



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
International General Certificate of Secondary Education

CANDIDATE  
NAME

CENTRE  
NUMBER

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**COMBINED SCIENCE**

**0653/21**

Paper 2 (Core)

**October/November 2012**

**1 hour 15 minutes**

Candidates answer on the Question Paper.

No Additional Materials are required.

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs, tables or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

**DO NOT WRITE IN ANY BARCODES.**

Answer **all** questions.

A copy of the Periodic Table is printed on page 20.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

For Examiner's Use	
1	
2	
3	
4	
5	
6	
7	
8	
9	
<b>Total</b>	

This document consists of **18** printed pages and **2** blank pages.



- 1 (a) Complete Table 1.1 by choosing one of the words from the list to match each statement.

ammeter      ampere      electron      insulator  
ohm      volt      voltmeter      watt

Table 1.1

statement	word
a particle with a negative electrical charge	
an instrument that measures electrical current	
the unit of potential difference	
a material that does not conduct electricity	

[4]

- (b) Fig. 1.1 shows two circuits, **A** and **B**. All the lamps and both cells are the same.

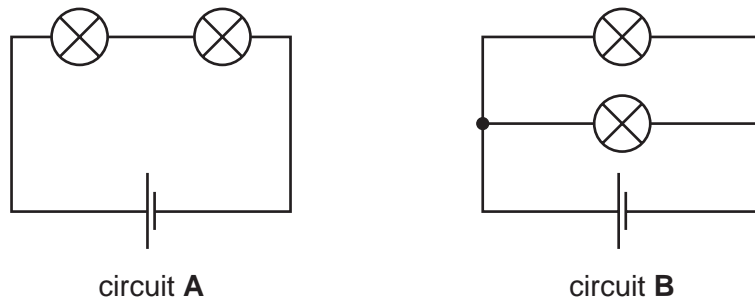


Fig. 1.1

- (i) One lamp is unscrewed from circuit **A**.

State what happens to the other lamp.

Explain your answer.

.....

.....

..... [2]

(ii) Explain why lights in a house are connected as in circuit **B** and **not** as in circuit **A**.

.....  
.....  
..... [2]

(iii) The resistance of each lamp is  $1.2\ \Omega$ .

Calculate the combined resistance of the two lamps in circuit **A**.

State the formula that you use and show your working.

formula used

working

.....  $\Omega$  [2]

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2 (a) Fig. 2.1 shows part of the carbon cycle.

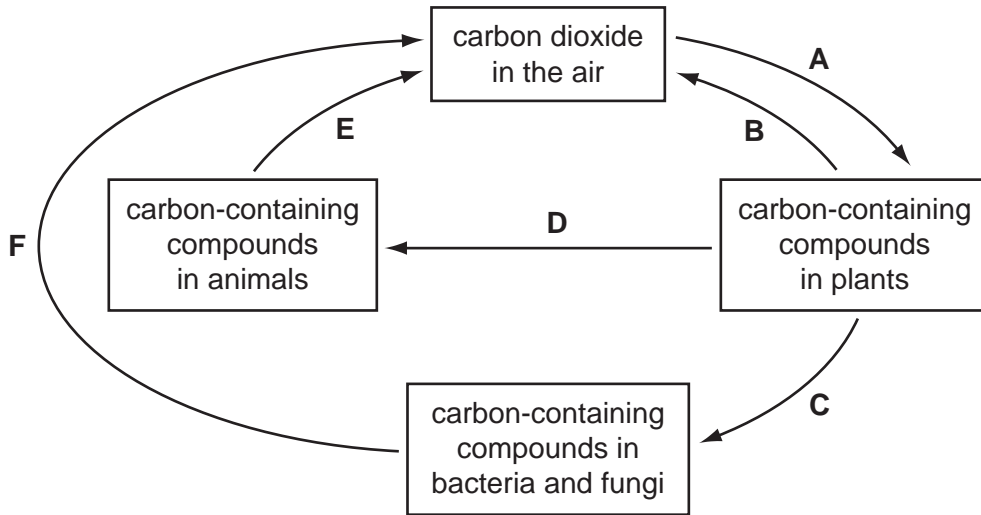


Fig. 2.1

(i) State the letter that represents photosynthesis in Fig. 2.1. .... [1]

(ii) State the **three** letters that represent respiration in Fig. 2.1.  
..... [1]

(iii) Name **one** carbon-containing compound in plants.  
..... [1]

(iv) State the approximate percentage of carbon dioxide in the air.  
..... [1]

(b) (i) Earthworms play an important part in the carbon cycle. They eat leaves, and egest material containing plant nutrients into the soil.

