found by plotting total volume of oxygen evolved against time.

This graph is shown in Fig. 8.1.

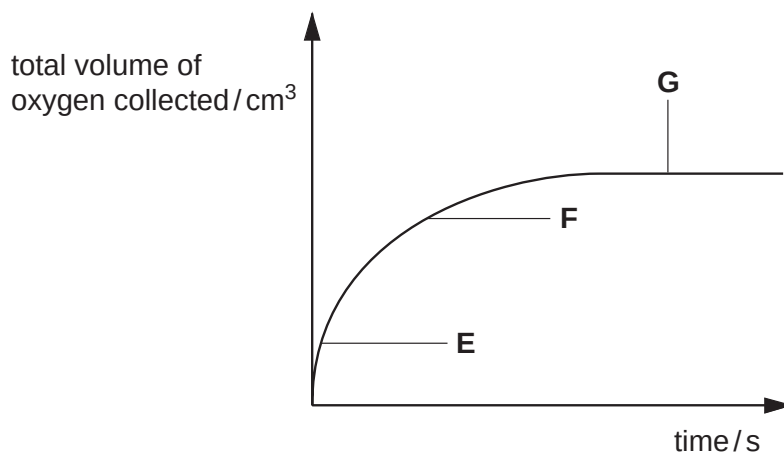


Fig. 8.1

- (a) Describe the rate of decomposition at points E, F and G.

(i) E,

(ii) F,

(iii) G. [3]

- (b) How does the rate of decomposition depend upon the number of molecules of hydrogen peroxide present in the solution?

..... [1]

Section B

Answer any **two** questions.

Write your answers on the lined pages provided and, if necessary, continue on separate answer paper.

- 9 Substance **H** is a mixture of three sodium salts. Fig. 9.1 shows a description written by students of how they attempted to identify the three salts.

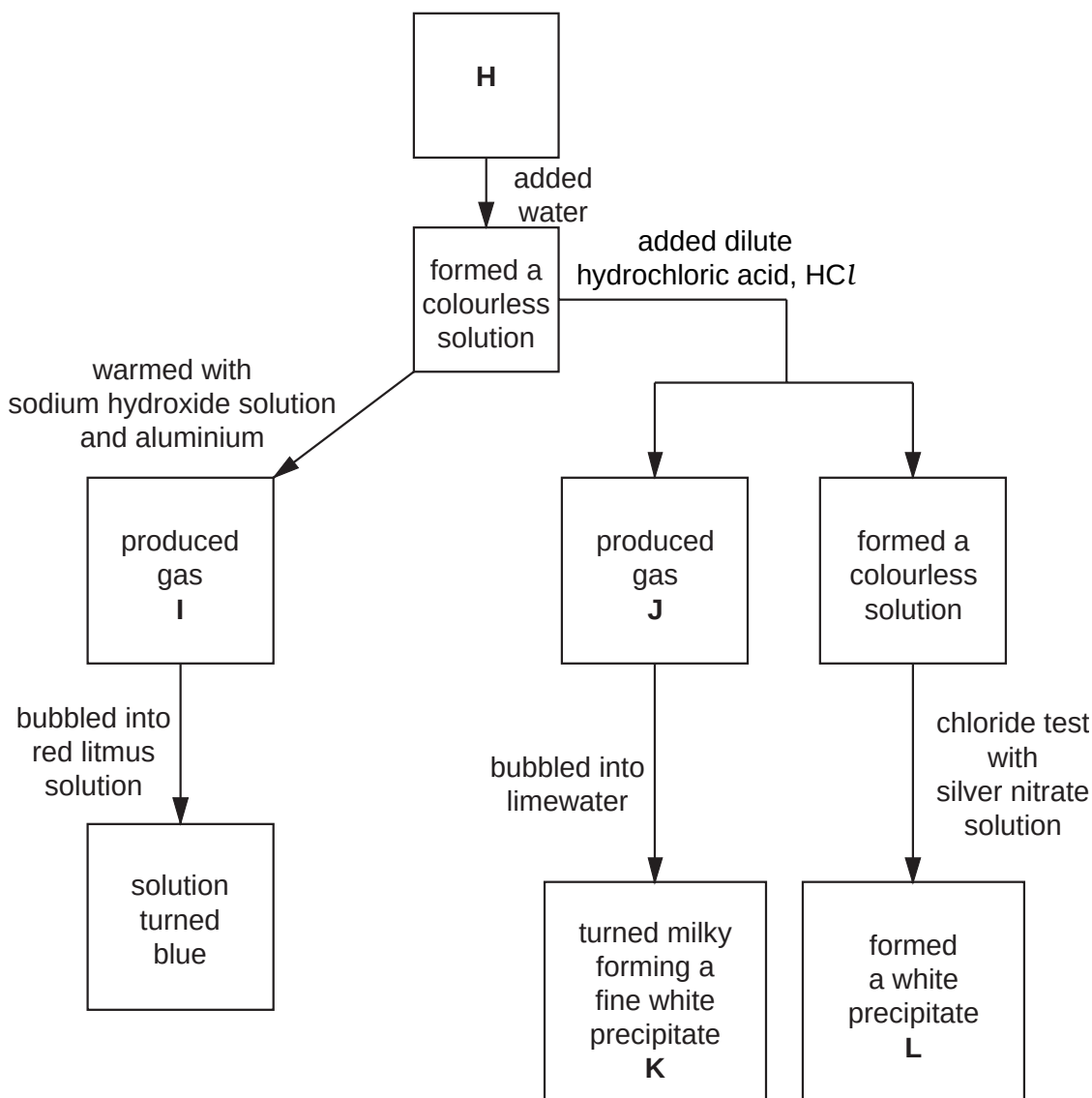


Fig. 9.1

- (a) (i) Name the gases **I** and **J** and the white precipitates **K** and **L**. [4]
- (ii) The formation of white precipitate **L** shows the presence of chloride ions. [2]
Why does this not prove that chloride ions are present in substance **H**? [2]
- (b) What two sodium salts must be present in the substance **H**? [2]
- (c) Write a chemical equation to represent any **one** of the reactions shown in Fig. 9.1. State symbols are **not** required. [2]

- 10 (a)** Name a mixture that is gaseous, a compound that is a liquid and an element that is a solid, at room temperature and pressure. [3]
- (b)** For **each** of the substances you have identified in **(a)**,
- (i)** name the atoms within the substance which are bonded together as molecules, if any,
 - (ii)** describe how the particles move within that substance. [7]
- 11 (a)** Iron is manufactured in a blast furnace using an iron ore, coke and limestone.
- Name the ore and give the formula for the main iron compound in this ore. [2]
- (b)** Describe the essential chemical reactions that take place in the blast furnace. Include chemical equations in your description. [5]
- (c)** The properties of a metal can be changed by alloying the metal with other elements.
- Name an alloy, state its components and give one of its uses. [3]

[illegible]

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[illegible]

[illegible]

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The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.).