



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CANDIDATE
NAME

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CENTRE
NUMBER

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CANDIDATE
NUMBER

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COMBINED SCIENCE

0653/22

Paper 2 (Core)

October/November 2012

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs, tables or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

A copy of the Periodic Table is printed on page 24.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

For Examiner's Use

1	
2	
3	
4	
5	
6	
7	
8	
9	
Total	

This document consists of **22** printed pages and **2** blank pages.



1 Fig. 1.1 shows a red blood cell and a root hair cell.

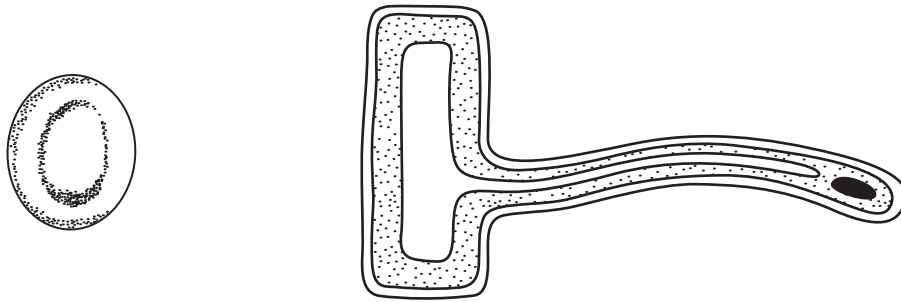


Fig 1.1

(a) Tick (✓) the boxes to show which structures are present in

- a red blood cell,
- a root hair cell.

structure	red blood cell	root hair cell
cell membrane	☐	☐
nucleus	☐	☐
chloroplast	☐	☐

[2]

(b) (i) Name the red protein found in the cytoplasm of the red blood cell.

..... [1]

(ii) State the function of a red blood cell.

..... [1]

(c) Name the colourless carbohydrate in the cell wall of the root hair cell.

..... [1]

(d) Fig. 1.2 shows a plant with its roots in a beaker of water containing a blue dye.

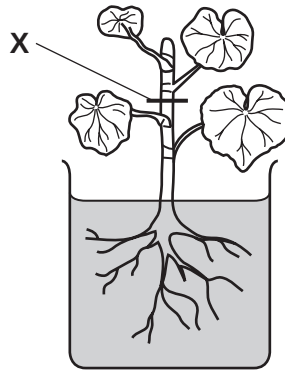


Fig. 1.2

After 10 minutes, the veins in the leaves of the plant became blue.

(i) Explain why the veins in the leaves became blue.

.....

.....

..... [2]

(ii) A student cut the stem of the plant at **X**. Fig. 1.3 shows the appearance of the cut stem seen through a microscope.

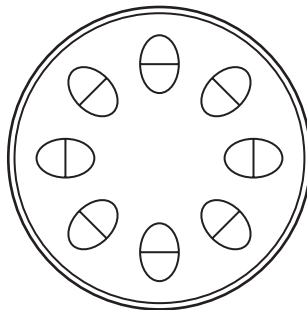


Fig. 1.3

On Fig. 1.3, use a pencil to shade the part that would look blue.

[1]

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- 2 (a) The proton (atomic) number of the element fluorine is 9. Fluorine is found in Period 2 and Group 7 of the Periodic Table.

- (i) Predict the number of electrons in one atom of fluorine.

Explain your answer.

total number of electrons

explanation

..... [2]

- (ii) Predict and explain, in terms of its position in the Periodic Table, whether this element would be an electrical conductor or an insulator.

.....

..... [1]

- (b) The halogens are reactive elements found in Group 7 of the Periodic Table.

Halogens combine vigorously with the alkali metals from Group 1 to form colourless ionic compounds. The halogens and alkali metals from Periods 2 to 4 are shown in Fig. 2.1.

Li	F
Na	Cl
K	Br

Fig. 2.1

- (i) The alkali metals react with water to produce an alkaline solution and a gaseous element.

