

UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS International General Certificate of Secondary Education

	CANDIDATE NAME		
	CENTRE NUMBER		CANDIDATE NUMBER
* 2	COMBINED SC	IENCE	0653/23
9 2 2	Paper 2 (Core)		October/November 2012
2 5 4			1 hour 15 minutes
7	Candidates ans	wer on the Question Paper.	
3 0 9	No Additional M	aterials are required.	

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in. Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs, tables or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer all questions. A copy of the Periodic Table is printed on page 20.

At the end of the examination, fasten all your work securely together. The number of marks is given in brackets [] at the end of each question or part question.

For Exam	For Examiner's Use			
1				
2				
3				
4				
5				
6				
7				
8				
9				
Total				

This document consists of 20 printed pages.



UNIVERSITY of CAMBRIDGE International Examinations 1 Flowers are organs in which sexual reproduction takes place. For Examiner's Use (a) (i) Complete the definition of sexual reproduction. Use words from the list. dissimilar female haploid identical ovary sperm zygote Sexual reproduction is the process involving the fusion of nuclei to form a diploid and the production of genetically offspring. [3] (ii) State the scientific term for the fusion of two nuclei. [1] (b) Fig. 1.1 shows a section through a flower. В D Fig. 1.1 (i) Name the parts labelled A and B. Α [2] В _____ (ii) State the letter of the part in which the male gametes are produced, a zygote is produced. [2]

(c) After pollination, seeds are produced. A student set up an experiment to investigate the conditions needed for the germination of lettuce seeds.

3

He placed five lettuce seeds on cotton wool in each of five test-tubes. Fig. 1.2 shows the conditions present in each tube.

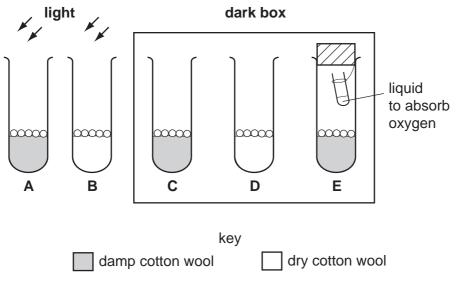




Table 1.1 shows his results.

Table	1.	1
-------	----	---

tube	conditions			number of seeds that germinated
Α	water	oxygen	light	5
В	no water	oxygen	light	0
С				5
D				0
E				0

- (i) Complete Table 1.1 to show the conditions present in each tube. Tubes A and B have been done for you. [2]
- (ii) What conclusions can the student make from these results?

[3]

For Examiner's Use

2	(a)	(i)	State the percentages of nitrogen and oxygen in the air.	For Examiner's
			nitrogen	Use
			oxygen [2]	
		(ii)	During a thunderstorm, energy from lightning causes nitrogen and oxygen to combine to form nitric oxide.	
			Explain why nitrogen is an example of an <i>element</i> and nitric oxide is an example of a <i>compound</i> .	
			[2]	
	(i	iii)	Nitric oxide has the chemical formula, NO.	
			Explain what is meant by this formula.	
			,	
			[2]	
	(i	iv)	What name is given to the type of chemical reaction that occurs when oxygen bonds to another element?	
			[1]	
			en magnesium burns in air, a white solid is formed. This white solid contains gnesium oxide, MgO.	
		(i)	Name the type of chemical bonding in magnesium oxide.	
			Explain your answer.	
			type of chemical bonding	
			explanation	

(ii) A student burned some magnesium in air and then added the white solid formed to water. Examiner's

She tested the solution with Universal (full range) Indicator and found that the pH was 9.

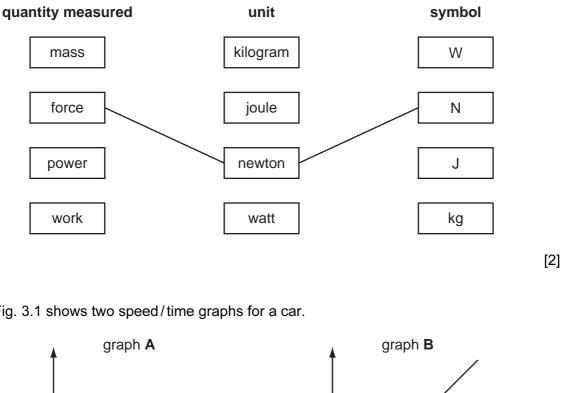
State a conclusion that the student can draw from this observation.

