



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CANDIDATE
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CO-ORDINATED SCIENCES

0654/22

Paper 2 (Core)

October/November 2011

2 hours

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs, tables or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

A copy of the Periodic Table is printed on page 24.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

For Examiner's Use	
1	
2	
3	
4	
5	
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7	
8	
9	
Total	

This document consists of **21** printed pages and **3** blank pages.



1 Fig. 1.1 shows five insects.

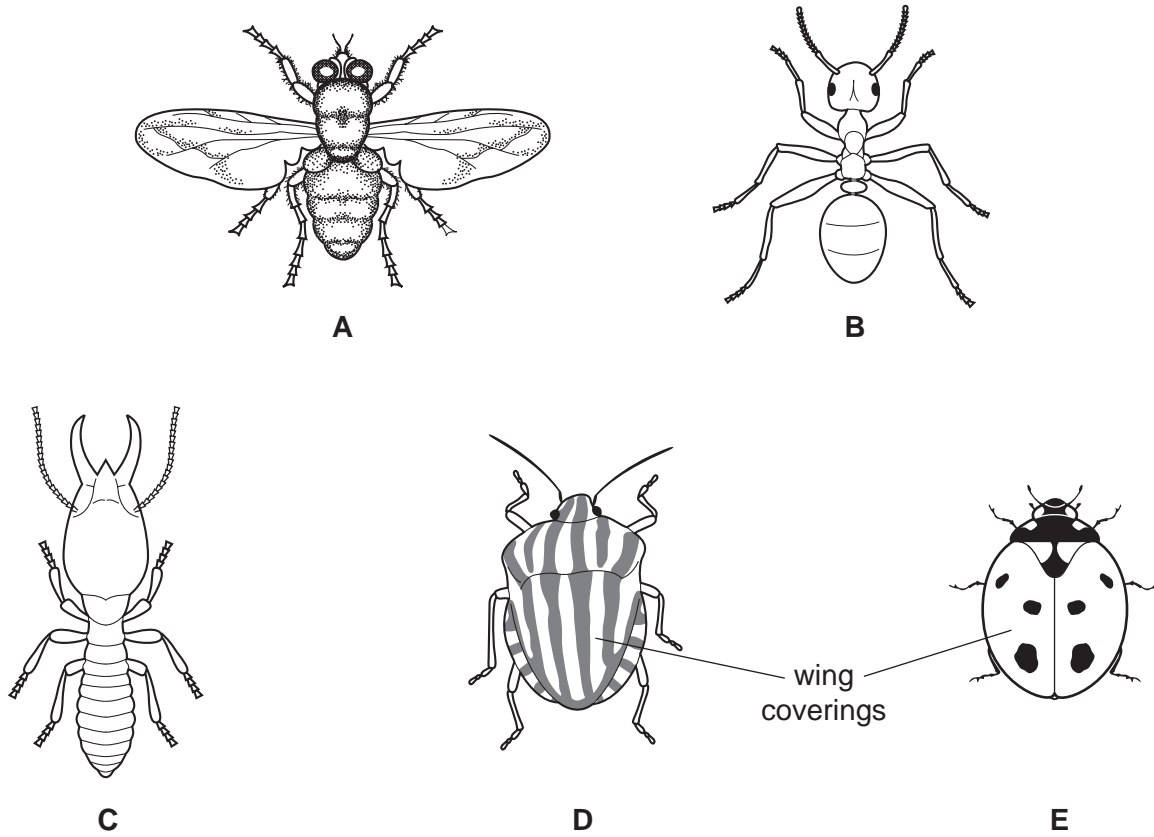


Fig. 1.1

(a) Use the key below to identify each insect. Fill in the table to show how you arrived at your identifications. The first one has been done for you.

- 1a has wings go to 2
- 1b does not have wings go to 3
- 2a wings are covered go to 4
- 2b wings are not covered *Musca*
- 3a head longer than front leg *Termes*
- 3b head shorter than front leg *Formica*
- 4a striped pattern on wing coverings *Graphosoma*
- 4b spots on wing coverings *Coccinella*

insect	1		2		3		4		name
	a	b	a	b	a	b	a	b	
A	✓			✓					<i>Musca</i>
B									
C									
D									
E									

[4]

(b) The common name for insect **A** is housefly. The complete binomial of insect **A** is *Musca domestica*.

Suggest why scientists may prefer to use the binomial of an insect, rather than its common name.