State and explain briefly which **one** of the alkali metals shown in Fig. 2.1 reacts most vigorously with water.

alkali metal

explanation

..... [2]

- (ii) Name the gas which is given off during the reaction in (i) and describe a test for this gas.

name

test

..... [3]

- (iii) Describe how potassium and bromine atoms become strongly bonded together when they react to form potassium bromide.

You may draw a diagram if it helps your answer.

.....

.....

.....

.....

..... [3]

- (c) A student adds a solution containing chlorine to a colourless solution of potassium bromide as shown in Fig. 2.2.

For
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Use

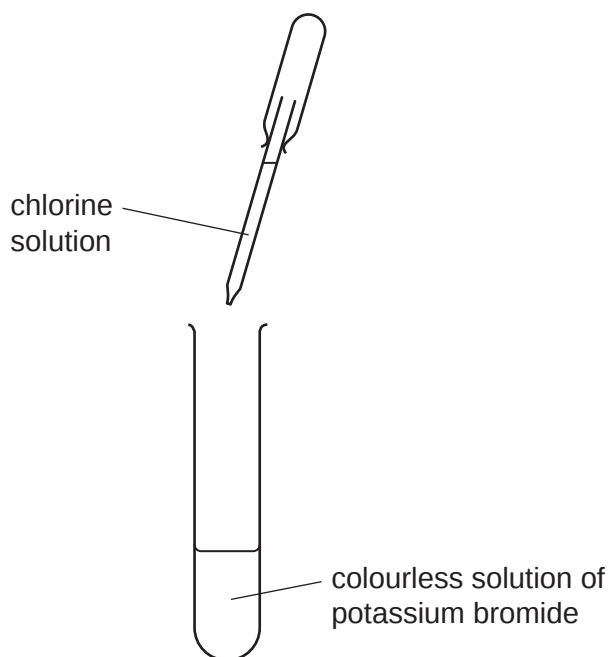


Fig. 2.2

Describe and explain briefly what is observed when chlorine and potassium bromide react.

observation

.....

explanation

..... [2]

- 3 Fig. 3.1 shows four swimmers at the start of a race.

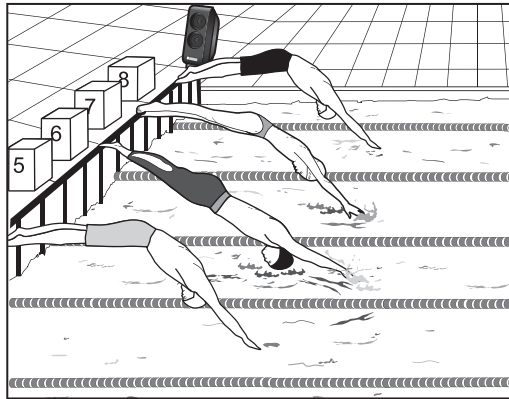


Fig. 3.1

- (a) State the form of energy which the swimmers lose as they fall from their starting positions into the water.

..... [1]

- (b) The swimmers start their race when they hear a loud, high-pitched sound from a loudspeaker.

- (i) Fig. 3.2 shows the trace of a sound wave as it appears on an oscilloscope screen.

On Fig. 3.2 draw another trace of a sound wave from a sound that is louder than the one shown, but has the same pitch.

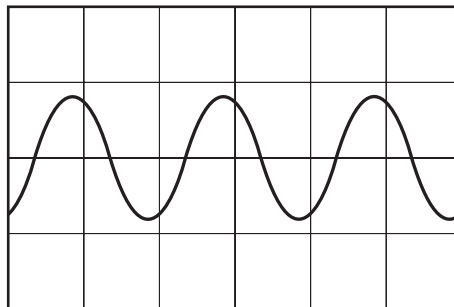


Fig. 3.2

[2]

For
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- (ii) Fig. 3.3 shows the trace of a sound wave as it appears on an oscilloscope screen.

On Fig. 3.3 draw another trace of a sound wave from a sound that has a higher pitch than the one shown, but has the same loudness.