Explain the meaning of the term *egest*.

.....  
.....  
..... [2]

(ii) Underline the **two** words that describe the position of an earthworm in a food chain.

**carnivore      consumer      herbivore      producer**

[1]

(iii) Fishermen catch large numbers of earthworms to use as bait.

There are concerns that too many earthworms are being collected.

Suggest why it is important to conserve earthworms.

.....  
.....  
..... [2]

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- 3 (a) Fig. 3.1 shows how a digital pH meter is used to measure the pH of some liquids.

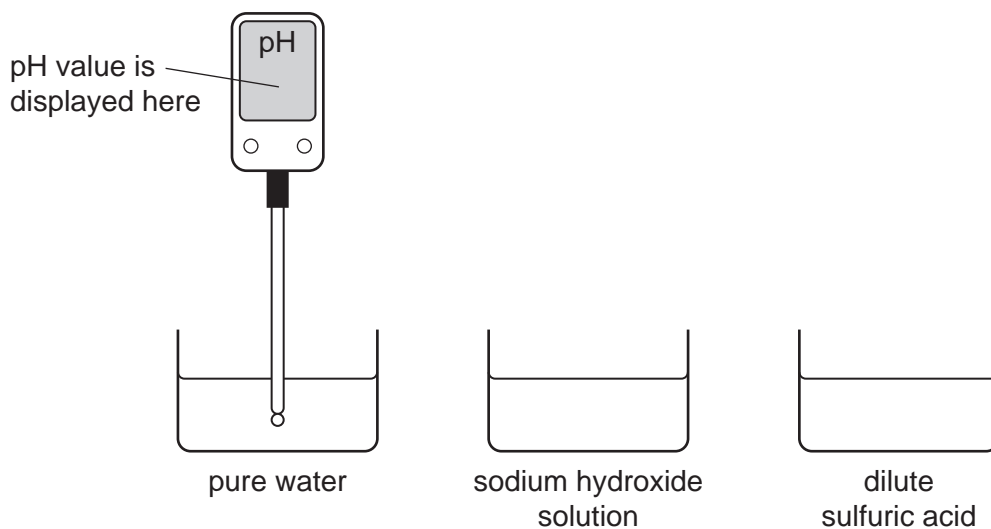


Fig. 3.1

- (i) Complete Table 3.1 by suggesting suitable pH values for the different liquids.

Table 3.1

liquid	pH
pure water	
sodium hydroxide solution	
dilute sulfuric acid	

[3]

- (ii) Suggest **one** advantage of using a digital pH meter rather than a piece of litmus paper to compare the acidity of two different acid solutions.

.....  
 .....  
 ..... [1]

- (b) Describe how a student could use a solution of acidified silver nitrate to find out whether or not an unlabelled solution contains sodium chloride.

.....  
 .....  
 ..... [2]

(c) When a reactive metal is added to a dilute acid, the metal reacts and dissolves and a gas is given off.

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(i) Name **one** reactive metal that must **not** be added to a dilute acid.

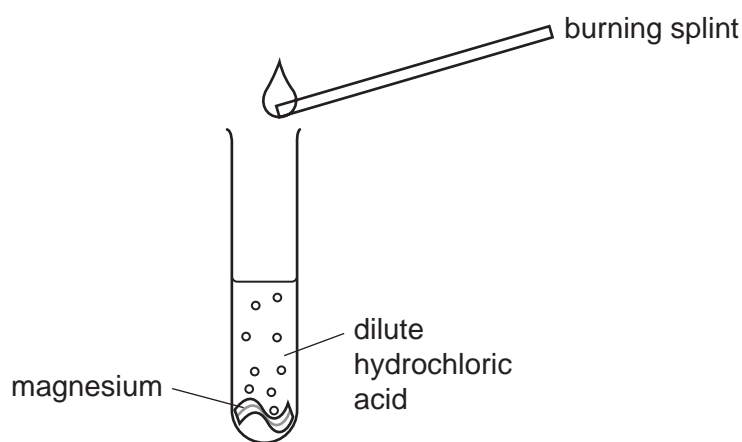
Explain why this metal should not be added to acid.

metal .....

explanation .....

