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0653/23

October/November 2013

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

You may lose marks if you do not show your working or if you do not use appropriate units.

A copy of the Periodic Table is printed on page 20.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

This document consists of **19** printed pages and **1** blank page.

1 Fig. 1.1 shows a root hair cell.

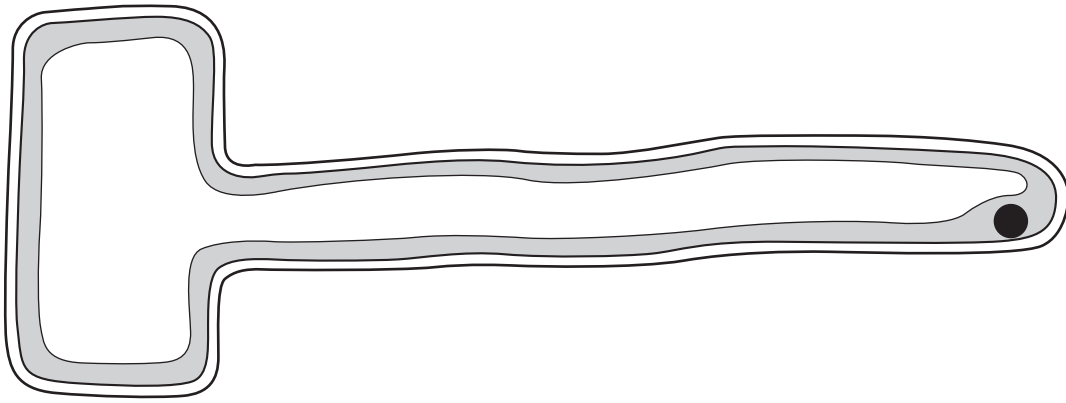


Fig. 1.1

(a) Use the letters **A**, **B** and **C** to label these parts of the root hair cell in Fig. 1.1.

A the cell membrane

B the part that contains chromosomes

C a structure that is **not** present in animal cells

[3]

(b) Name **two** substances that are absorbed by root hair cells.

1

2

[2]

- (c) Fig. 1.2 shows part of a plant stem from which the outer layer, including the phloem, has been removed.

For
Examiner's
Use

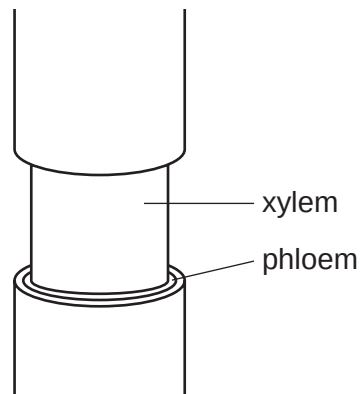


Fig. 1.2

- (i) State the function of phloem.

.....

.....

..... [2]

- (ii) Suggest why this treatment would cause the roots of the plant to die.

.....

.....

.....

..... [2]

- 2 (a) Table 2.1 shows information about some chemical elements and their positions in the Periodic Table.

For
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Use

Table 2.1

element	group number in the Periodic Table
oxygen	6
calcium	2
lithium	1
sulfur	6
fluorine	7

- (i) State the noble (inert) gas that is in the same period of the Periodic Table as sulfur.

..... [1]

- (ii) Select **two** elements from Table 2.1 whose atoms form ionic chemical bonds with each other and explain your answer.

..... and

explanation

..... [2]

- (b) Fig. 2.1 shows a diagram of an atom.

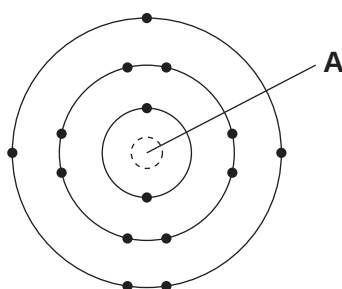


Fig. 2.1

- (i) Name structure **A** in Fig 2.1. [1]

- (ii) State the proton number of the atom in Fig. 2.1.

Explain your answer briefly.

proton number

explanation

..... [2]

*For
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Use*

- 3 Fig. 3.1 shows a circuit used to measure the current passing through a resistor when the voltage across it is changed.

