



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

COMBINED SCIENCE

0653/11

Paper 1 Multiple Choice

May/June 2012

45 minutes

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

* 4 7 7 8 3 1 2 2 0 1 *

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

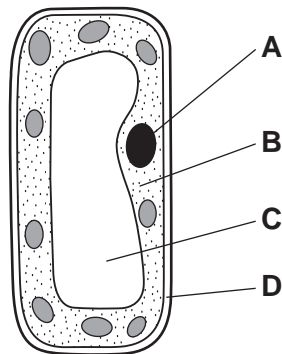
A copy of the Periodic Table is printed on page 20.

This document consists of **17** printed pages and **3** blank pages.

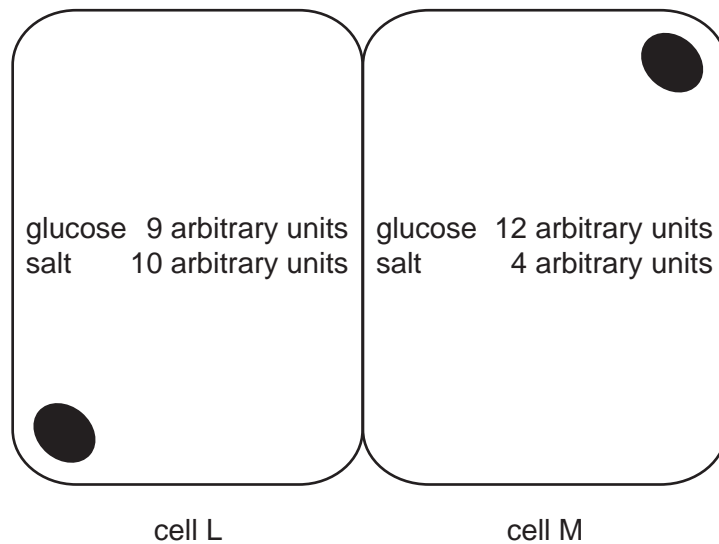


- 1 The diagram shows a cell from the mesophyll of a leaf.

Which part contains DNA?



- 2 The diagram shows two cells in contact with one another, and the concentrations of glucose and salt in each of the cells.



Which movements would occur by diffusion?

- A glucose from cell L to cell M, salt from M to L
- B glucose and salt from L to M
- C glucose from cell M to cell L, salt from L to M
- D glucose and salt from M to L

3 The statements are about enzymes.

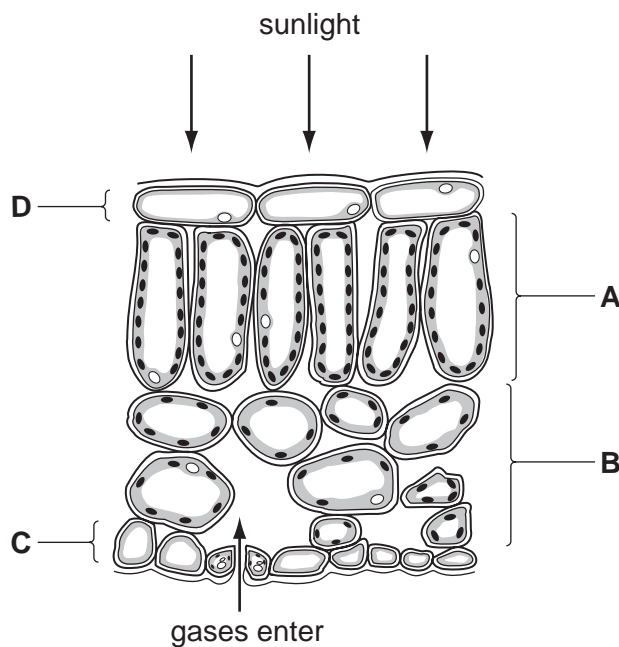
- 1 act as catalysts
- 2 can be denatured by heat
- 3 composed of complex carbohydrates
- 4 not affected by pH
- 5 produced by cells

Which statements are correct?