.....
..... [1]

(c) Houseflies feed by spitting saliva onto food, such as meat. Enzymes in the saliva turn insoluble substances into soluble ones. The flies can then suck the liquid into their digestive system.

(i) Suggest **one** enzyme in a housefly's saliva that could digest a substance in meat.

..... [1]

(ii) State the soluble product or products that this enzyme would produce.

..... [1]

(d) Houseflies spread diseases such as typhoid fever. They leave harmful microorganisms on food that will later be eaten by a person.

(i) Name the cells in the human body that can help to prevent microorganisms causing infections.

..... [1]

(ii) Pesticides are sometimes used to kill houseflies and therefore reduce the risk of spreading disease.

Give **one** reason why pesticides should **not** be used more than necessary.

.....
..... [1]

(e) When a housefly flies, its wings produce a buzzing sound.

(i) Suggest how a movement such as that of a fly's wings produces sound.

.....
.....
..... [2]

- (ii) A housefly beats its wings about 200 times per second. A midge (a small insect) beats its wings about 1000 times per second.

State and explain how the sound produced by a flying midge will differ from the sound produced by a flying housefly.

.....

.....

..... [2]

*For
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2 (a) Fig. 2.1 shows an aircraft moving along a runway.

(i) Draw and label arrows on Fig. 2.1 to show the directions of the driving and friction forces acting on the aircraft. [1]

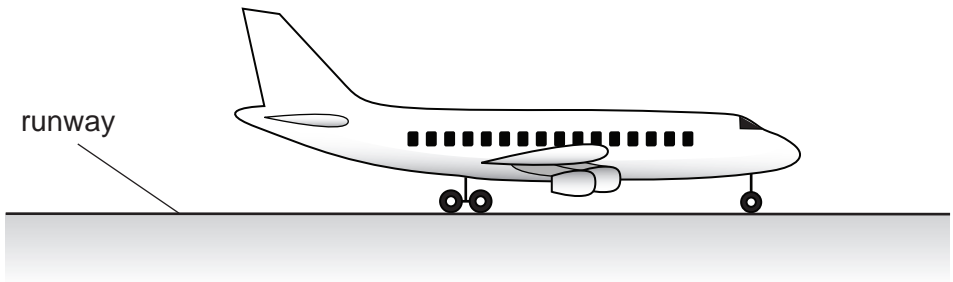


Fig. 2.1

(ii) The driving and friction forces are balanced.

Explain what is meant by the phrase *forces are balanced*.

.....
..... [1]

(iii) Describe the movement of the aircraft when these forces are balanced.

.....
..... [1]

(b) People who fly frequently have greater exposure to ionising radiation than those who do not fly.

(i) Explain why exposure to ionising radiation may be harmful.

.....
.....
..... [2]

(ii) This ionising radiation is cosmic radiation from outer space. This is one source of background radiation.

State **one** other natural source of background radiation.

..... [1]

- (c) The aircraft is able to navigate using radar. This involves using microwaves. These are part of the electromagnetic spectrum.

Name **one** other wave which is part of the electromagnetic spectrum and give a use for this radiation.

name

use [2]

- (d) Potato snacks are packed in airtight packets and filled with nitrogen gas at atmospheric pressure.



- (i) Suggest why nitrogen gas is used rather than air.

.....
..... [2]

- (ii) A passenger has a packet of potato snacks in his hand luggage on the aircraft. During the flight, the aircraft cabin is at a pressure less than normal atmospheric pressure.

The passenger notices that the packet has expanded.

State why this happens.

.....
..... [1]

3 Hydrocarbons are compounds which contain only the elements hydrogen and carbon.

For
Examiner's
Use

(a) (i) State the number of electrons in the outer shell of a carbon atom.

..... [1]

(ii) Another element, **X**, has atoms whose nuclei contain 14 protons.

Name element **X** and explain whether or not atoms of **X** have the same number of outer electrons as a carbon atom.

name of element **X**

explanation

..... [2]

(iii) Name the **least** reactive element which is in the same period of the Periodic Table as carbon.

..... [1]

(b) The simplest hydrocarbon is methane which is an important gaseous fuel.

(i) State **two** natural sources of methane.

1

2 [2]

(ii) A fuel such as methane combines with oxygen in a chemical reaction. When the reaction is occurring, a large amount of heat is given off each second.