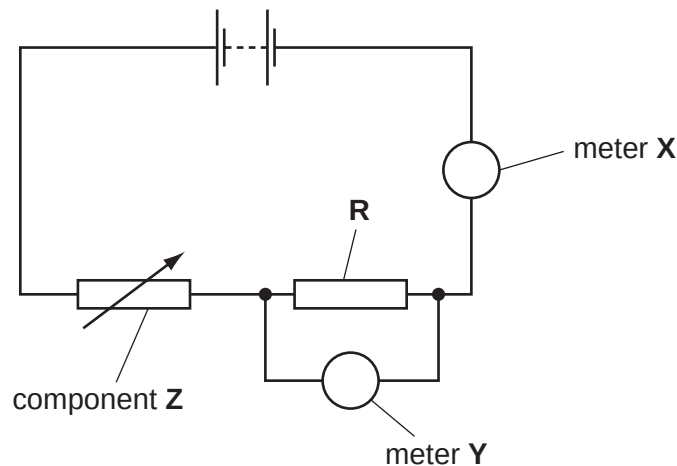


Fig. 3.1

- (a) Describe the purpose of component **Z** in the circuit.

..... [1]

- (b) The meters shown in the circuit give readings of 0.6 A and 8.0 V.

State which meter, **X** or **Y**, gives the reading of 0.6 A.

Explain your answer.

meter

explanation

..... [1]

- (c) Use the formula

$$\text{resistance} = \text{potential difference} / \text{current}$$

to calculate the resistance of the resistor.

State the units for your answer.

working

..... unit [2]

4 Soya beans are an important crop in Brazil.

- (a) Table 4.1 contains information about the tests used and results obtained when testing soya beans for protein, fat and starch.

Table 4.1

nutrient tested for	reagent	result	conclusion
protein		purple	
starch			contains starch
fat		milky white	

Complete the table.

[3]

- (b) Explain why protein is an important part of a balanced diet.

.....

 [2]

- (c) When a person eats soya beans, the beans are chewed in the mouth.

Explain why this makes it easier for enzymes in the digestive system to digest the beans.

.....

 [2]

- (d) Large areas of rainforest have been cleared in Brazil, to provide more land for growing soya beans.

State **two** ways in which cutting down the rainforest can harm the environment.

1

 2
 [2]

For
Examiner's
Use

- 5 (a) A student placed four equally-sized pieces of different metals into colourless liquids contained in four test-tubes **P**, **Q**, **R** and **S**.

For
Examiner's
Use

Fig. 5.1 shows what the student observed.

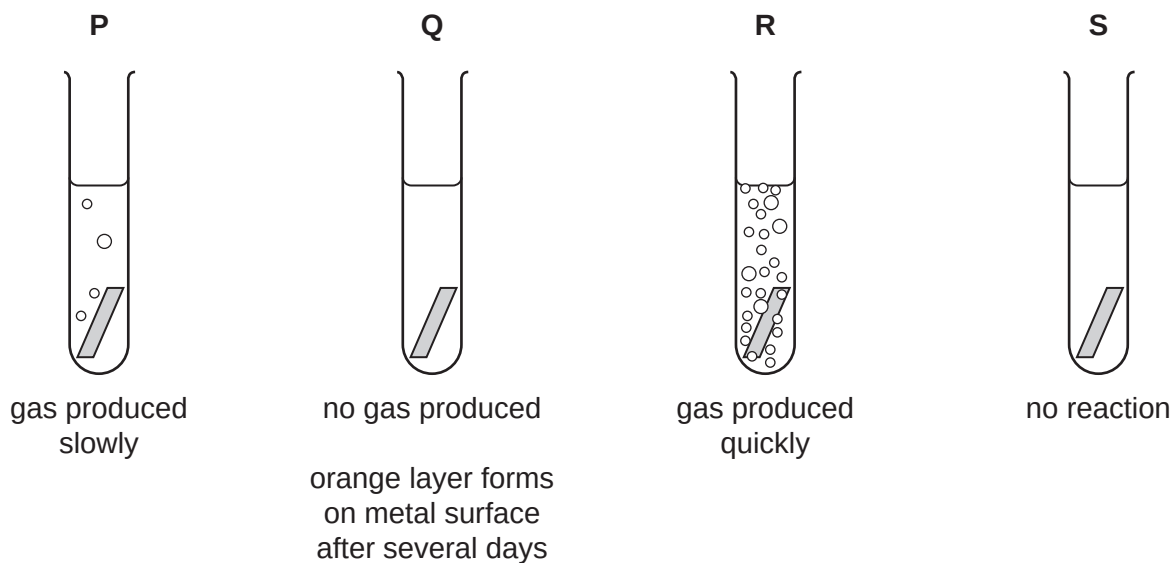


Fig. 5.1

- (i) Suggest which of the test-tubes in Fig. 5.1 contained water to which a piece of iron was added.

