

COMBINED SCIENCE

Paper 1 Multiple Choice

0653/13 May/June 2012 45 minutes

Additional Materials:	Multiple Choice Answer Sheet
	Soft clean eraser
	Soft pencil (type B or HB is recommended)

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

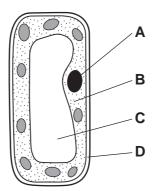
Each correct answer will score one mark. A mark will not be deducted for a wrong answer. Any rough working should be done in this booklet. A copy of the Periodic Table is printed on page 20.

This document consists of 18 printed pages and 2 blank pages.



1 The diagram shows a cell from the mesophyll of a leaf.

Which part contains DNA?

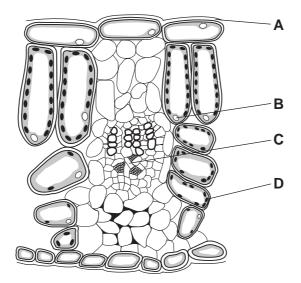


2 Which of these structures does an animal cell have?

	cell surface membrane	cell wall
Α	\checkmark	\checkmark
в	\checkmark	x
С	x	\checkmark
D	×	×

3 The diagram shows part of a cross-section of a leaf.

Where is light energy used?

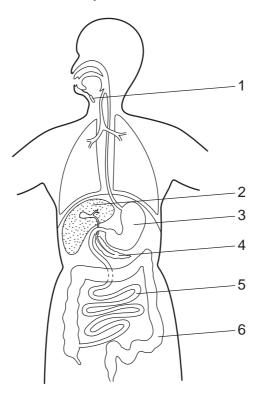


- 4 The statements are about enzymes.
 - 1 act as catalysts
 - 2 can be denatured by heat
 - 3 composed of complex carbohydrates
 - 4 not affected by pH
 - 5 produced by cells

Which statements are correct?

Α	1, 2 and 5	В	1, 4 and 5	С	2, 3 and 4	D	3 and 5 only
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5 The diagram shows the human alimentary canal.

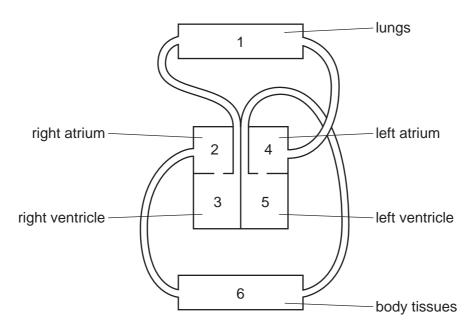


Which numbered structures secrete chemicals that are involved in fat digestion?

A 1 and 3 **B** 2 and 4 **C** 3 and 5 **D** 5 and 6

- 6 In which organisms does respiration take place?
 - A animals and plants, in both the dark and light
 - B animals in both the dark and light, but never in plants
 - **C** animals in the dark and light, plants only in the dark
 - D animals in the dark and light, plants only in the light

7 Which sequence of numbers represents the path taken by a deoxygenated red blood cell?

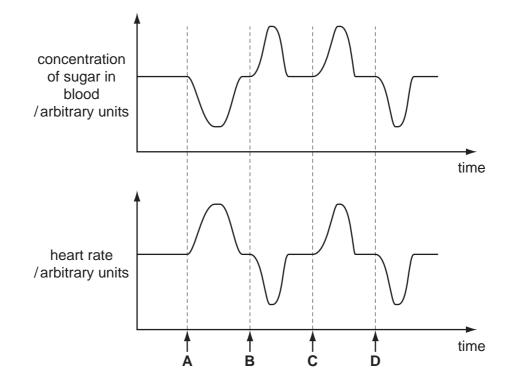


- $\textbf{A} \quad 1 \rightarrow 3 \rightarrow 2 \rightarrow 6$
- $\textbf{B} \quad 1 \rightarrow 4 \rightarrow 5 \rightarrow 6$
- $\textbf{C} \quad 6 \rightarrow 2 \rightarrow 3 \rightarrow 1$
- $\textbf{D} \quad 6 \rightarrow 5 \rightarrow 4 \rightarrow 1$
- 8 In which state does water enter and which state does water leave a plant?