- A** 1, 2 and 5 **B** 1, 4 and 5 **C** 2, 3 and 4 **D** 3 and 5 only

4 The diagram shows some cells in a leaf of a green plant.

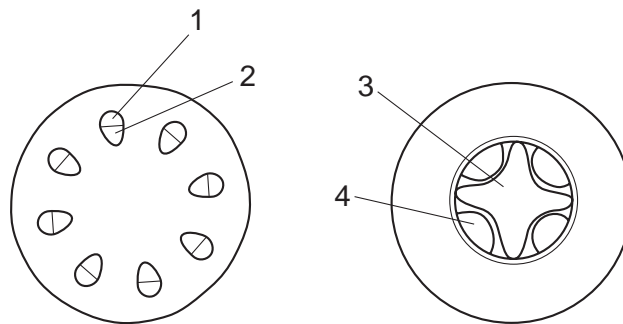
In which layer of cells does most photosynthesis occur?



5 Which substance must be present in the diet to prevent scurvy?

- A** calcium
- B** iron
- C** vitamin C
- D** vitamin D

- 6 Which process in the human body does **not** depend on energy from respiration?
- A cell division
 B diffusion
 C muscle contraction
 D passage of a nerve impulse
- 7 Through which vessel does oxygenated blood enter the heart?
- A aorta
 B pulmonary artery
 C pulmonary vein
 D vena cava
- 8 The diagram shows a cross-section of a stem and a cross-section of a root.

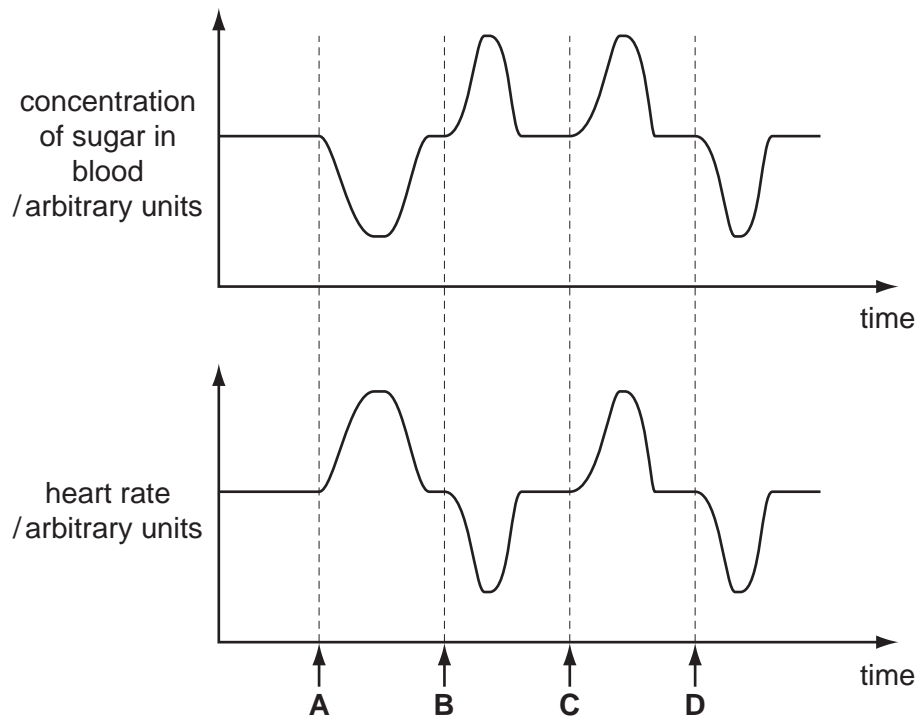


Which row identifies the tissues in the stem and root?

	tissues			
	1	2	3	4
A	phloem	xylem	phloem	xylem
B	phloem	xylem	xylem	phloem
C	xylem	phloem	phloem	xylem
D	xylem	phloem	xylem	phloem

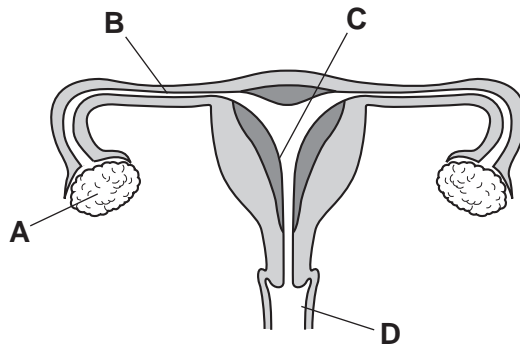
- 9 The graphs show changes in the rate of heartbeat and in the concentration of sugar in the blood over the same period of time.

When was adrenaline secreted?