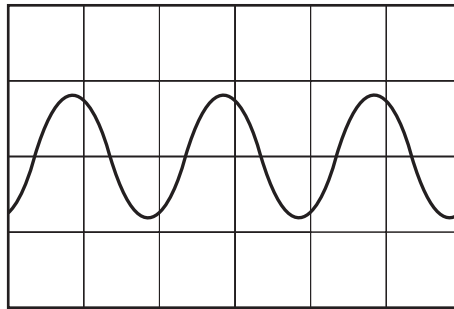


Fig. 3.3

[2]

- (iii) The swimmers can hear the sound from the loudspeaker only if the frequency of the sound lies within a range of frequencies which the human ear can detect.

State this range of frequencies.

..... Hz to Hz [1]

- (c) Sound travels at 330 m/s in air. One swimmer is 0.4 m from the loudspeaker when he hears the sound.

Calculate the time taken for the sound to travel from the loudspeaker to the swimmer.

State the formula that you use and show your working.

formula used

working

..... s [2]

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Use

- (d) When the swimmers have finished their race, they leave the pool. The water on their bodies evaporates.

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Explain in terms of particles how this evaporation takes place.

.....

.....

.....

..... [3]

- 4 (a) Fig. 4.1 shows part of a food web in a forest ecosystem.

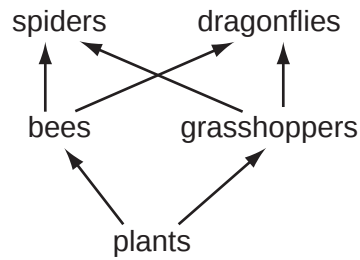


Fig. 4.1

- (i) Plants are the producers in this food web.

Define the term *producer*.

.....

 [2]

- (ii) Name **one** organism in the food web that is a carnivore.

..... [1]

- (iii) What do the arrows in the food web represent?

..... [1]

- (b) The food web shows that bees depend on plants. Some flowering plants also depend on bees and other insects to help them to reproduce.

- (i) Complete the sentences, using words from the list.

anthers	asexual	diploid	haploid
ovary	petals	sexual	stigma

Flowers are organs in which reproduction takes place.

Pollen grains are made in the

During pollination, insects carry pollen grains from one flower to another. The

pollen grains are transferred to the [3]

- (ii) After they have been pollinated, flowers produce seeds.

List **two** environmental conditions that all seeds need for germination.

1

2

[2]

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- 5 Acid indigestion is caused by unusually high levels of stomach acid. This condition may be treated by taking an antacid tablet.

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One type of antacid tablet contains a mixture of sodium hydrogencarbonate, calcium carbonate and magnesium carbonate.

A student investigated the reaction between these antacid tablets and dilute hydrochloric acid.

Fig. 5.1 shows one of the experiments the student carried out.

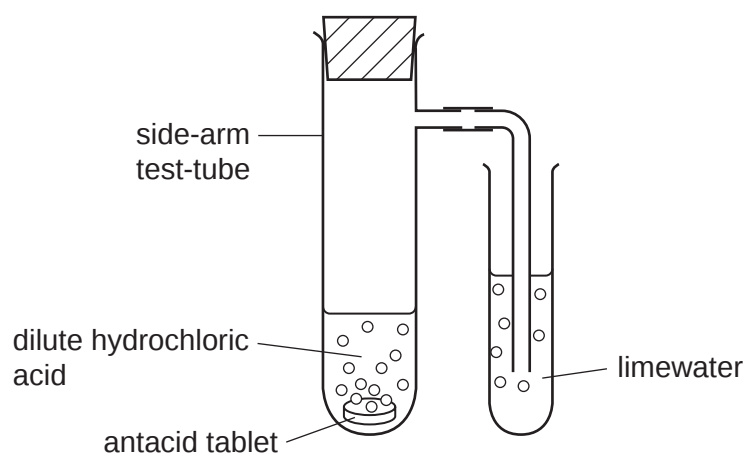


Fig. 5.1

A gas was given off when the antacid tablet reacted with the dilute hydrochloric acid. This gas reacted with the limewater.