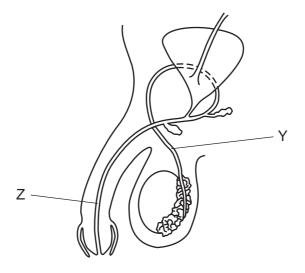
	enters	leaves
Α	liquid	liquid
в	liquid	vapour
С	vapour	liquid
D	vapour	vapour

9 The graphs show changes in the rate of heartbeat and in the concentration of sugar in the blood over the same period of time.

When was adrenaline secreted?



10 The diagram shows the male reproductive system.

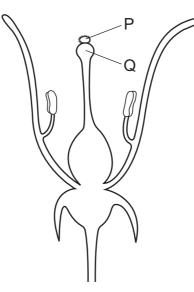


What are Y and Z?

	Y	Z
Α	prostate gland	urethra
в	urethra	prostate gland
С	sperm duct	prostate gland
D	sperm duct	urethra

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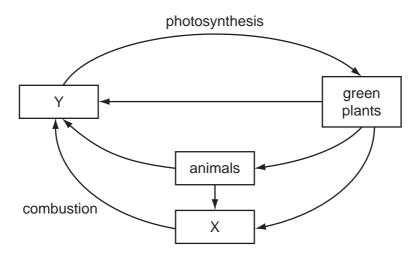
11 The diagram shows a section through a flower immediately after pollination.



Which row identifies structures P and Q?

	Р	Q
Α	pollen grain	anther
в	pollen grain	stigma
С	zygote	anther
D	zygote	stigma

12 The diagram shows part of the carbon cycle.



What are X and Y?

	Х	Y
Α	carbon dioxide	oxygen
в	fossil fuel	carbon dioxide
С	fossil fuel	oxygen
D	oxygen	carbon dioxide

13 Which are possible harmful effects of deforestation?

	global warming	species extinction
Α	\checkmark	✓
в	\checkmark	×
С	×	✓
D	×	x

14 The melting points of three substances are given in the table.

	melting point/°C
W	-219
X	0
Z	3500

What could the substances be?

	W	Х	Z
Α	oxygen	graphite	water
В	oxygen	water	graphite
С	graphite	water	oxygen
D	water	graphite	oxygen

- 15 Which statement correctly describes the bonding in compounds?
 - A An ionic compound contains two metallic elements.
 - **B** In a compound formed between a metal and a non-metal, the metal has a negative charge.
 - **C** When a metal combines with a non-metal, electrons are shared between the elements.
 - **D** When two non-metals combine, molecules are formed.
- **16** Four minerals containing calcium are listed.

mineral	chemical formula
calcite	CaCO ₃
dolomite	CaCO ₃ .MgCO ₃
fluorite	CaF₂
gypsum	CaSO ₄

What is the total number of different elements contained in these minerals?

A 4 **B** 5 **C** 6 **D** 7

17 Copper(II) oxide reacts as shown.

 $CuO \ + \ H_2 \ \rightarrow \ Cu \ + \ H_2O$

Which statement about the reaction is correct?

- **A** Both oxidation and reduction take place.
- **B** Copper(II) oxide is a reducing agent.
- **C** Copper(II) oxide is oxidised to copper.
- D Hydrogen is reduced to water.
- 18 Which change does not alter the rate of reaction between zinc and dilute sulfuric acid?
 - **A** addition of a catalyst
 - **B** change in concentration of the acid
 - **C** change in atmospheric pressure
 - D change in temperature
- **19** Barium oxide is a metallic oxide.

A solution of barium oxide is1..... with a pH2..... than seven.

Which words correctly complete gaps 1 and 2?