..... [2]

(ii) Fig. 3.2 shows how a student tested the gas given off when magnesium was added to dilute hydrochloric acid.



**Fig. 3.2**

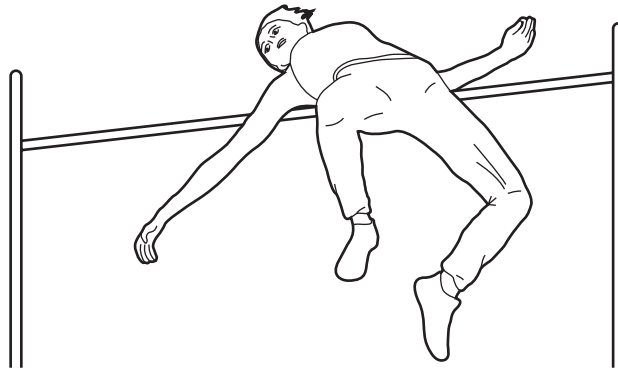
State and explain what the student observed when he carried out this test.

observation .....

explanation .....

[2]

4 An athlete competes in the high jump.



(a) Describe the energy changes that take place between the athlete taking off and landing after the high jump.

.....  
.....  
.....  
..... [3]

(b) After jumping, the athlete is sweating.

(i) Describe, in terms of particles, how evaporation occurs from the surface of a liquid.

.....  
.....  
..... [2]

(ii) Explain how this process will cool down the athlete.

.....  
.....  
..... [1]



**Please turn over for Question 5.**

5 Table 5.1 shows some of the nutrients contained in 100 g of five foods.

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Table 5.1

food	nutrients			
	sugar/g	starch/g	protein/g	fat/g
A	0	0	13	10
B	14	6	7	0
C	0	0	14	6
D	6	8	12	14
E	9	14	3	0

(a) (i) Which **two** nutrients listed in Table 5.1 are carbohydrates?

..... and ..... [2]

(ii) Which nutrient listed in Table 5.1 contains nitrogen atoms in its molecules?

..... [1]

(iii) State the letters of **two** foods in Table 5.1 that could have come from animals.

..... and ..... [1]

(iv) State the letter of **one** food that would appear orange-brown when tested with iodine solution, and give a purple colour when tested with biuret reagent.

..... [1]

(v) State the letter of the food that provides the most energy per 100g.

..... [1]

(b) Table 5.1 does **not** contain information about vitamins or minerals.

Outline the symptoms that a person may develop if their diet is deficient in

(i) vitamin D, .....

..... [1]

(ii) iron. ....

..... [1]

(c) Explain why eating a lot of foods containing sugar can increase the risk of tooth decay.

.....

.....

.....

.....

..... [3]

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- 6 Some types of firework are made by filling a cardboard tube with firework mixture. Firework mixture is made from several solid substances which have been powdered and mixed together.

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Fig. 6.1 shows a typical firework.

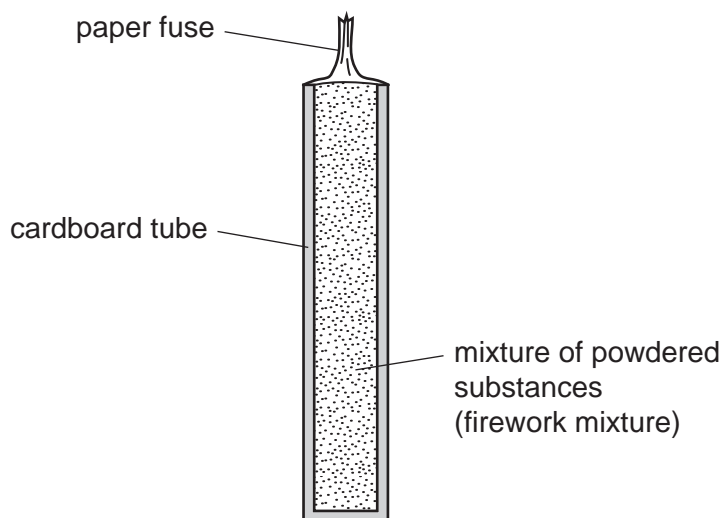


Fig. 6.1

When the paper fuse is lit, exothermic chemical reactions occur inside the firework.