.....[1] For

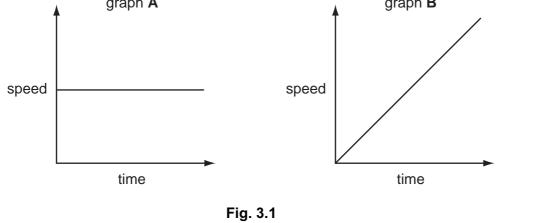
Use

3 (a) Draw lines to connect each quantity measured to its correct unit and symbol.

One has been done for you.



(b) Fig. 3.1 shows two speed/time graphs for a car.



Describe the motion of the car in graph A, [2] graph **B**.

For Examiner's Use

(c)	The	e car travels at 20 m/s for 90 seconds.		For
	Cal	culate the distance covered.		Examiner's Use
	Sta	te the formula that you use and show your working.		
		formula used		
		working		
		working		
		m	[2]	
(d)	On	e of the car's headlamps has a current of 2A, when the voltage across it is 12V.		
	(i)	Show that the resistance of the headlamp is 6Ω .		
		State the formula that you use and show your working.		
		formula used		
		working		
			[2]	
	(ii)	The car has two of these identical headlamps connected in series.		
		Calculate the total resistance of these two headlamps.		
		State the formula that you use and show your working.		
		formula used		
		working		
			101	
		Ω	[2]	

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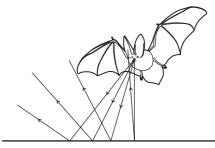
4	Many bats are predators that fly at night. They eat moths and other insects. (a) Underline the two words that describe the position of a bat in a food chain.					For Examiner's Use		
	C	arnivore	consumer	herbivore	producer		[1]	
	(b) Bats emit ultrasound.							
	(i) Ultrasound is sound that has a frequency too high for a human to hear.							
		Suggest a freque	ency for the ultraso	und emitted by ba	ts.	Hz	[1]	
	(ii) Underline the one word that correctly describes an ultrasound wave.							
		electroma	agnetic Ion	gitudinal	transverse		[1]	

(c) Bats use echo location to detect objects around them.

The reflected ultrasound waves are detected by special cells in the bat's head.

Fig. 4.1 shows how ultrasound waves are reflected from a rough surface and from a smooth surface. The arrows show the direction in which the sound waves travel.

rough surface



smooth surface

Fig. 4.1

(i) Use the information in Fig. 4.1 to describe what happens to the ultrasound waves when they hit

a rough surface,	
a smooth surface.	 •••••
	 [1]

8

(ii) Suggest how the bat can tell if it is flying over a rough surface or a smooth surface, even when it is completely dark.
[1]
(d) Many kinds of bat live in trees in forests.
List three ways in which deforestation can damage the environment.
1
2
3
[3]

0653/23/O/N/12

9

For

Examiner's Use 5 (a) In many countries, river water is collected and treated to make it safe for humans to For drink. Examiner's Use State and explain which two of the processes shown below are used to treat river water so that it becomes safe to drink. chlorination crystallisation filtration evaporation first process reason why this process is carried out second process reason why this process is carried out [4]

(b) Fig. 5.1 shows chromatography being used by a student to investigate mixtures of dyes (coloured compounds) used to colour sweets.

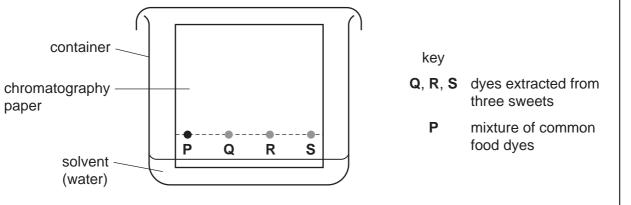
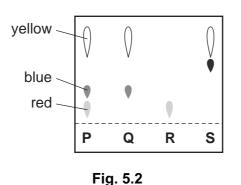




Fig. 5.2 shows the appearance of the chromatography paper after several minutes.