	1	2
Α	acidic	greater
в	acidic	less
С	alkaline	greater
D	alkaline	less

20 Acid X reacts with metal Y.

A colourless gas is given off and a pale green solution is produced.

Two tests are carried out on the solution.

test	reagent(s) added	result
1	aqueous silver nitrate and nitric acid	white precipitate
2	aqueous sodium hydroxide	green precipitate

What are acid X and metal Y?

	acid	metal
Α	hydrochloric	iron
В	hydrochloric	zinc
С	sulfuric	iron
D	sulfuric	zinc

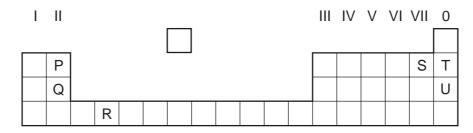
21 The properties of the elements in Groups I and VII of the Periodic Table change down the group.

Which row shows how the melting points change down these groups?

	Group I Group VI	
Α	decrease	decrease
в	decrease	increase
С	increase	decrease
D	increase	increase

22 The diagram shows elements in part of the Periodic Table.

The letters are **not** the actual symbols of the elements.

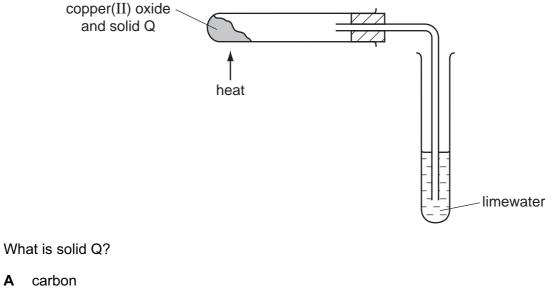


Which statement about the elements shown is correct?

- **A** P and S are in the same group of the Periodic Table.
- **B** P is more reactive than Q.
- **C** R is likely to form coloured compounds.
- **D** T and U are halogens.
- 23 Which property of a metal determines the method used to extract the metal from its ore?
 - **A** the melting point of the metal
 - **B** the position of the metal in the Periodic Table
 - **C** the reactivity of the metal
 - **D** the relative atomic mass, A_r , of the metal

24 Copper(II) oxide is mixed with a solid Q.

On heating the mixture, a reaction occurs and the limewater turns milky.



- **B** iron
- C sulfur
- D zinc
- **25** The diagrams show molecules of four gases present in clean air. Different circles represent atoms of different elements.







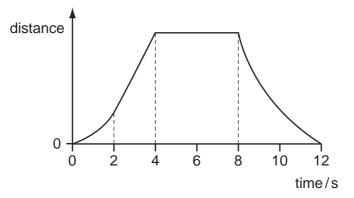
Which elements are shown as \bullet and \bigcirc ?

		0			
Α	hydrogen	nitrogen			
В	hydrogen oxygen				
С	oxygen	hydrogen			
D	oxygen	nitrogen			

26 Which property of the compounds in petroleum is used to separate it into useful fractions?

A boiling point

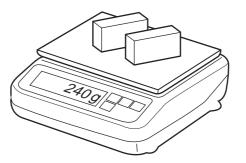
- B density
- C melting point
- **D** solubility
- 27 Which equation shows the complete combustion of a hydrocarbon?
 - $\label{eq:constraint} \begin{array}{cccc} \mbox{\bf A} & C_2 H_4 \ + \ 2 O_2 & \rightarrow \ 2 CO \ + \ 2 H_2 O \end{array}$
 - $\textbf{B} \quad C_2H_4 \ \textbf{+} \ 3O_2 \quad \rightarrow \ 2CO_2 \ \textbf{+} \ 2H_2O$
 - $\label{eq:constraint} \mbox{C} \quad C_2 H_6 O \mbox{ + } 2O_2 \mbox{ \rightarrow } 2CO \mbox{ + } 3H_2 O$
 - $\textbf{D} \quad C_2H_6O \ \textbf{+} \ \ 3O_2 \ \rightarrow \ 2CO_2 \ \textbf{+} \ \ 3H_2O$
- 28 The graph shows how the distance of an object changes with time.