- (a) Describe and explain the change in appearance of the limewater during the experiment.

.....

 [2]

- (b) The student used excess acid in the reaction shown in Fig. 5.1, which caused the antacid tablet to react and dissolve completely.

State the names of **two** salts that remain in the solution when the reaction is finished.

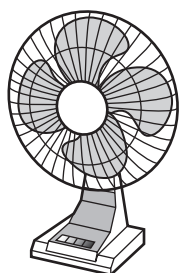
1
 2 [2]

- 6 (a) The appliances shown convert electrical energy into other forms of energy.

Complete the sentences next to each diagram to show the useful form of energy released.

For
Examiner's
Use

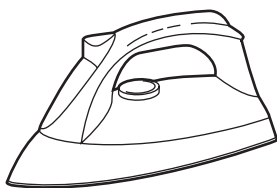
(i)



A fan converts electrical energy into

..... energy. [1]

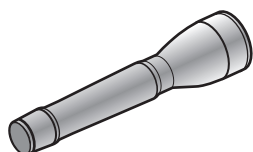
(ii)



An iron converts electrical energy into

..... energy. [1]

(iii)



A torch (flashlight) converts electrical energy into

..... energy. [1]

- (b) There are several precautions that are necessary to avoid getting an electric shock or starting a fire when using electrical appliances.

(i) State **one** precaution that must be taken when using an electrical appliance.

.....
..... [1]

(ii) For the precaution described in (i), explain why it is important.

.....
.....
..... [1]

- (c) Some torches (flashlights) use a filament lamp. Fig. 6.1 shows a circuit for measuring the current through a filament lamp as the potential difference is changed.

For
Examiner's
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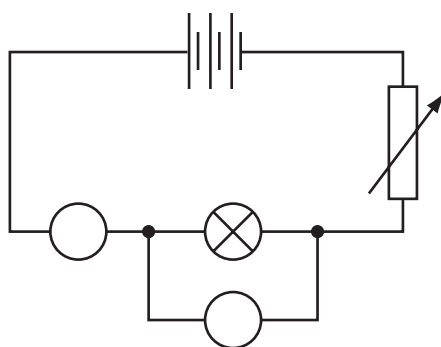


Fig. 6.1

Write the letters **A** and **V** in the two circles on the diagram. They should show the correct positions of the ammeter **A** and voltmeter **V**. [1]

- (d) Fig. 6.2 shows a graph of the results.