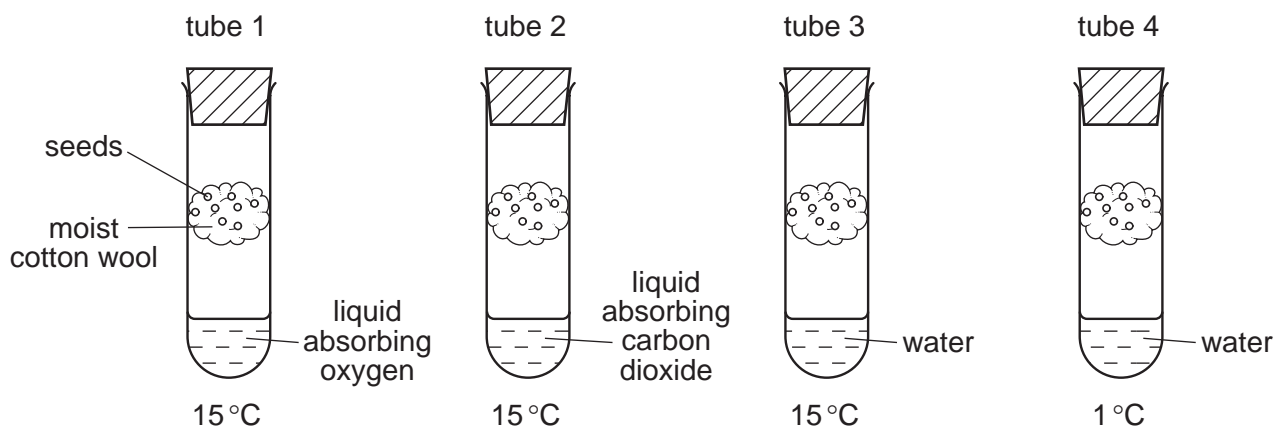


- 10 The diagram shows the female reproductive system.

Where does successful implantation normally occur?



11 The diagram shows four test-tubes set up to investigate the conditions needed for germination.



In which test-tubes will the seeds germinate?

- A** 1 and 2 **B** 2 and 3 **C** 3 and 4 **D** 4 and 1

12 Which process takes carbon dioxide out of the air?

- A** combustion
B decomposition
C photosynthesis
D plant respiration

13 Which are possible harmful effects of deforestation?

	global warming	species extinction
A	✓	✓
B	✓	x
C	x	✓
D	x	x

14 Which substance is liquid at 25°C?

	melting point /°C	boiling point /°C
A	-182	-161
B	-100	80
C	-77	-34
D	44	280

- 15 Two oxides have the formulae CaO and NO.

Which statement about the bonding in these oxides is correct?

- A Both CaO and NO are covalent.
- B Both CaO and NO are ionic.
- C CaO is covalent and NO is ionic.
- D CaO is ionic and NO is covalent.

- 16 Information about two minerals is given below.

name of mineral	formula
anorthite	$\text{CaAl}_2\text{Si}_2\text{O}_8$
orthoclase	KAlSi_3O_8

Which information about anorthite and orthoclase is correct?

- A Both of them contain the same number of atoms of aluminium.
- B Both of them contain the same number of atoms of oxygen.
- C Orthoclase contains the same number of atoms of both aluminium and silicon.
- D Orthoclase contains twice as many atoms of aluminium as anorthite.

- 17 The equations show some chemical reactions.

- 1 copper + oxygen → copper oxide
- 2 copper carbonate → copper oxide + carbon dioxide
- 3 copper oxide + carbon → copper dioxide + water

In which equations has the underlined substance been reduced?

- A 1 only B 1 and 2 C 2 and 3 D 3 only

- 18 Which would decrease the speed of a reaction?

- A adding a suitable catalyst
- B decreasing the concentration
- C decreasing the particle size
- D raising the temperature

19 Four pairs of oxides are listed.

- W calcium oxide and sodium oxide
 X calcium oxide and sulfur dioxide
 Y nitrogen dioxide and sodium oxide
 Z sulfur dioxide and nitrogen dioxide

Which pairs of oxides would neutralise each other?