Explain your answer.

test-tube

explanation

.....

.....

..... [3]

- (ii) The colourless liquid in test-tube **R** was dilute hydrochloric acid.

Suggest the name of the metal that was added to test-tube **R** and name the gas that was produced.

metal

gas [2]

- (iii) Test-tube **P** contained the same concentration of dilute hydrochloric acid at the same temperature as test-tube **R**.

Suggest a reason why gas was produced more slowly in test-tube **P** than in test-tube **R**.

.....
 [1]

- (b) Gasoline and diesel are mixtures of liquid hydrocarbons obtained from petroleum by the process of fractional distillation.

- (i) State **one** difference in the properties of the hydrocarbons in gasoline that allows them to be separated by fractional distillation.

.....
 [1]

- (ii) State the main use of gasoline and explain, in terms of its chemical properties, why it is suitable for this use.

use

explanation

..... [2]

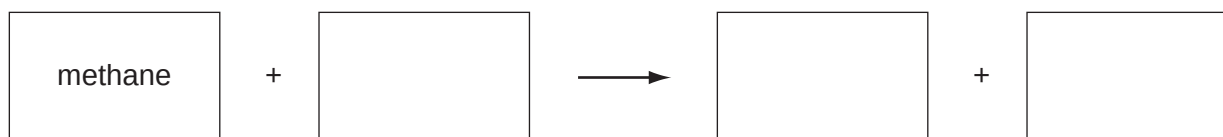
- (c) Natural gas contains mainly methane.

- (i) Complete the diagram of the structure of one molecule of methane.



[1]

- (ii) Complete the **word** chemical equation for the complete combustion of methane.



[2]

- 6 (a) Fig. 6.1 gives information about the uses of different types of electromagnetic waves and their effects on living tissue.

For
Examiner's
Use

Draw lines to link each electromagnetic wave with its effect on living tissue and its use. One has been completed as an example.

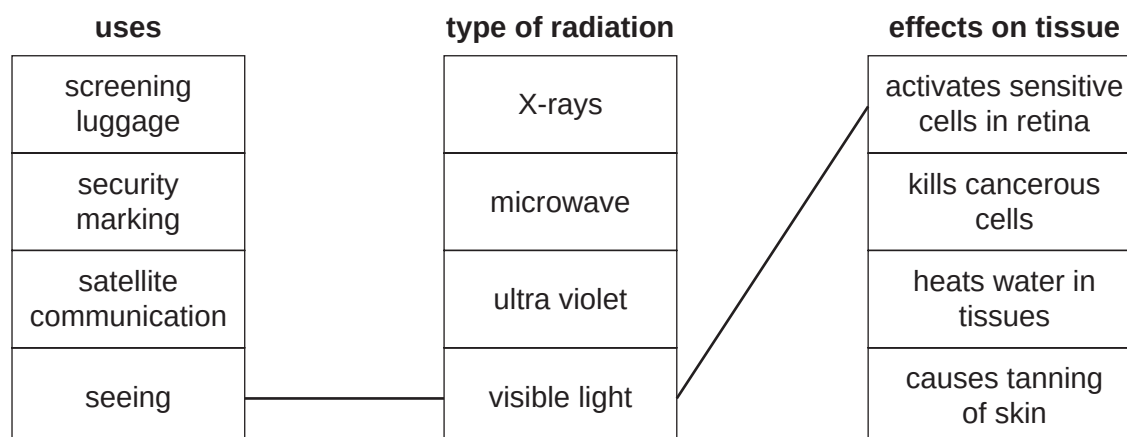


Fig. 6.1

[4]

- (b) Electromagnetic waves are transverse waves. Water waves are also transverse.

Draw a diagram of a transverse wave on the axes below. Label the amplitude and **one** wavelength on your diagram.



