



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

CHEMISTRY

0620/11

Paper 1 Multiple Choice

May/June 2013

45 Minutes

Additional Materials: Multiple Choice Answer Sheet
 Soft clean eraser
 Soft pencil (type B or HB is recommended)



READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.

This document consists of **15** printed pages and **1** blank page.



1 The diagram shows a cup of tea.



Which row describes the water particles in the air above the cup compared with the water particles in the cup?

	moving faster	closer together
A	✓	✓
B	✓	x
C	x	✓
D	x	x

2 Crystals of sodium chloride were prepared by the following method.

- 1 25.0 cm³ of dilute hydrochloric acid was accurately measured into a conical flask.
- 2 Aqueous sodium hydroxide was added until the solution was neutral. The volume of sodium hydroxide added was measured.
- 3 The solution was evaporated and the crystals washed with approximately 15 cm³ of water.

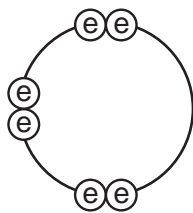
Which row shows the pieces of apparatus used to measure the 25.0 cm³ of hydrochloric acid, the volume of aqueous sodium hydroxide and the 15 cm³ of water?


	25.0 cm ³ of hydrochloric acid accurately	the volume of aqueous sodium hydroxide added	15 cm ³ of water approximately
A	burette	pipette	measuring cylinder
B	measuring cylinder	burette	pipette
C	pipette	burette	measuring cylinder
D	pipette	measuring cylinder	burette

6 For which substance is the type of bonding **not** correct?

	substance	type of bonding		
		ionic	covalent	metallic
A	chlorine		✓	
B	potassium bromide	✓		
C	sodium			✓
D	sodium chloride		✓	

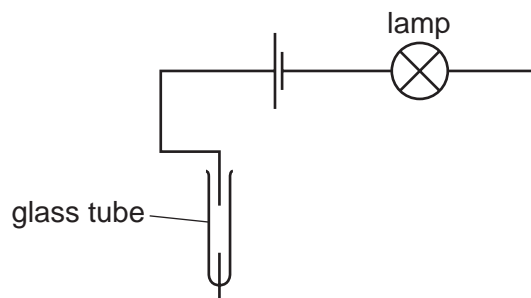
7 Element X has six electrons in its outer shell.



key
 = electron

How could the element react?

- A** by gaining two electrons to form a positive ion
 - B** by losing six electrons to form a negative ion
 - C** by sharing two electrons with two electrons from another element to form two covalent bonds
 - D** by sharing two electrons with two electrons from another element to form four covalent bonds
- 8 The diagram shows an incomplete circuit.



Which substance causes the lamp to light when added to the glass tube?

- A** aqueous sodium chloride
- B** aqueous sugar
- C** solid sodium chloride
- D** solid sugar

9 A compound with the formula XF_2 has a relative formula mass of 78.

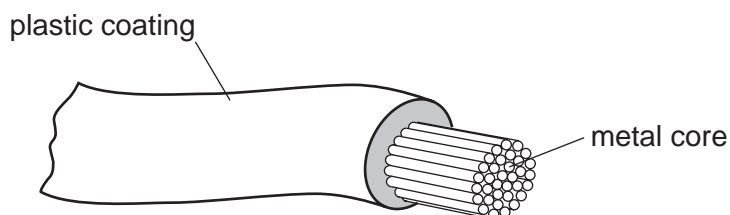
What is element X?

- A argon
- B calcium
- C neon
- D zirconium

10 What is the balanced chemical equation for the reaction between calcium and water?

- A $\text{Ca} + \text{H}_2\text{O} \rightarrow \text{CaOH} + \text{H}_2$
- B $\text{Ca} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$
- C $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{CaOH} + \text{H}_2$
- D $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$