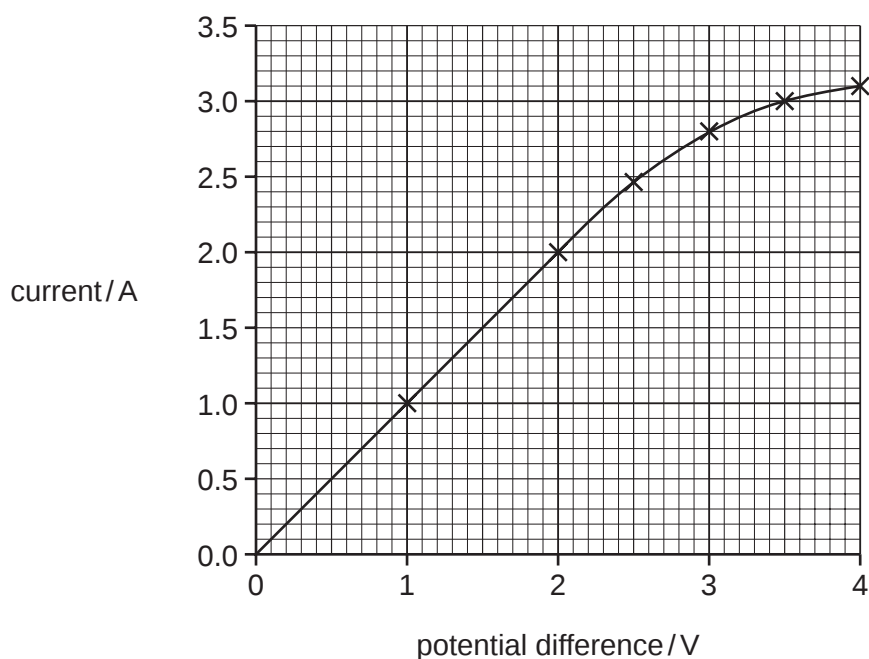


Fig. 6.2

- (i) Use the graph to find the current when the potential difference is 1.5V.

Show your working on the graph.

..... A [1]

- (ii) Describe how the current through the filament lamp changes as the voltage increases above 2.0V.

For
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Use

.....
..... [1]

- (e) A single ray of light from a torch is shone onto a mirror as shown in Fig. 6.3.

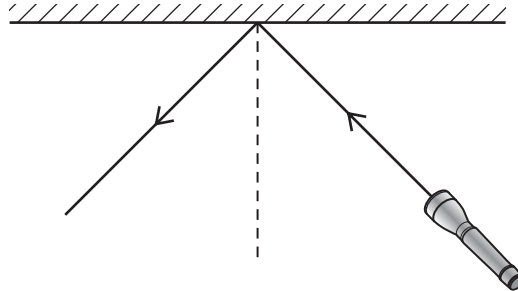


Fig. 6.3

- (i) On Fig. 6.3, label the angle of incidence and angle of reflection. [1]

- (ii) The angle of incidence = 45° .

Write down the value of the angle of reflection. [1]

7 (a) Fig. 7.1 shows the human alimentary canal.

For
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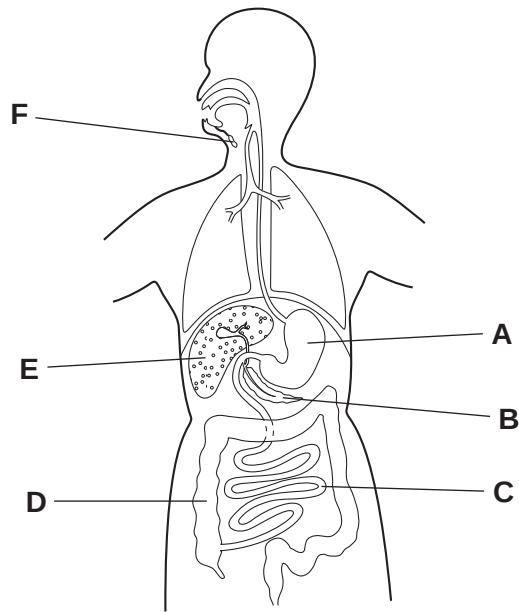


Fig. 7.1

(i) Name

part A,

part D.

[2]

(ii) State the **letter** that indicates

the liver,

the area where digested food is absorbed.

[2]

(b) Describe how the molar teeth help in the digestion of food.

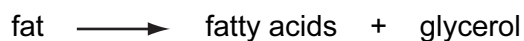
.....

.....

.....

[2]

- (c) Lipase is an enzyme that catalyses the breakdown of fats to fatty acids and glycerol.



A student carried out an experiment to investigate the effect of temperature on the rate of the breakdown of fats by lipase. Fig. 7.2 shows how she set up two test-tubes.