[3]

(c) Fig. 6.2 shows a person looking into a mirror and seeing an image.

For
Examiner's
Use

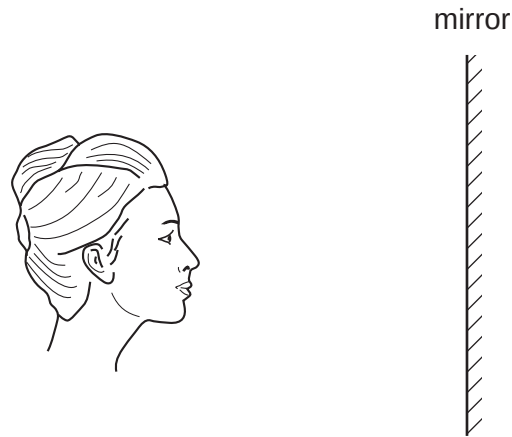


Fig. 6.2

- (i) Write the letter **X** on Fig. 6.2 to show the position of the image of the person's nose. [2]
- (ii) Select **three** words or phrases from the list that describe the image correctly.

larger than object	real	same size as object
smaller than object	upright	upside down
		virtual

.....

..... [3]

7 Fig. 7.1 shows the contents of the human thorax (chest).

For
Examiner's
Use

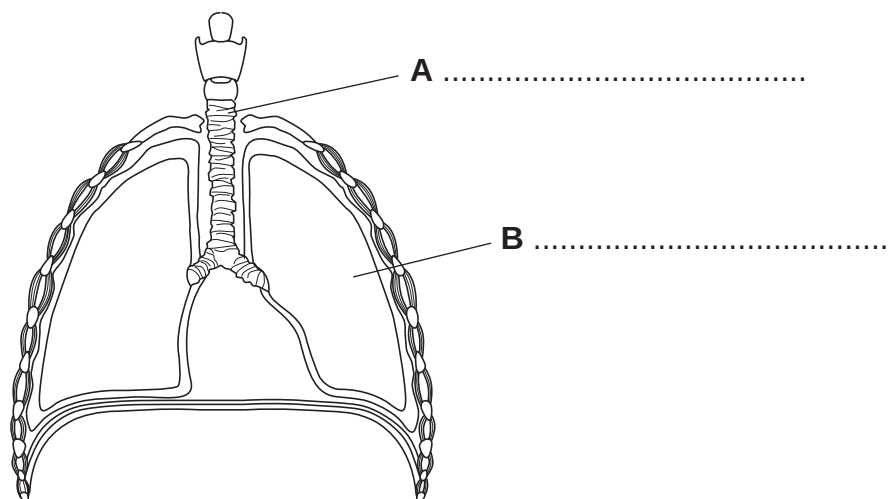


Fig. 7.1

(a) On Fig. 7.1, name structures **A** and **B**. [2]

(b) Oxygen diffuses into the blood from the alveoli inside the lungs. Carbon dioxide diffuses into the alveoli from the blood.

(i) Define the term *diffusion*.

.....

 [2]

(ii) Name the component of blood that transports dissolved carbon dioxide.

..... [1]

(iii) When a person is doing vigorous exercise, the concentration of carbon dioxide in the blood increases.

Explain why this happens.

.....

 [2]

- (iv) Suggest how this will affect the rate of diffusion of carbon dioxide from the blood to the alveoli.

*For
Examiner's
Use*

Explain your answer.

effect on rate of diffusion

explanation

..... [2]

- 8 (a) Fig. 8.1 shows apparatus a student used to investigate the reaction between dilute nitric acid and excess calcium carbonate.

For
Examiner's
Use

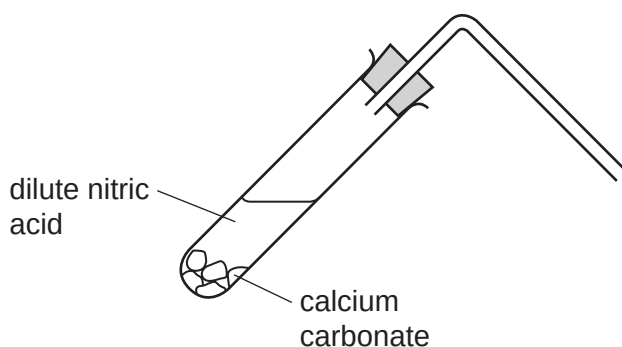


Fig. 8.1

- (i) Describe how the student could show that the reaction in Fig. 8.1 produced carbon dioxide. You may complete the diagram to help you answer this question.