(i) Deduce and explain the colour of the sweet which contains only one dye.

.....[1]

11

For

Examiner's Use 6 (a) Fig. 6.1 shows a washing machine.





Complete the sentence below using $\ensuremath{\textit{two}}$ of the words in the list.

		heat	kinetic	light	potential	sound	
1	Αw	ashing machine	e is designed to	o transform el	ectrical energy into		
•			ene	rgy and		energy.	[2]
(b)	(i)	Some of the w	ater inside the	washing mac	hine evaporates.		
		Explain the pro	ocess of evapor	ration in term	s of particles.		
							[2]
(ii)	Explain why ev	vaporation has	a cooling effe	ct.		
							[1]

For Examiner's

Use

(c) The casing of the washing machine is a solid. The water used in it is a liquid.

Complete the diagrams below to show the arrangement of particles in a solid and in a liquid.

solid liquid

[2]

For Examiner's Use

(d) Before buying a washing machine, a person may research several types to find out which washing machine has the greatest energy efficiency.

Explain the meaning of the term efficiency.

[1]

7 (a) Fig. 7.1 shows two human teeth.

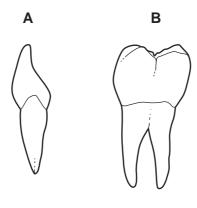


Fig. 7.1

(i) Name the **two** types of teeth shown in Fig. 7.1.

tooth A ______

(ii) Explain how tooth **B** helps to digest a food such as bread.

[2]

(b) For each part of the digestive system in the list below, tick (✓) the correct function or functions.

part	ingestion	digestion	absorption	
mouth				
stomach				
small intestine				

[3]

[2]

For Examiner's Use Please turn over for Question 8.

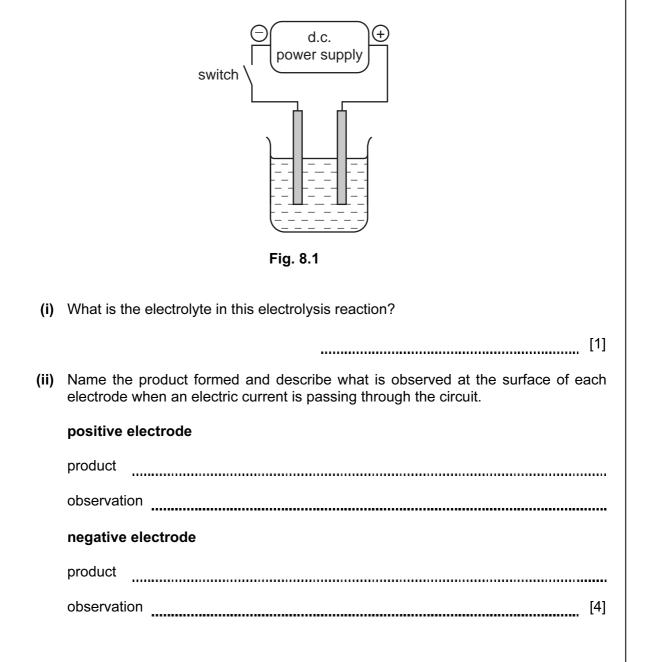
Metallic copper is a very important material that has been extracted from copper 8 For compounds for thousands of years. Examiner's Use (a) (i) The wires used in many electrical devices are made from copper. State the two properties of metals such as copper, that make them suitable for making electrical wires. 1 [2] 2 (ii) Copper wires are connected to the mains electrical supply using brass plugs. Brass is an alloy. copper wire brass plug Explain the meaning of the term alloy and state one difference in the physical properties of brass compared to copper. meaning of alloy difference in physical property [2] (iii) One of the processes used in the extraction of copper involves heating copper(I) sulfide in air. One of the reactions that occurs is between copper(I) sulfide and oxygen. This reaction also produces sulfur dioxide. Construct the **word** chemical equation for this reaction.[1]

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(b) Copper may also be formed by the electrolysis of an aqueous solution of copper chloride using electrodes made of graphite (carbon).