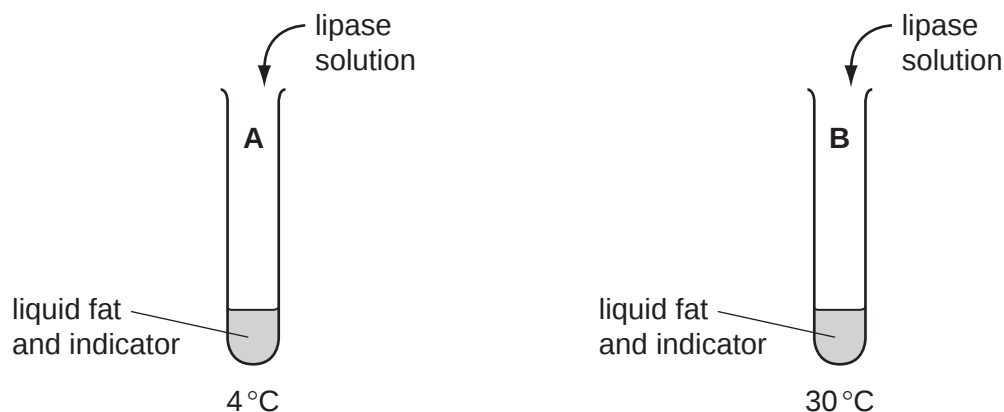


Fig. 7.2

The indicator that the student used changes colour from blue to yellow when the pH falls below 5.

Table 7.1 shows her results.

Table 7.1

time / minutes	tube A (4 °C)	tube B (30 °C)
0	blue	blue
5	blue	yellow
10	blue	yellow
15	yellow	yellow

- (i) Explain why the indicator eventually changed to yellow in both tubes.

..... [1]

- (ii) Explain the reason for the difference between the results for tube **A** and tube **B**.

.....

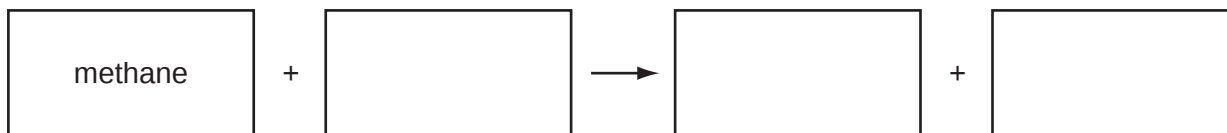
 [2]

- 8 Large amounts of chemical energy are stored in the world's reserves of fossil fuels such as natural gas and petroleum (crude oil).

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(a) Methane is found in natural gas.

(i) Complete the **word** chemical equation for the complete combustion of methane.



[3]

(ii) State the term used to describe chemical reactions that release heat.

..... [1]

(b) Petroleum is a mixture of a very large number of compounds.

Fig. 8.1 shows a diagram of the industrial process used to separate petroleum into mixtures that are more useful.

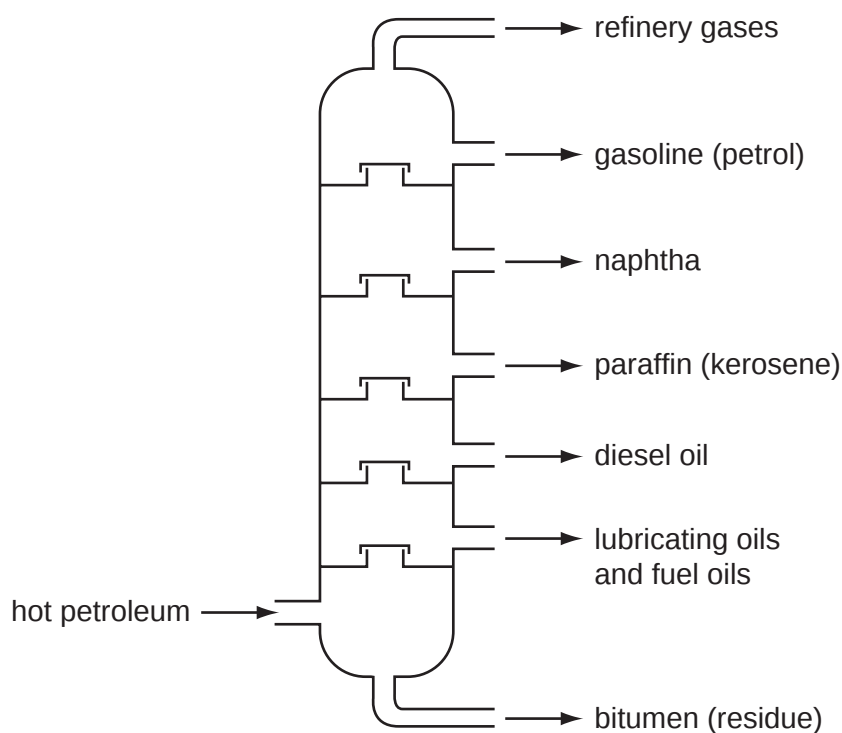
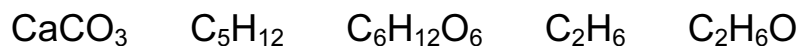