Suggest and explain which **one** of the sentences, **A** to **D**, accurately describes the reaction between a typical fuel and oxygen.

A The reaction is endothermic and has a very high rate.

B The reaction is exothermic and has a very high rate.

C The reaction is exothermic and has a very low rate.

D The reaction is endothermic and has a very low rate.

sentence

explanation

.....

..... [2]

(c) Some types of oil and grease contain hydrocarbons. Oil and grease stick to clothes and make them look dirty. Washing with water alone does not remove oil and grease from clothes.

For
Examiner's
Use

(i) State the type of substance which could be added to water so that washing does remove oil and grease.

..... [1]

(ii) Suggest **one** possible disadvantage, other than cost, of using large amounts of the substance given in answer (c)(i) over a long period of time.

.....

..... [1]

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Please turn over for Question 4.

4 Yaks are animals that live in the cold mountainous region of the Himalayas.

Fig. 4.1 shows a yak.

For
Examiner's
Use

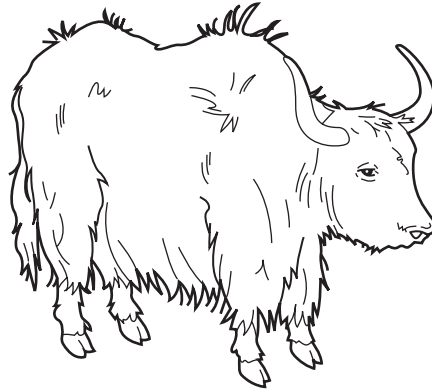


Fig. 4.1

(a) Explain how the long hair of the yak keeps it warm during the cold weather.

.....
.....
..... [2]

(b) Yaks are used as 'beasts of burden'. They can be ridden or used to carry or pull heavy objects.

A yak of mass 1000 kg is carrying a load of 80 kg.

(i) Calculate the combined weight of the yak and load.

The Earth's gravitational field strength is 10 N/kg.

..... N [1]

- (ii) The yak carries its load up a mountain slope and finishes 100 m higher up the mountain.

Calculate the work done in gaining this height.

State the formula that you use and show your working.

formula used

working

..... J [2]

- (iii) While the yak is carrying the load, it travels at a speed of 0.2 m/s.

Calculate the kinetic energy of the yak and its load at this time.

State the formula that you use and show your working.

formula used

working

..... J [2]

- (iv) The yak then travels 1000 m in 4000 s.

Calculate the average speed of the yak.

State the formula that you use and show your working.

formula used

working

..... m/s [2]

(c) Yak dung is commonly burned as a fuel in the Himalayas. However, it is slowly being replaced by kerosene.

State and explain **one** environmental reason why yak dung is a better fuel than kerosene.

.....
..... [1]

*For
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Please turn over for Question 5.