- (a) (i) State **two** forms of energy that are released when the firework mixture reacts.

1 .....

2 .....

[2]

- (ii) State the effect on the rate of reaction of using firework mixture in the form of a powder.

..... [1]

(b) Some firework mixtures contain aluminium which is oxidised when the firework is lit.

Table 6.1 shows the numbers of protons and electrons in four particles, **A**, **B**, **C** and **D**, which are involved in the oxidation of aluminium.

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**Table 6.1**

particle	number of protons	number of electrons
<b>A</b>	8	10
<b>B</b>	13	13
<b>C</b>	8	8
<b>D</b>	13	10

(i) Atoms of the element aluminium have the proton number 13.

State and explain which particle, **B** or **D**, in Table 6.1 is an **atom** of aluminium.

particle .....

explanation .....

..... [1]

(ii) State and explain which **two** particles in Table 6.1 could be found bonded together in aluminium oxide.

particles ..... and .....

explanation .....

.....

..... [3]

(c) Firework mixtures contain the compound potassium perchlorate,  $KClO_4$ .

When potassium perchlorate is heated, a colourless gas is given off which re-lights a glowing splint.

(i) State the name of this gas. .... [1]

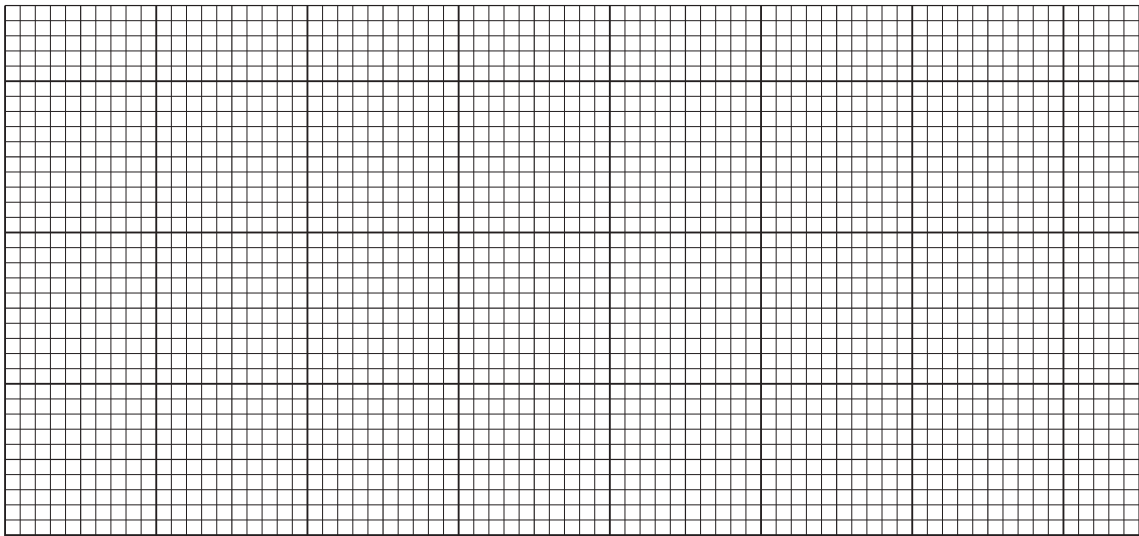
(ii) Suggest how potassium perchlorate in the firework mixture helps the mixture to burn.

.....

.....

..... [2]

- 7 (a) On the grid below, draw a wave with an amplitude of 2 cm and a wavelength of 4 cm.  
On your diagram, clearly label the amplitude and the wavelength.



[3]

- (b) (i) Two sound waves, **A** and **B**, have the same frequency. **A** has a greater amplitude than **B**.

What difference would you hear?

..... [1]

- (ii) Two sound waves, **X** and **Y**, have the same amplitude. **X** has a greater frequency than **Y**.

What difference would you hear?

..... [1]

- (c) Energy travels to the Earth from the Sun.