Between which two times is the object moving with a non-zero constant speed?

- A between 0s and 2s
- B between 2s and 4s
- C between 4s and 8s
- D between 8s and 12s

29 A shop-keeper places two identical blocks of cheese on a set of scales and notices that their combined mass is 240g. Each block measures $2.0 \text{ cm} \times 5.0 \text{ cm} \times 10.0 \text{ cm}$.



What is the density of the cheese?

A 0.42 g/cm^3 **B** 0.83 g/cm^3 **C** 1.2 g/cm^3 **D** 2.4 g/cm^3

30 Which row shows the unit of energy, the unit of power and the unit of work?

	unit of energy	unit of power	unit of work
Α	joule	joule watt	
в	joule	watt	joule
С	watt	joule	watt
D	watt	watt	joule

31 Warm water in a bowl evaporates.

Which row shows where the evaporation occurs and what effect the evaporation has on the temperature of the remaining water?

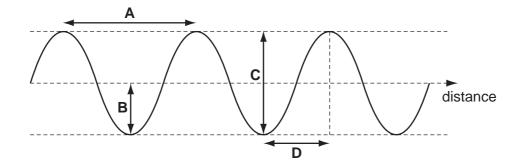
	where evaporation occurs	effect on water temperature
Α	only on the surface	decreases
В	only on the surface	no change
С	throughout the water	decreases
D	throughout the water	no change

32 In which state(s) of matter can convection occur?

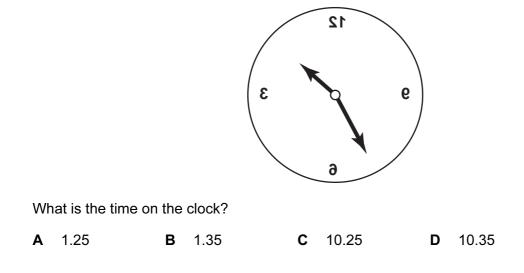
	in a solid	in a liquid	in a gas yes	
Α	no	no	yes	
В	no	yes	yes	
С	yes	no	no	
D	yes	yes	no	

33 The diagram represents a wave.

Which arrow represents the amplitude of the wave?

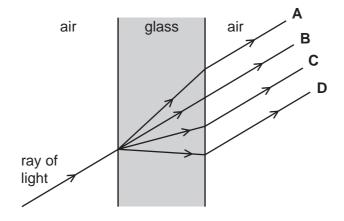


34 The image of a clock face as seen in a plane mirror is shown.

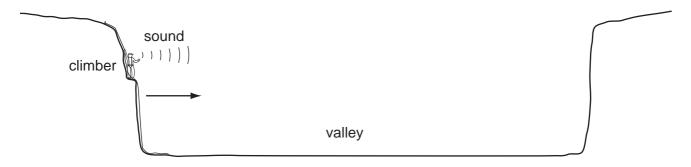


35 A ray of light passes through a glass window.

Which path does it take?



- 36 Which statement about the electromagnetic spectrum is correct?
 - **A** Gamma rays have the highest frequency.
 - **B** Microwaves have the smallest wavelength.
 - C Ultraviolet waves have the largest wavelength.
 - **D** Visible light waves have the lowest frequency.
- **37** A climber stands on a ledge on one side of a valley. He claps his hands and hears an echo from the opposite side of the valley 1.6 s later.

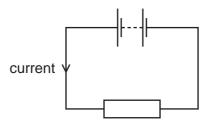


The speed of sound is 330 m/s.

How wide is the valley, to the nearest metre?