5 Fig. 5.1 shows two plants that are grown as crops.

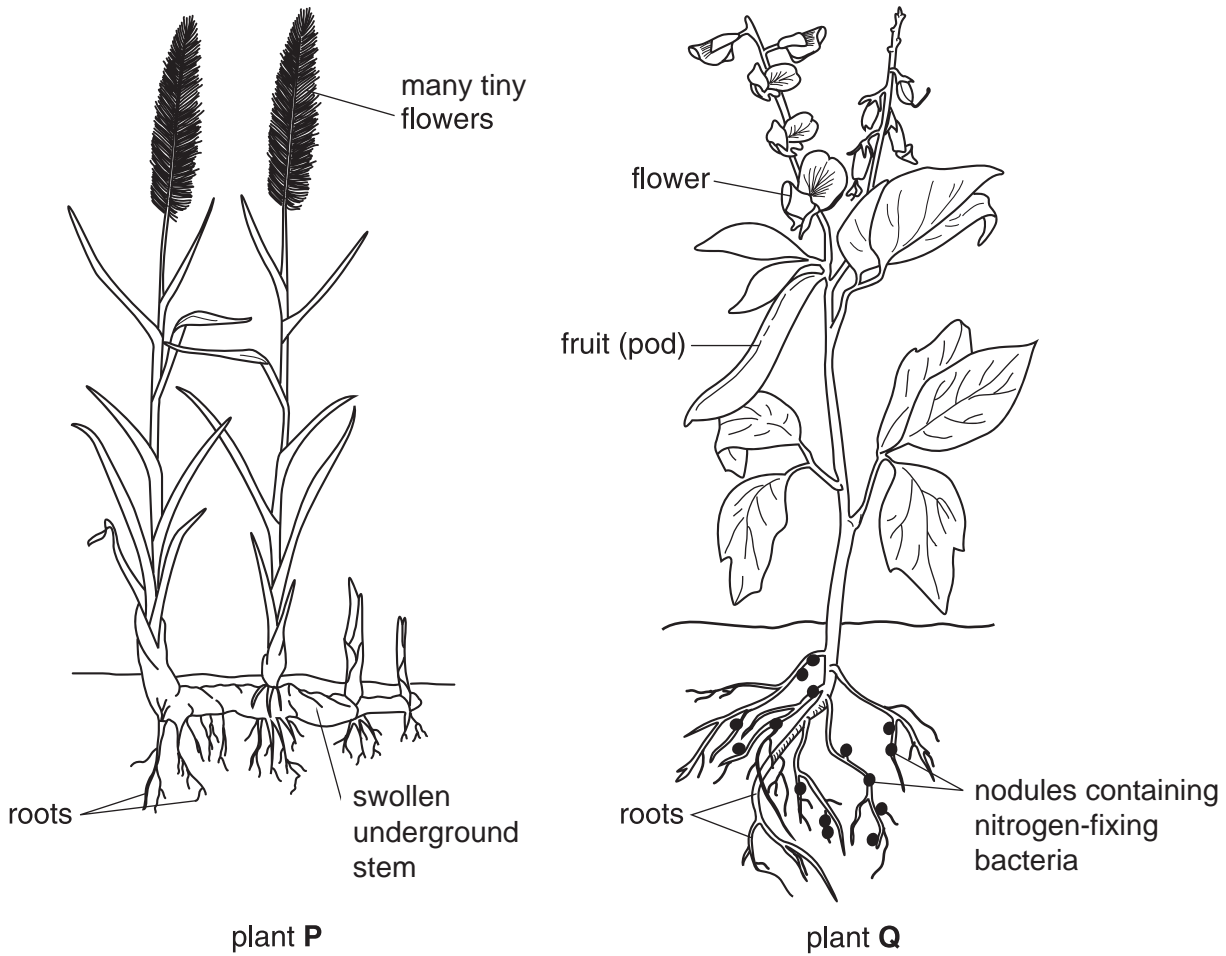


Fig. 5.1

(a) Plant P reproduces using flowers and also by growing new plants from its underground stem.

State the **type** of reproduction that each of these processes involves.

using flowers

growing new plants from underground stems [1]

(b) The flowers of plant Q are pollinated by insects.

(i) State **one** feature, visible on Fig. 5.1, that would attract insects to the flowers.

..... [1]

(ii) After pollination, fertilisation takes place in the flower.

Describe what happens during fertilisation.

.....
.....
..... [2]

(iii) Name the part of the flower that would develop into a fruit, following fertilisation.

..... [1]

(iv) What structures are present inside all fruits?

..... [1]

(c) Farmers often add fertilisers containing nitrates to the soil where they grow crops.

(i) Explain why this is done.

.....
.....
..... [2]

(ii) Explain why fields in which plant **Q** is growing will require less nitrate fertiliser than fields in which plant **P** is growing.

.....
.....
..... [2]

- 6 Nordic gold is an alloy of four metals used to make coins.



For
Examiner's
Use

Table 6.1 shows information about the metals contained in Nordic gold.

Table 6.1

metal	% by mass in Nordic gold
aluminium	5
copper	
tin	1
zinc	5

- (a) (i) Complete Table 6.1 by stating the percentage of copper in Nordic gold. [1]

- (ii) Suggest how Nordic gold could be made.

.....
..... [1]

- (iii) Nordic gold has properties which make it suitable for making coins.

Suggest **one** property Nordic gold is likely to have, other than its appearance, that makes it suitable for making coins.

Explain briefly why this property is important.

property

importance

..... [2]

- (b) (i) Tin may be extracted from tin oxide by heating a mixture of tin oxide and carbon. The other product of this reaction is carbon monoxide.

Write a **word** chemical equation for this reaction.

..... [1]

(ii) State and explain which substance is **oxidised** when tin is extracted from tin oxide.
 substance which is oxidised

explanation

..... [2]

(c) Aluminium is extracted from aluminium oxide, Al_2O_3 , by electrolysis. Aluminium oxide is an ionic compound.