.....

 [2]

- (ii) At the end of the reaction the test-tube in Fig. 8.1 contains a solution of the compound calcium nitrate.

State the general name for compounds like calcium nitrate which are produced when an acid reacts with a metal carbonate.

..... [1]

- (iii) The chemical formula of calcium nitrate is $\text{Ca}(\text{NO}_3)_2$.

State the total number of atoms and the number of different elements that are shown combined together in this formula.

total number of atoms

number of different elements [2]

- (b) The student then carried out an investigation into the way that the rate of the reaction in (a) changed when he varied the concentration of the nitric acid.

For
Examiner's
Use

Fig. 8.2 shows the apparatus the student used to measure the rate of reaction.

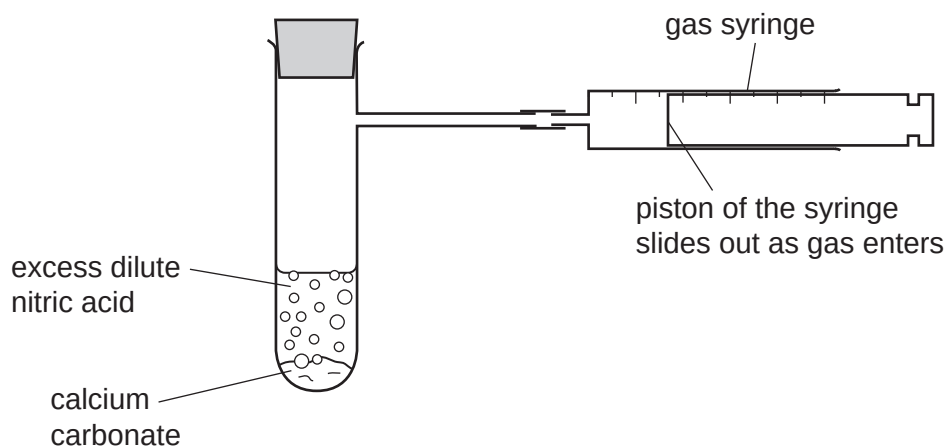


Fig. 8.2

The student measured the rate of reaction by finding how long it took for the gas syringe to fill with gas.

- (i) After he had completed several measurements, the student wrote the following correct conclusion in his notebook.

	Conclusion
	The higher the pH of the dilute nitric acid
	the longer it took for the gas syringe to
	fill with gas.

Explain this conclusion briefly.

.....

.....

..... [2]

- (ii) State **two** other variables that can affect the rate of reaction between dilute nitric acid and calcium carbonate.

1

2 [2]

- 9 Fig. 9.1 shows a solar- powered golf cart used to carry golfers around a golf course.

For
Examiner's
Use

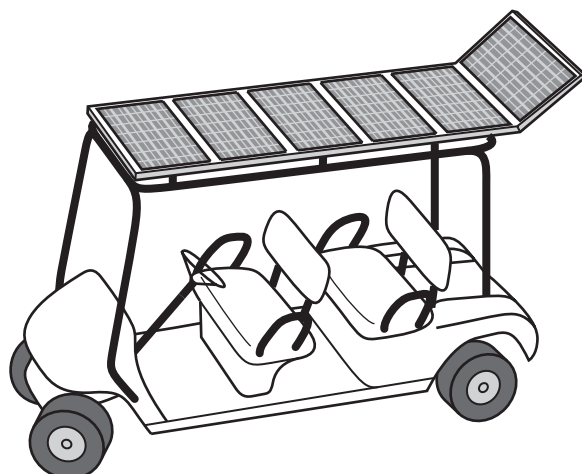


Fig. 9.1

- (a) As the cart moves around the course, the motion of the cart is measured.

Fig. 9.2 shows a distance / time graph for a small part of the journey lasting 60 seconds.