Α	103 m	В	206 m	C 264 m	D	528 m
A	10311	D	200111	C 204111	U	52

38 A battery is connected to a resistor.

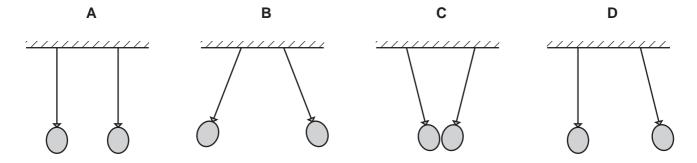


Which changes to potential difference across the resistor, and to the resistance of the resistor, will produce the smallest current?

	potential difference	resistance
Α	decrease	decrease
в	decrease	increase
С	increase	decrease
D	increase	increase

39 Two balloons hang near each other on insulating threads. One balloon is positively charged and the other is negatively charged.

Which diagram shows how they hang?



40 How are lamps usually connected in a lighting circuit, and what is one advantage of this?

	how lamps are connected	advantage
Α	in parallel	they each receive the full supply voltage
в	in parallel	they share the supply voltage equally
С	in series	they each receive the full supply voltage
D	in series	they share the supply voltage equally

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18

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19

	0	4 Helium 2	20 Neon 40	Ar Argon 18	84 Krypton 36	131 Xe 54	Radon 86		175 Lu Lutetium 71	Lr Lawrencium 103
	١١		9 35.5	C1 Chlorine	80 Bromine 35	127 lodine 53	At Astatine 85		173 Yb ^{Ytterbium} 70	Nobelium 102
	N		16 8 Oxygen 32	Sulfur 16	79 Se Selenium 34	128 Te Tellurium	Po Polonium 84		169 Tm ¹⁶⁹	Mendelevium 101
	>		14 7 Nitrogen 31	Phosphorus 15	75 AS Arsenic 33	122 Sb Antimony 51	209 Bi Bismuth		167 Er Erbium 68	Fermium 100
	\geq		6 Carbon 6 28	Silicon 14	73 Ge Germanium 32	119 Sn 50	207 Pb Lead 82		165 Holmium 67	Ensteinium 99
	=		11 5 ^{Boron} 27	A1 Auminium 13	70 Ga Gallium 31	115 n Indium	204 T1 Thalium 81		162 Dy Dysprosium 66	Cf Californium 98
21115					65 Zn 30	112 Cd Cadmium 48	201 Hg ^{Mercury}		159 Tb 65	BK Berkelium 97
Group					64 Cu Copper	108 Ag Silver	197 Au Gold 79		157 Gd Gadolinium 64	ocurium 96
Group					59 Nickel 28	106 Pd Paladium 46	195 Pt Platinum 78		152 Eu Europium	Am Americium 95
					59 Co ^{Cobalt}	103 Rh odium 45	192 r ridium		150 Samarium 62	
		¹ Hydrogen			56 Fe Iron 26	101 Ru Ruthenium 44	190 OS Osmium 76		Promethium 61	Neptunium 93
					55 Mn ^{Manganese} 25	TC Technetium 43	186 Re Rhenium 75		144 Neodymium 60	238 U Uranium 92
					52 Cr Chromium 24	96 Mo Molybdenum 42	184 V Tungsten 74		141 Pr 59	Pa Protactinium 91
					51 V Vanadium 23	93 Nb Niobium 41	181 Ta Tantalum 73		140 Ce Cerium 58	232 Th orium 90
					48 Tranium 22	91 Zr Zirconium 40	178 Hafnium 72		_	nic mass bol nic) number
					45 Scandium 21	89 Yttrium 39	139 La Lanthanum 57 *	227 Actinium 89 †	*58-71 Lanthanoid series 190-103 Actinoid series	a = relative atomic mass X = atomic symbol b = proton (atomic) number
				5	~ E	m L			noid s bid	p × a
	=		9 Berylium 24	Mg Magnesium 12	40 Calcium 20	88 St rontium 38	137 Ba Barium 56	226 Radium 88	*58-71 Lanthanoid serie 190-103 Actinoid series	т Х

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