State whether this transfer of energy is by conduction, convection or radiation.

Explain your answer.

.....  
.....  
..... [2]

8 Fig. 8.1 shows the male reproductive system.

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Examiner's  
Use

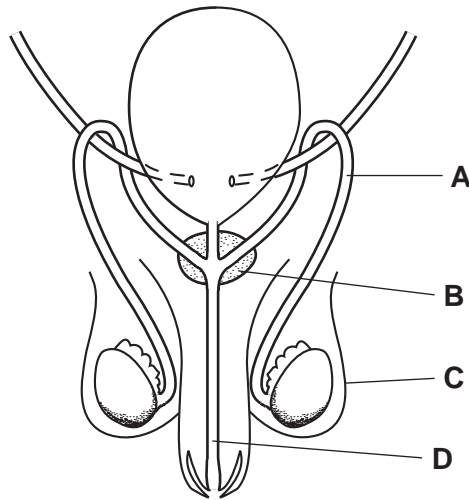


Fig. 8.1

(a) (i) Name parts C and D.

C .....

D .....

[2]

(ii) State the functions of parts A and B.

A .....

B .....

[2]

(iii) On Fig. 8.1, use a label line and the letter S to indicate where male gametes are made. [1]

(b) The human immunodeficiency virus (HIV) can be transmitted during sexual intercourse.

Outline **two** other ways in which HIV can be transmitted.

1 .....

.....

2 .....

.....

[2]

9 Chlorine is released when hydrochloric acid reacts with the compound, manganese dioxide.

(a) (i) Explain why chlorine is an example of an *element* and **not** a *compound*.

.....

.....

.....

..... [2]

(ii) Describe a safe test for chlorine gas.

.....

..... [2]

(b) Chlorine is produced in the chemical industry by electrolysis.

A simplified diagram of the apparatus used to produce chlorine is shown in Fig. 9.1.

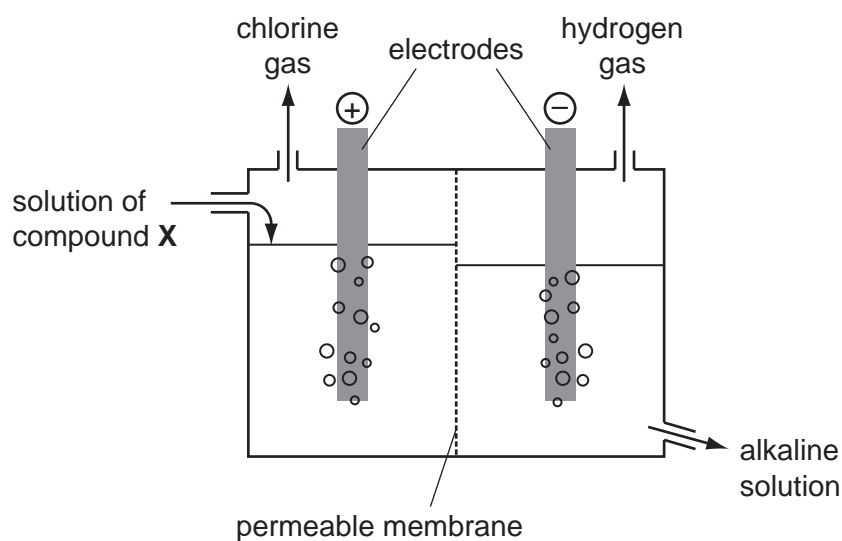


Fig. 9.1

(i) State the meaning of the term *anode*.

.....

..... [1]



- (ii) A student knows that compound **X** in Fig. 9.1 is either sodium hydroxide, NaOH, or sodium chloride, NaCl.

Using information from Fig. 9.1, deduce whether compound **X** is sodium hydroxide or sodium chloride.

Explain your answer.

**X** is .....

explanation .....

..... [1]

- (c) Chlorine is found in Group 7 of the Periodic Table. Two of the other elements in Group 7 are bromine and iodine.

- (i) Chlorine is a gas at room temperature.

What are the physical states of bromine and iodine at room temperature?

bromine .....

iodine ..... [2]

- (ii) Explain briefly why a solution of sodium bromide turns orange when chlorine is bubbled through it.

.....

..... [2]