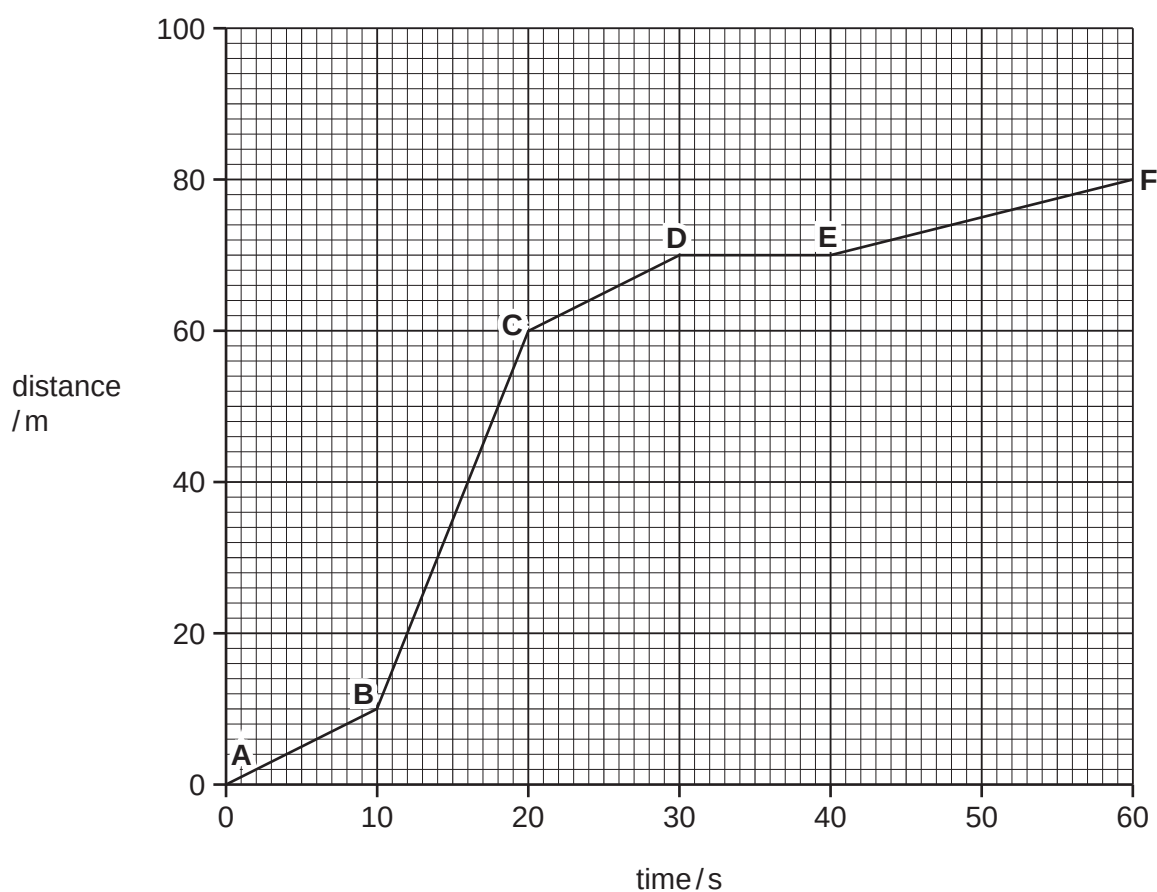


Fig. 9.2

- (i) Write down the total distance covered in 60 s. m [1]

- (ii) Describe the motion of the cart between **D** and **E**.

.....
 [1]

- (iii) During another part of the journey, the cart is accelerating.

State whether the forces acting on the cart are balanced or unbalanced.

Explain your answer.

.....
 [1]

- (b) The cart is powered by solar cells on its roof. The solar cells produce electrical energy used to charge the rechargeable batteries in the cart.

Name **one** other renewable energy resource that could produce electrical energy.

..... [1]

- (c) The golfer hits a golf ball with his club. The ball flies through the air.

- (i) State the form of energy given to the golf ball when the ball is hit.

..... [1]

- (ii) State the form of energy gained by the golf ball as it rises into the air after being hit.

..... [1]

- (d) The mass of a golf ball is 45g. The volume of a golf ball is 36 cm^3 .

Calculate the density of the golf ball.

State the formula that you use and show your working.

formula

working

..... g/cm^3 [2]

- (e) The head of the golf club is made of solid metal. The air the golf ball is travelling through is a gas.