For Examiner's Use

Fig. 8.1 shows a laboratory apparatus a student used to carry out this electrolysis reaction.



9 (a) Complete Table 9.1 to show the circuit symbol for each of the named components.

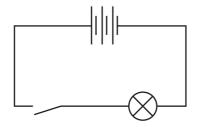
component	symbol
ammeter	
fuse	
variable resistor	

Table 9.1

For Examiner's Use

[3]

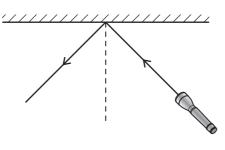
(b) Fig. 9.1 shows an electrical circuit for a torch (flashlight).





(i)	How many cells are fitted in the torch?	[1]
(ii)	A voltmeter is used to check the voltage across the light bulb.	
	Draw the symbol for the voltmeter in the correct position on the circuit.	[1]

(c) A single ray of light from a torch is shone onto a mirror as shown in Fig. 9.2.





- (i) On Fig. 9.2, label the angle of incidence and angle of reflection.
- (ii) The angle of incidence = 45° .

Write down the value of the angle of reflection.

For Examiner's Use

[1]

[1]

.....

	0	4 Helium 2	20 Neon Neon	40 Ar Argon	84 Krypton 36	131 Xe 54	Radon 86		175 Lu Lutetium 71	Lr Lawrencium 103
The Periodic Table of the Elements Group	١١		9 Fluorine 9	35.5 C1 Chlorine	80 Bromine 35	127 lodine 53	At Astatine 85		173 Yb ^{Ytterbium} 70	Nobelium 102
	N		16 Oxygen 8	32 Sultur 16	79 Selenium 34	128 Te Tellurium 52	Po Polonium 84		169 Tm ¹⁶⁹	Mendelevium 101
	>		14 Nitrogen 7	31 Phosphorus 15	75 AS Arsenic 33	122 Sb Antimony 51	209 Bi Bismuth		167 Er Erbium 68	Fm Fermium 100
	2		6 Carbon	28 Si licon 14	73 Ge Germanium 32	119 Sn 50	207 Pb Lead 82		165 Holmium 67	Es Einsteinium 99
	≡		ۍ Boron ± 1	27 A1 Aluminium 13	70 Ga Gallium 31	115 n Indium	204 T 1 Thallium 81		162 Dy Dysprosium 66	Californium 98
					65 Zn 30	112 Cd Cadmium 48	201 Hg ^{Mercury}		159 Tb 65	BK Berkelium 97
					64 Cu Copper	108 Ag Silver	197 Au Gold 79		157 Gd Gadolinium 64	Curium 96
					59 Nickel	106 Pd Palladium 46	195 Pt Platinum 78		152 Eu Europium	Americium 95
					59 CO 27 27	103 Rh odium 45	192 r ridium 77		150 Samarium 62	
		Hydrogen			56 Fe Iron	101 Ru Ruthenium 44	190 OS Osmium 76		Promethium 61	Neptunium 93
					55 Manganese 25	TC Technetium 43	186 Re Rhenium 75		144 Neodymium 60	238 U Uranium 92
					52 Cr Chromium 24	96 Mo Molybdenum 42	184 V Tungsten 74		141 Pr 59	Protactinium 91
					51 V Vanadium 23	93 Nb Niobium 41	181 Ta Tantalum 73		140 Ce Cerium 58	232 Th orium 90
					48 Trtanium 22	91 Zr Zirconium 40	178 Hafnium 72		_	nic mass bol nic) number
					45 Scandium 21	89 Y ttrium 39	139 La Lanthanum 57 *	227 Actinium 89 †	*58-71 Lanthanoid series 190-103 Actinoid series	a = relative atomic mass X = atomic symbol b = proton (atomic) number
				ε	.	etim			noid s bid	p × a
	=		9 Beryllium 4	24 Magnesium 12	40 Calcium 20	88 St rontium 38	137 Ba Barium 56	226 Ra Radium 88	*58-71 Lanthanoid serie 190-103 Actinoid series	<i>∝</i> ★

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20