- A** pair W and pair X
B pair W and pair Z
C pair X and pair Y
D pair Y and pair Z

20 Aqueous sodium hydroxide and aqueous ammonia are added separately to aqueous solutions X and Y.

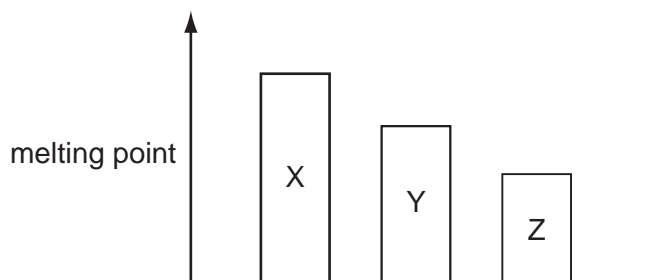
The results are shown below.

	aqueous sodium hydroxide added	aqueous ammonia added
solution X	light blue precipitate, insoluble in excess	light blue precipitate, soluble in excess to form a dark blue solution
solution Y	green precipitate, insoluble in excess	green precipitate, insoluble in excess

Which metals ions are present in solution X and solution Y?

	solution X	solution Y
A	Cu(II)	Fe(III)
B	Cu(II)	Fe(II)
C	Fe(II)	Cu(II)
D	Fe(III)	Cu(II)

21 The diagram shows the trend in melting point for three elements X, Y and Z.



What could X, Y and Z be?

	X	Y	Z
A	Cl	Ar	K
B	Cl	Br	I
C	H	Li	C
D	Li	Na	K

22 The table gives information about four elements.

Which element is most likely to be a transition metal?

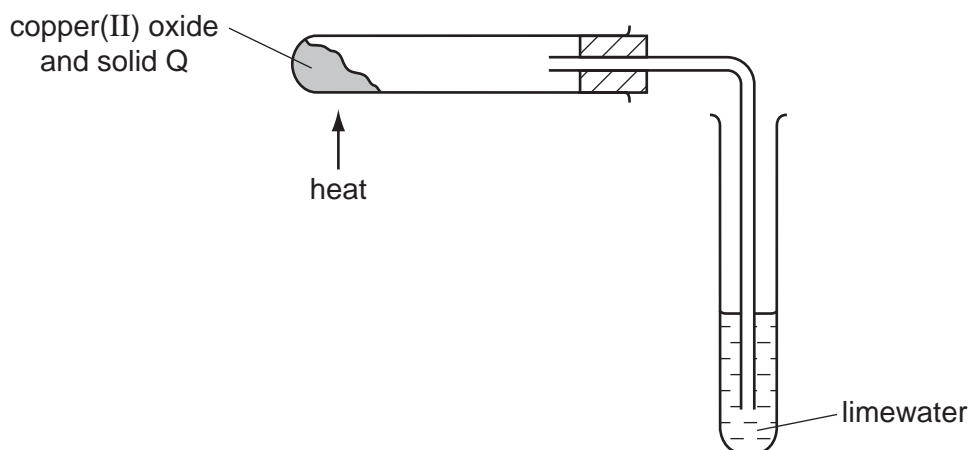
	appearance of its compound	density (g/cm ³)
A	coloured	0.97
B	coloured	7.2
C	white	0.97
D	white	7.2

23 Which property of a metal determines the method used to extract the metal from its ore?

- A** the melting point of the metal
- B** the position of the metal in the Periodic Table
- C** the reactivity of the metal
- D** the relative atomic mass, A_r , of the metal

24 Copper(II) oxide is mixed with a solid Q.

On heating the mixture, a reaction occurs and the limewater turns milky.



What is solid Q?

- A carbon
- B iron
- C sulfur
- D zinc

25 The diagrams show molecules of four gases present in clean air. Different circles represent atoms of different elements.



Which elements are shown as ● and ○?

	●	○
A	hydrogen	nitrogen
B	hydrogen	oxygen
C	oxygen	hydrogen
D	oxygen	nitrogen

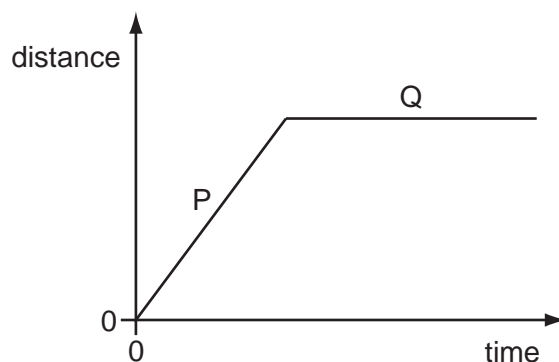
26 Which property of the compounds in petroleum is used to separate it into useful fractions?

