UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS

International General Certificate of Secondary Education

MARK SCHEME for the October/November 2010 question paper for the guidance of teachers

0653 COMBINED SCIENCE

0653/32

Paper 3 (Extended Theory), maximum raw mark 80

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

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Page 2		2	Mark Schei				Syllabus	S	Paper
			IGCSE – Od	ctober/No	vember 20	010	0653		32
1	(a) ide	a of re	estoring full/correct	number (of 46) in the	e zygote ;			[1]
	(b) (i)	ovar	y ;						[1]
	(ii)	ovid	uct / Fallopian tube ;						[1]
			s / contains, amniotic / supports, embryo ;						[2]
	(d) (i)	T, be	ecause Tt does not	have thala	assaemia/	owtte;			[1]
	(ii)	pher	notypes of parents		man with thalassae			n withou ssaemia	
		gend	otypes of parents		Tt			Tt	
		gam	etes	(T and	t	(T)	and (t	
						gametes f	rom woman		
							t		
				gametes	T	тт	Tt		
				from mar	t	Tt	tt thalassaemia		
		gam offsp	ental genotype ; ete genotypes ; oring genotypes ; I with thalassaemia	identified	•				[4]
	(iii)		moglobin transports lood) ;	oxygen/	person with	h thalassa	aemia has les	s oxyge	n
	so less respiration (in cells)/description of respiration;				[2 max]				
									[Total: 12]

Page 3	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – October/November 2010	0653	32

- 2 (a) (i) pink/orange/brown/copper (layer);
 - (ii) 2+;

two negative charges from chloride must balance the charge on the copper ion / owtte :

[2]

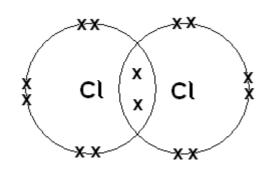
[1]

(iii) (L)

it is a negative ion/has a negative charge/has more electrons than protons; reference to attraction between opposite charges; (points separately marked)

[2]

(iv)



one shared pair; all other electrons correctly shown;

[2]

[2]

(b) (i) carbon dioxide;

[1]

(ii) 2PbO + C → 2Pb + CO₂;;(correct formulae and balanced)

[Total: 10]

3 (a) (i)

	description	charge	range in air	ionising ability
alpha	helium nucleus	positive	5 cm	very strong
beta	electron	negative	50 cm	medium
gamma	electromagnetic wave	none	many kilometres	weak

(the wording for ionising ability **must** show beta lies between alpha and gamma) ;;;;

[4]

(ii) alpha particles have low penetration in air/absorbed by casing/will not reach people living in house/smoke detectors are a long way from people;

[1]

(b) working (on graph or numerically); 5 hours;

[2]

[Total: 7]

	Page 4			Mark Scheme: Teachers' version	Syllabus	Paper
				IGCSE – October/November 2010 0653		32
4	(a)	bund plou	ds/e gh a cro	/building of walls (qualified); embankments/ditches; long slope (not up and down); p cover; es;		[max 2]
	(b)		kills does ecos	antage more pests / can completely destroy pest population s not introduce a (potentially) damaging new system);	_	the
		disadvantage may kill other beneficial/all insects/toxic to humans/have to apply severa times/development of resistance; bioaccumulation/persistence provided related directly to DDT;				veral
				ore refs to costs unless related to reason) lax for advantage, 1 max for disadvantage)		[2]
		` ,	absc	ning orbed (by plant) and transported (in phloem) ; thes all parts of plant ;		
			can	antage kill pests even if it does not directly hit them; affects insects feeding on the plant;		[2]
			(1 m	ax for advantage, 1 max for disadvantage)		
						[Total: 6]
5	(a)	(i)				[1]
				hts up/on ; nd L go off ;		[2]
	(b)			resistors;		
		in pa		el; on to show this;		[3]
	(c)	(i)	coil (cuts magnetic field / coil experiences changing mag	netic field ;	[1]
				ction of magnetic field relative to coil changes (eve otion of coil through magnetic field changes / revers	•	ction [1]
						[Total: 8]

Page 5	Mark Scheme: Teachers' version	Syllabus	Paper
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- 6 (a) $(H^+ +) OH^- \rightarrow H_2O ;;$ [2]
 - (b) (acid is added) until indicator / solution changes colour; colour change correct allow blue to either red or reasonable intermediate; [2]
 - (c) no indicator added/use of pH meter to show neutrality; ref. to same amount/volume of sodium hydroxide solution/alkali (as in (b)); ref. to same amount/volume of acid (as in (b)); evaporate/heat/boil off the water (from the solution);

[Total: 7]

[max 3]

- 7 (a) (i) traps layer of air; acts as insulator / reduces convection and conduction; [2]
 - (ii) white surfaces <u>radiate</u> less heat than black surfaces; less heat is lost; [2]
 - (b) (i) below 20 Hz; lowest frequency of human hearing is 20 Hz / below range of human hearing; [1]
 - (ii) (number of) waves / oscillations produced per unit time / wavelengths passing a point per unit time; [1]
 - (iii) waves have same amplitude; less waves shown on trace; [2]
 - (c) (i) 1.6 cm; [1]
 - (ii) both rays drawn backwards to meet; image labelled / clearly and unambiguously visible on diagram; [2]
 - (iii) image which cannot be projected (onto a screen)/light (rays) does not pass through it; [1]

[Total: 12]

Page 6	Mark Scheme: Teachers' version	Syllabus	Paper
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8 (a)

(2C and 4H bonded and double bond shown)

[2]

(b) (catalytic/thermal) cracking; fractions are boiled/vaporised/heated; passed over (hot) catalyst/subjected to very high temp. and pressure; (allow named catalyst e.g. alumina, silica, pumice, porcelain)

[3]

(c) double bonds become single; single bonds form between molecules to form a long chain; (marks can be obtained by clear diagrams)

[2]

[2]

(d)
$$A_r C = 12$$
 and $H = 1$;
 $(12 \times 2) + (1 \times 4) = 28$;

[Total: 9]

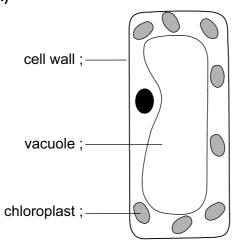
9 (a) water <u>vapour</u> lost from plant('s leaves); correct ref. to transpiration; condensation;

water vapour cooled; gas changed to liquid;

ref. to particles and (kinetic) energy;

[max 4]

(b) (i)



[max 2]

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(ii) water moved out of the cell;

down a water potential gradient/from where there was a lot of water to where there was less/from dilute solution to concentrated solution; through partially permeable cell membrane; so volume of cell/vacuole shrank; strong cell wall cannot change shape (much) so cytoplasm/cell membrane pulls away from it;

[max 3]

[Total: 9]