Fig. 8.1

(i) State the full name of the process shown in Fig. 8.1.

..... [1]

- (ii) The list below shows the chemical formulae of five compounds.



State and explain which of these formulae represent compounds that are found in petroleum.

formulae

explanation [2]

- (iii) State **one** use of refinery gas.

..... [1]

- (iv) Refinery gas contains the compound ethane.

Complete the diagram of the structure of one molecule of ethane which has been started below.



[2]

- 9 Fig. 9.1 shows a toy car of mass 0.5 kg travelling over a plastic surface.

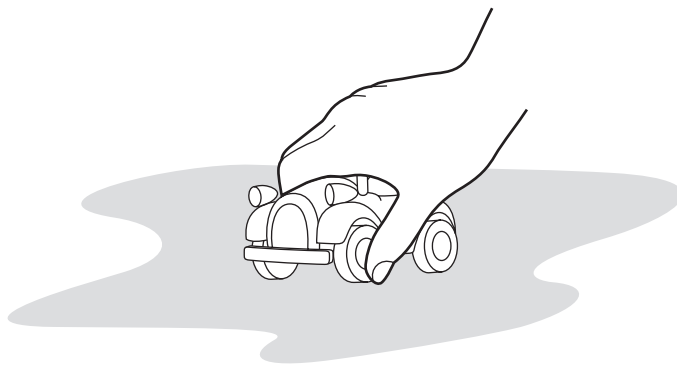


Fig. 9.1

- (a) While the car is moving the wheels are rubbing against the plastic surface. The car becomes electrostatically charged with a positive charge.

Explain how this happens.

.....

.....

..... [3]

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- (b) A speed – time graph for the car is shown in Fig. 9.2. It shows the motion of the car over a 25 second period.

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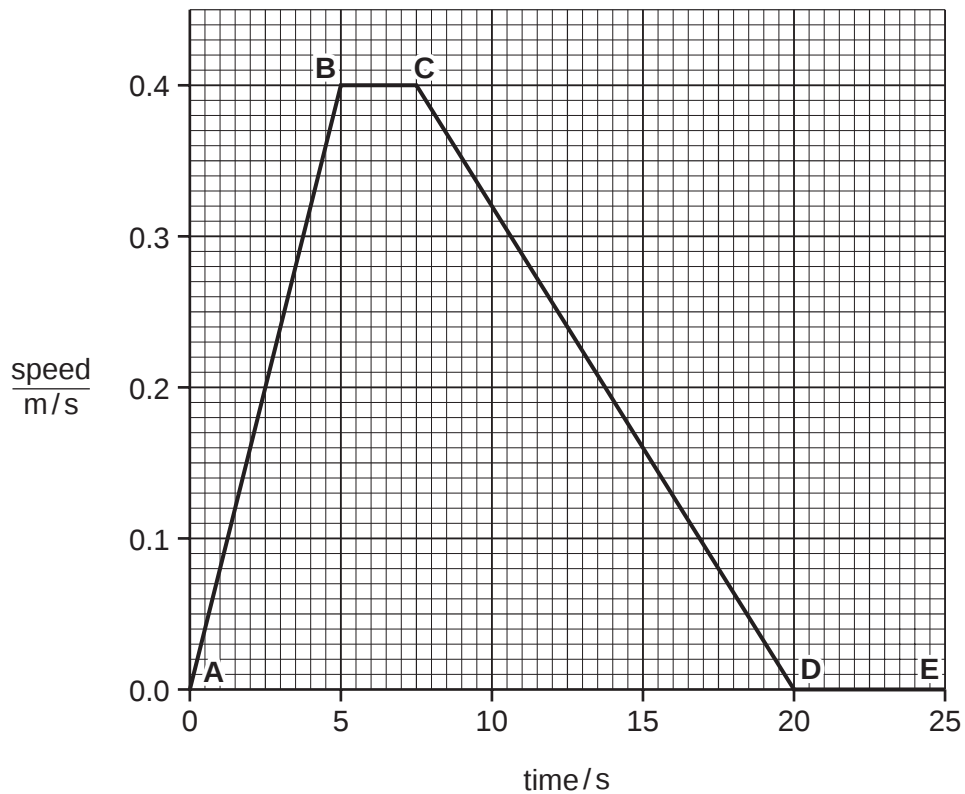