- A boiling point
- B density
- C melting point
- D solubility

27 Which equation shows the complete combustion of a hydrocarbon?

- A $C_2H_4 + 2O_2 \rightarrow 2CO + 2H_2O$
- B $C_2H_4 + 3O_2 \rightarrow 2CO_2 + 2H_2O$
- C $C_2H_6O + 2O_2 \rightarrow 2CO + 3H_2O$
- D $C_2H_6O + 3O_2 \rightarrow 2CO_2 + 3H_2O$

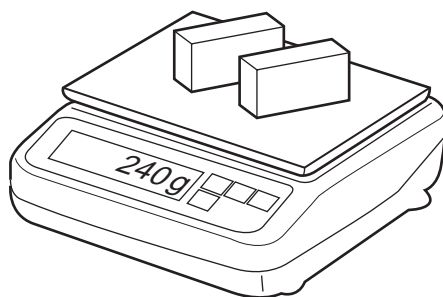
28 The graph is the distance/time graph for a bicycle journey.



Which row describes the behaviour of the bicycle in part P and the behaviour of the bicycle in part Q of the graph?

	part P	part Q
A	moving at constant speed	moving at constant speed
B	moving at constant speed	not moving
C	moving at increasing speed	moving at constant speed
D	moving at increasing speed	not moving

- 29 A shop-keeper places two identical blocks of cheese on a set of scales and notices that their combined mass is 240g. Each block measures 2.0 cm × 5.0cm × 10.0 cm.



What is the density of the cheese?

- A 0.42g/cm³ B 0.83g/cm³ C 1.2g/cm³ D 2.4g/cm³
- 30 In which situation is gravitational energy converted into kinetic energy?
- A diving from a high platform
 B kicking a football along the ground
 C lifting a book on to a high shelf
 D pumping water up to a storage tank
- 31 The temperature of a liquid is below its boiling point. The liquid evaporates.

Which row is correct about where the evaporation occurs and what effect this has on the temperature of the liquid?

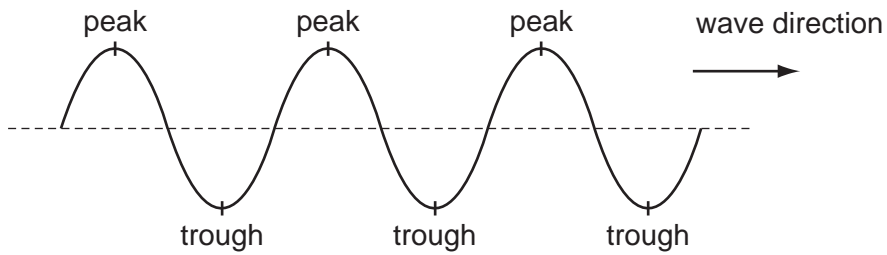
	where evaporation occurs	effect on the temperature of the liquid
A	in all parts of the liquid	temperature falls
B	in all parts of the liquid	temperature rises
C	only on the surface of the liquid	temperature falls
D	only on the surface of the liquid	temperature rises

- 32 Hot drinks may be bought in cups with a plastic lid.

What is the effect of the lid on heat transfer?

- A It mainly reduces conduction.
 B It mainly reduces convection.
 C It mainly reduces radiation.
 D It reduces conduction, convection and radiation equally.

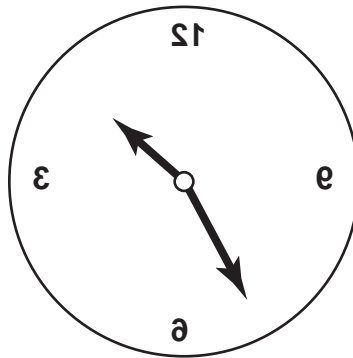
33 The diagram shows the peaks and troughs of a water wave.



What is the wavelength of the water wave?

