## MARK SCHEME for the October/November 2012 series

## 0620 CHEMISTRY

0620/51

Paper 5 (Practical), maximum raw mark 40

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2012 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.



	Page 2		Mark Scheme	Syllabus	Paper			
			IGCSE – October/November 2012	0620	51			
1	(e)	e) Table of results for Experiments						
		all initial temperature boxes completed correctly as instructed (1)						
		all final temperature boxes completed correctly not more than 20 °C below original (1)						
		all average temperatures completed correctly (1)						
		times co	ompleted in seconds (1) <b>ignore:</b> dps					
		<u>descend</u>	<u>ling</u> in order (comparable to supervisor) (1)		[5]			
	(f)	points pl	lotted correctly (4)					
		smooth I	line graph (1)		[5]			
	(g)	average	temperature 72 °C (1)					
		value fro	om graph (1)					
		extrapola	ation shown on grid (1)		[3]			
	(h)	as an inc	dicator/check presence of iodine owtte (1)		[1]			
	(i)	(i) expe	eriment 5/when temperature is 70 (1)		[1]			
		(ii) high	nest temperature (1)					
		parti	ticles have more energy/more collisions (1)		[2]			
	(j)	time long	ger/more/increase (1)					
	speed slower/decrease (1)		lower/decrease (1)		[2]			
	(k)	more <u>acc</u>	curate (1)		[1]			
2	(a)	pH 5–7 (	(1) ignore colours		[1]			
	(b)	(i) white	te (1) precipitate (1) dissolves owtte (1)		[3]			
		(ii) white	te (1) precipitate (1) dissolves owttte (1)		[3]			
	(c)	no reacti	ion/no change/no precipitate/no observation (1)		[1]			
	(d)	white (1)	) precipitate (1)		[2]			

Pag	je 3	Mark Scheme	Syllabus	Paper
		IGCSE – October/November 2012	0620	51
(e)	litmus tu	rns red (1) then bleached/white (1)		[2]
(f)	bubbles/	fizz etc. (1)		
	glowing	splint (1) glows brighter/relights (1)		[3]
(g)	zinc (1) s	sulfate (1)		[2]
(h)	oxygen (	1)		[1]
(i)	transitior	n metal present (1) catalyst (1)		
	mangane	ese/copper (1) oxide (1) max 2		[2]
				[Total: 40]