For
Examiner's
Use

Complete Fig. 9.3 below to show the arrangement of particles in a gas. The diagram for a solid has been done for you.

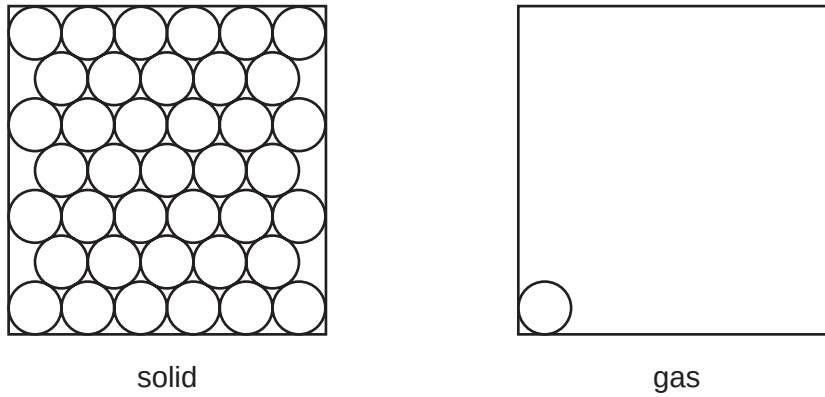


Fig. 9.3

[2]

DATA SHEET
The Periodic Table of the Elements

Group																	
I	II											III	IV	V	VI	VII	0
1 H Hydrogen 9 Be Beryllium 23 Na Sodium 39 K Potassium 85 Rb Rubidium 133 Cs Cesium Fr Francium 7 Li Lithium 11 Mg Magnesium 19 Ca Calcium 37 Sr Strontium 55 Ba Barium Ra Radium 4 B Boron 12 Al Aluminium 20 Si Silicon 38 Ga Gallium 56 Ge Germanium 84 As Arsenic 101 Se Selenium 127 Br Bromine 157 Kr Krypton 184 Xe Xenon 222 Rn Radon 5 C Carbon 13 N Nitrogen 21 O Oxygen 35 F Fluorine 53 Ne Neon 71 Ar Argon 89 Cl Chlorine 107 S Sulfur 125 P Phosphorus 151 Sb Antimony 179 Te Tellurium 210 I Iodine 289 At Astatine 6 N Nitrogen 14 P Phosphorus 22 As Arsenic 34 Sb Antimony 51 Bi Bismuth 83 Po Polonium 115 At Astatine 151 Fr Francium 187 Ac Actinium 227 Ra Radium 285 Fr Francium 315 Ac Actinium 11 B Boron 19 Al Aluminium 27 Ga Gallium 35 In Indium 49 Tl Thallium 81 Pb Lead 127 Bi Bismuth 159 Po Polonium 209 At Astatine 269 Rn Radon 12 C Carbon 20 Si Silicon 28 Ge Germanium 36 Se Selenium 52 Te Tellurium 82 Po Polonium 114 At Astatine 150 Fr Francium 186 Ac Actinium 226 Ra Radium 286 Fr Francium 316 Ac Actinium 13 N Nitrogen 21 P Phosphorus 29 As Arsenic 37 Sb Antimony 51 Bi Bismuth 83 Po Polonium 115 At Astatine 151 Fr Francium 187 Ac Actinium 227 Ra Radium 287 Fr Francium 317 Ac Actinium 14 O Oxygen 22 S Sulfur 34 Se Selenium 52 Te Tellurium 82 Po Polonium 114 At Astatine 150 Fr Francium 186 Ac Actinium 226 Ra Radium 286 Fr Francium 316 Ac Actinium 15 F Fluorine 23 Cl Chlorine 35 Br Bromine 53 I Iodine 85 At Astatine 117 Fr Francium 153 Ac Actinium 189 Ra Radium 229 Fr Francium 289 Ac Actinium 319 Ac Actinium 16 Ne Neon 24 Ar Argon 36 Kr Krypton 54 Xe Xenon 86 Rn Radon 118 Fr Francium 154 Ac Actinium 190 Ra Radium 230 Fr Francium 290 Ac Actinium 320 Ac Actinium																	

*58-71 Lanthanoid series
†90-103 Actinoid series

a	X	a = relative atomic mass
b	X	X = atomic symbol
	b	b = proton (atomic) number

Key

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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