(i) Explain the meanings of the following terms.
 cathode

electrolyte

..... [2]

(ii) Describe briefly the change in electronic structure which occurs when an aluminium atom becomes an aluminium **ion**.

..... [1]

(iii) Calculate the number of oxide ions which will be combined with 10 aluminium ions in aluminium oxide.

..... [1]

(d) Zinc oxide is a white solid used in sunscreen to protect human skin from the harmful effects of ultraviolet light from the sun.

(i) In a typical sunscreen, tiny pieces of zinc oxide are dispersed in a liquid.
 Name the type of mixture in which a solid is dispersed in a liquid.

..... [1]

(ii) Suggest why a sunscreen is able to protect human skin from ultraviolet light.

..... [1]

7 Most cells obtain energy from carbohydrates and other nutrients by aerobic respiration.

(a) Describe how a cell in a human muscle obtains the oxygen that it needs for respiration.

.....
.....
..... [2]

(b) When a person carries out exercise, muscle cells use energy to contract and produce movement.

(i) State **two** uses of energy in the human body, other than the contraction of muscles.

1
2 [2]

(ii) Some of the energy in exercising muscles is released as heat. Sweating helps to prevent the internal body temperature from rising too high.

State the correct biological term for the maintenance of a constant internal environment.

..... [1]

(iii) Suggest why an athlete running a long race, such as a marathon, needs to drink fluids during the run.

.....
.....
..... [2]

(c) Regular exercise can help to reduce the risk of having a heart attack.

Describe the events that lead to a heart attack.

.....
.....
.....
.....
..... [3]

Please turn over for Question 8.

- 8 Fig. 8.1 shows the inside of a refrigerator. The temperature inside the freezing compartment is $-20\text{ }^{\circ}\text{C}$ and the temperature in the rest of the refrigerator is $+5\text{ }^{\circ}\text{C}$.

For
Examiner's
Use

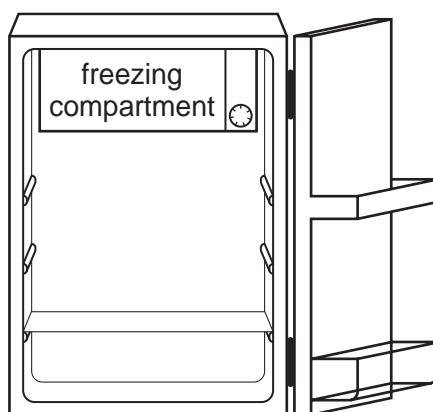


Fig. 8.1

- (a) (i) Draw arrows on Fig. 8.1 to show what happens to the air cooled by the freezing compartment. [1]
- (ii) Name this method of heat transfer. [1]
- (iii) Explain why this happens, using the idea of density.

.....
..... [1]

- (b) Fig. 8.2 shows an ice cube with sides of 2 cm. The ice cube has a mass of 7.4 g.

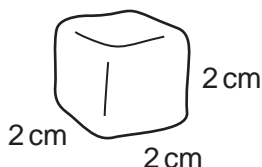


Fig. 8.2

- (i) Calculate the density of the ice.

State the formula that you use and show your working. State the units of your answer.

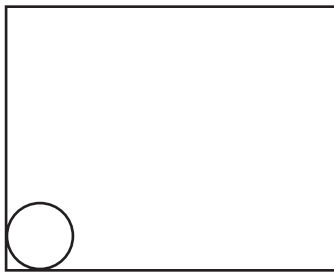
formula used

working

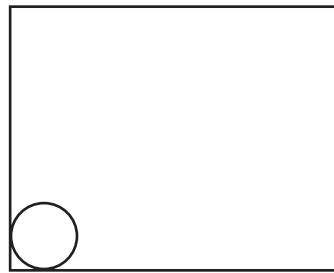
..... [3]

- (ii) The ice cube has been made by freezing some water.

Draw diagrams to show the arrangement of water molecules in solid ice and in liquid water. One molecule has been drawn for you in each box.



solid ice



liquid water

[2]

- (c) The refrigerator has a lamp inside. The supply voltage is 250 V and the current passing through the lamp when lit is 0.05 A.

- (i) Show that the resistance of the lamp when lit is 5000 Ω .

State the formula that you use and show your working.

formula used

working

..... [2]

- (ii) Two lamps with a resistance of 5000 Ω are connected together in series.