- A the depth of one wave trough
- B the distance between one wave trough and the next
- C the distance travelled by one wave trough in one second
- D the number of wave troughs passing a point in one second

34 The image of a clock face as seen in a plane mirror is shown.

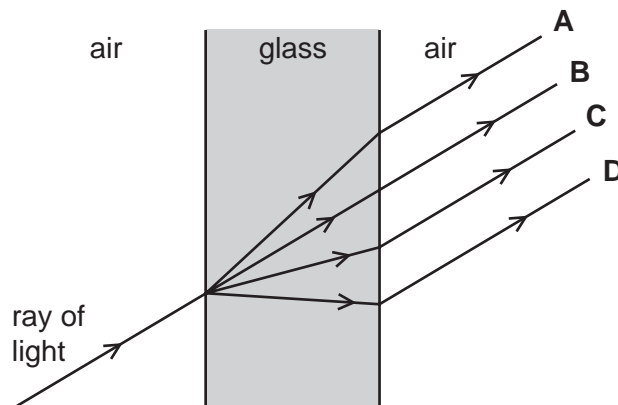


What is the time on the clock?

- A 1.25
- B 1.35
- C 10.25
- D 10.35

35 A ray of light passes through a glass window.

Which path does it take?

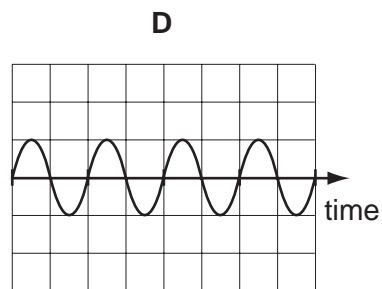
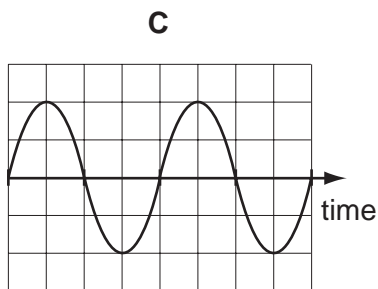
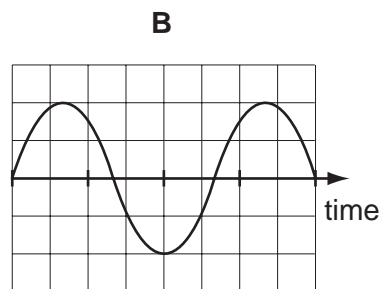
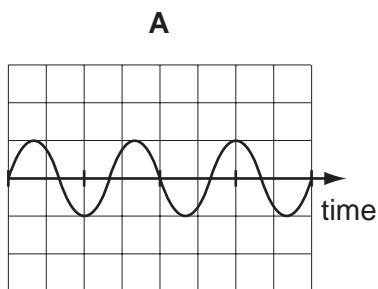


36 Which statement about the electromagnetic spectrum is correct?

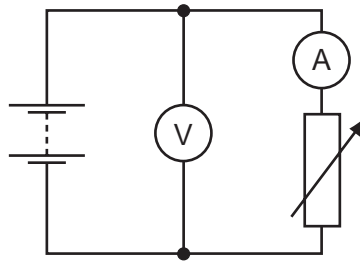
- A Gamma rays have the highest frequency.
- B Microwaves have the smallest wavelength.
- C Ultraviolet waves have the largest wavelength.
- D Visible light waves have the lowest frequency.

37 The diagrams represent four different sound waves. The scales are the same in all the diagrams.

Which sound has the lowest pitch?



- 38 The diagram represents a circuit which includes a battery, an ammeter, a voltmeter and a variable resistor.



What happens to the readings on the meters as the resistance of the variable resistor is increased?

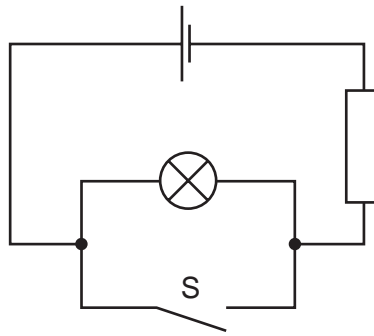
	ammeter reading	voltmeter reading
A	decreases	decreases
B	decreases	stays constant
C	increases	decreases
D	increases	stays constant

- 39 An electric fire draws a current of 10 A. Fuses are available rated at 5 A and 13 A.

Which fuse rating should be used with this electric fire, and where should it be fitted in the circuit?

	fuse rating	where in circuit
A	5 A	live wire
B	5 A	neutral wire
C	13 A	live wire
D	13 A	neutral wire

40 The diagram shows a circuit.



What happens to the lamp when switch S is closed?

- A It does not light at all.
- B It lights and then blows.
- C It lights more brightly.
- D It lights more dimly.