Fig. 9.2

- (i) Use the graph to find one time when the car is not moving. Write down this time.

..... [1]

- (ii) Determine **one** part of the graph when the car was travelling at constant speed and write down the value of this speed.

part of graph

speed [2]

DATA SHEET
The Periodic Table of the Elements

Group																	
I	II											III	IV	V	VI	VII	0
												1 H Hydrogen 1					
7 Li Lithium 3	9 Be Beryllium 4																
23 Na Sodium 11	24 Mg Magnesium 12																
39 K Potassium 19	40 Ca Calcium 20	45 Sc Scandium 21	48 Ti Titanium 22	51 V Vanadium 23	52 Cr Chromium 24	55 Mn Manganese 25	56 Fe Iron 26	59 Co Cobalt 27	59 Ni Nickel 28	64 Cu Copper 29	65 Zn Zinc 30	70 Ga Gallium 31	73 Ge Germanium 32	75 As Arsenic 33	79 Se Selenium 34	80 Br Bromine 35	84 Kr Krypton 36
85 Rb Rubidium 37	88 Sr Strontium 38	89 Y Yttrium 39	91 Zr Zirconium 40	93 Nb Niobium 41	96 Mo Molybdenum 42	101 Ru Ruthenium 44	101 Rh Rhodium 45	106 Pd Palladium 46	108 Ag Silver 47	112 Cd Cadmium 48	115 In Indium 49	119 Sn Tin 50	122 Sb Antimony 51	127 Te Tellurium 52	131 I Iodine 53	136 Xe Xenon 54	136 Rn Radon 86
133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium 75	190 Os Osmium 76	192 Ir Iridium 77	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84	210 At Astatine 85	222 Rn Radon 86
Fr Francium 87	88 Ra Radium 88	89 Ac Actinium 89															
58-71 Lanthanoid series																	
90-103 Actinoid series																	
175 Lu Lutetium 71																	
173 Yb Ytterbium 70																	
169 Tm Thulium 69																	
167 Er Erbium 68																	
165 Ho Holmium 67																	
162 Dy Dysprosium 66																	
159 Tb Terbium 65																	
157 Gd Gadolinium 64																	
152 Eu Europium 63																	
150 Sm Samarium 62																	
144 Nd Neodymium 60																	
141 Pr Praseodymium 59																	
140 Ce Cerium 58																	
238 U Uranium 92																	
238 Pa Protactinium 91																	
232 Th Thorium 90																	
232 Pu Plutonium 94																	
238 Am Americium 95																	
238 Cm Curium 96																	
238 Bk Berkelium 97																	
238 Cf Californium 98																	
238 Es Einsteinium 99																	
238 Fm Fermium 100																	
238 Md Mendelevium 101																	
238 No Nobelium 102																	
238 Lr Lawrencium 103																	

a

X

b

a = relative atomic mass
X = atomic symbol
b = proton (atomic) number

*58-71 Lanthanoid series
†90-103 Actinoid series

a	X	a = relative atomic mass
Key	X	X = atomic symbol
b		b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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