Calculate the combined resistance of the two lamps.

State the formula that you use and show your working.

formula used

working

..... Ω [2]

For
Examiner's
Use

9 Coral reefs are large rocky structures found in shallow seawater. The reefs are formed from the skeletons of small animals (coral polyps).

For
Examiner's
Use

(a) Seawater is a mixture which contains many dissolved compounds. The coral polyps extract the compound calcium carbonate from seawater and use it to build their skeletons.

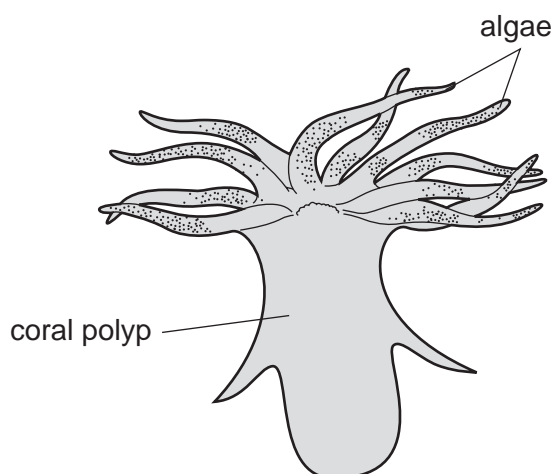
Choose **two** statements from the list below that describe compounds but which do **not** describe mixtures.

- A They have a chemical formula.
- B They can contain any amounts of two or more substances.
- C Their properties are different from those of the substances used to make them.
- D Their formation does **not** normally produce a significant change in temperature.

statements and [2]

(b) Certain algae (microscopic plants) live in the coral polyps, and these organisms help each other to survive.

The algae produce oxygen in the presence of sunlight. The coral polyps use oxygen and produce carbon dioxide as a waste product.



(i) Name the processes which are occurring

in the algae to produce oxygen,

.....

in the coral polyps to produce carbon dioxide.

..... [2]

(ii) Name the compound which is produced by the algae in addition to oxygen.

..... [1]

(iii) Suggest **one** way that the coral polyps and the algae help each other to survive.

.....
..... [1]

(c) In recent years, the amount of carbon dioxide in the atmosphere has increased. This has caused a decrease in the average pH of seawater.

During this period, many coral reefs have become damaged or have stopped growing.

(i) State and explain **one** example of human activity which has caused the amount of carbon dioxide in the atmosphere to increase in recent years.

.....
.....
..... [2]

(ii) Explain why increased levels of carbon dioxide in the atmosphere cause the average pH of seawater to decrease.

.....
.....
..... [2]

(iii) Suggest a possible reason why a decrease in the average pH of seawater could damage coral reefs.

.....
..... [1]

DATA SHEET
The Periodic Table of the Elements

		Group													
		I	II	III	IV	V	VI	VII	VIII	IX	X	0			
		1 H Hydrogen 1													
7 Li Lithium 3	9 Be Beryllium 4												4 He Helium 2		
23 Na Sodium 11	24 Mg Magnesium 12												20 Ne Neon 10		
39 K Potassium 19	40 Ca Calcium 20	51 V Vanadium 23	48 Ti Titanium 22	56 Fe Iron 26	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36	115 In Indium 49	122 Sb Antimony 51	127 I Iodine 53	131 Xe Xenon 54		
85 Rb Rubidium 37	88 Sr Strontium 38	91 Zr Zirconium 40	96 Mo Molybdenum 42	101 Ru Ruthenium 44	112 Cd Cadmium 48	122 Sb Antimony 51	128 Te Tellurium 52	131 Xe Xenon 54	146 Pb Lead 82	150 Hg Mercury 80	157 Gd Gadolinium 64	162 Dy Dysprosium 66	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71
133 Cs Caesium 55	137 Ba Barium 56	140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	162 Dy Dysprosium 66	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71	181 Fr Francium 87	226 Ra Radium 88	227 Ac Actinium 89	227 Ac Actinium 89
133 Cs Caesium 55	137 Ba Barium 56	140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	162 Dy Dysprosium 66	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71	181 Fr Francium 87	226 Ra Radium 88	227 Ac Actinium 89	227 Ac Actinium 89
<p>*58-71 Lanthanoid series †90-103 Actinoid series</p>															
		a		X		b									
		a = relative atomic mass		X = atomic symbol		b = proton (atomic) number									
		Key		X		b									

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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