DATA SHEET
The Periodic Table of the Elements

		Group										
		I	II	III	IV	V	VI	VII	VIII	IX	X	
		1 H Hydrogen 1										
7	9											
Li Lithium 3	Be Beryllium 4											
23	24											
Na Sodium 11	Mg Magnesium 12											
39	40											
K Potassium 19	Ca Calcium 20	45	48	51	52	55	56	59	59	64	65	
		Sc Scandium 21	Ti Titanium 22	V Vanadium 23	Cr Chromium 24	Mn Manganese 25	Fe Iron 26	Co Cobalt 27	Ni Nickel 28	Cu Copper 29	Zn Zinc 30	
85	88	89	91	93	96	101	101	103	106	108	112	
Rb Rubidium 37	Sr Strontium 38	Y Yttrium 39	Zr Zirconium 40	Nb Niobium 41	Mo Molybdenum 42	Ru Ruthenium 44	Rh Rhodium 45	Pd Palladium 46	Ag Silver 47	Cd Cadmium 48		
133	137	139	178	181	184	190	190	192	195	197	201	
Cs Caesium 55	Ba Barium 56	La Lanthanum 57	Hf Hafnium 72	Ta Tantalum 73	W Tungsten 74	Os Osmium 76	Ir Iridium 77	Pt Platinum 78	Au Gold 79	Hg Mercury 80		
226	227	227										
Fr Francium 87	Ra Radium 88	Ac Actinium 89										
*58-71 Lanthanoid series †90-103 Actinoid series												
		140	141	144	150	152	157	159	162	165	167	
		Ce Cerium 58	Pr Praseodymium 59	Nd Neodymium 60	Pm Promethium 61	Sm Samarium 62	Eu Europium 63	Gd Gadolinium 64	Dy Dysprosium 66	Ho Holmium 67	Er Erbium 68	
		232	238	238	238	238	238	238	238	238	238	
		Th Thorium 90	Pa Protactinium 91	U Uranium 92	Np Neptunium 93	Pu Plutonium 94	Am Americium 95	Cm Curium 96	Cf Californium 98	Es Einsteinium 99	Fm Fermium 100	
		169	173	173	173	173	173	173	173	173	173	
		Tm Thulium 69	Yb Ytterbium 70	Lu Lutetium 71	No Nobelium 102	Md Mendelevium 101	Fm Fermium 100	Er Erbium 68	Ho Holmium 67	Dy Dysprosium 66	Ho Holmium 67	
		86	86	86	86	86	86	86	86	86	86	
		Rn Radon 86	At Astatine 85	Po Polonium 84	Bi Bismuth 83	Pb Lead 82	Tl Thallium 81	Pt Platinum 78	Au Gold 79	Hg Mercury 80	Pb Lead 82	
		131	131	131	131	131	131	131	131	131	131	
		Xe Xenon 54	I Iodine 53	Te Tellurium 52	Sb Antimony 51	Sn Tin 50	In Indium 49	Ag Silver 47	Cd Cadmium 48	Hg Mercury 80	Pb Lead 82	
		36	36	36	36	36	36	36	36	36	36	
		Kr Krypton 36	Br Bromine 35	Se Selenium 34	As Arsenic 33	Ge Germanium 32	Ga Gallium 31	Cu Copper 29	Zn Zinc 30	Hg Mercury 80	Pb Lead 82	
		18	18	18	18	18	18	18	18	18	18	
		Ar Argon 18	Cl Chlorine 17	S Sulfur 16	P Phosphorus 15	Si Silicon 14	Al Aluminium 13	B Boron 5	C Carbon 6	N Nitrogen 7	O Oxygen 8	
		10	10	10	10	10	10	10	10	10	10	
		Ne Neon 10	F Fluorine 9	O Oxygen 8	N Nitrogen 7	C Carbon 6	B Boron 5	He Helium 2	Ne Neon 10	O Oxygen 8	F Fluorine 9	
		2	2	2	2	2	2	2	2	2	2	
		He Helium 2	Ne Neon 10	Ar Argon 18	Kr Krypton 36	Xe Xenon 54	Rn Radon 86	At Astatine 85	Po Polonium 84	Bi Bismuth 83	Pb Lead 82	

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

	a	X	a = relative atomic mass
Key	b	X	X = atomic symbol
			b = proton (atomic) number

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