11 The diagram shows an electrical cable.

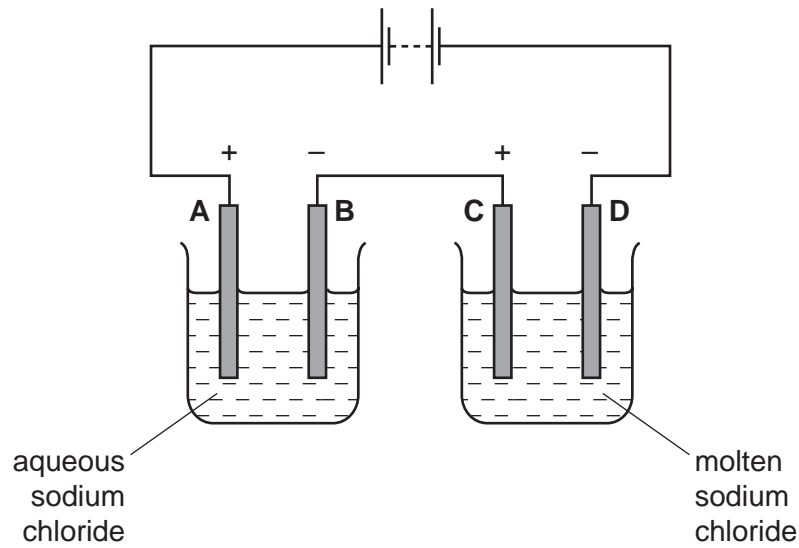


Which statement about the substances used is correct?

- A The coating is plastic because it conducts electricity well.
- B The core is copper because it conducts electricity well.
- C The core is copper because it is cheap and strong.
- D The core is iron because it is cheap and strong.

12 The diagram shows an electrolysis circuit.

At which electrode is hydrogen formed?



13 Some white anhydrous copper(II) sulfate powder is put into a beaker of water and stirred.

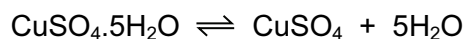
What would show that the process was exothermic?

- A A blue solution is formed.
- B The beaker feels cooler.
- C The beaker feels warmer.
- D The powder dissolves in the water.

14 Which substance does **not** require oxygen in order to produce energy?

- A coal
- B hydrogen
- C natural gas
- D ^{235}U

- 15 The equation shows the formation of anhydrous copper(II) sulfate from hydrated copper(II) sulfate.



Statements 1, 2 and 3 refer to this reaction.

- 1 Hydrated copper(II) sulfate is reduced to anhydrous copper(II) sulfate.
- 2 The (II) in the name copper(II) sulfate refers to the oxidation state of the metal.
- 3 The reaction is reversible.

Which statements are correct?

- A** 1 only **B** 1 and 2 **C** 2 and 3 **D** 3 only

- 16 Calcium carbonate reacts with hydrochloric acid to form carbon dioxide.

Which changes would slow this reaction down?

- 1 decreasing the concentration of hydrochloric acid
- 2 decreasing the particle size of calcium carbonate
- 3 decreasing the temperature

- A** 1 and 2 only **B** 1 and 3 only **C** 2 and 3 only **D** 1, 2 and 3

- 17 The equations represent redox reactions.

In which equation is the underlined substance acting as a reducing agent?

- A** 3CO + Fe₂O₃ → 2Fe + 3CO₂
- B** CO₂ + C → 2CO
- C** CuO + H₂ → Cu + H₂O
- D** CaO + H₂O → Ca(OH)₂

- 18 Ant stings hurt because of the methanoic acid produced by the ant.

Which substance could, **most safely**, be used to neutralise the acid?

	substance	pH
A	baking soda	8
B	car battery acid	1
C	lemon juice	3
D	oven cleaner	14

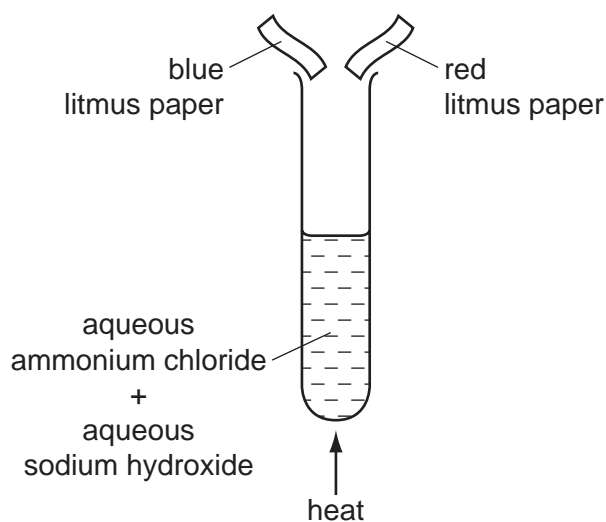
19 The diagram shows one period of the Periodic Table.

Li	Be	B	C	N	O	F	Ne
----	----	---	---	---	---	---	----

Which two elements form acidic oxides?

- A carbon and lithium
- B carbon and neon
- C carbon and nitrogen
- D nitrogen and neon

20 The diagram shows an experiment.



What happens to the pieces of litmus paper?

	blue litmus paper	red litmus paper
A	changes colour	changes colour
B	changes colour	no colour change
C	no colour change	changes colour
D	no colour change	no colour change

- 21 Two indicators, bromophenol blue and Congo red, show the following colours in acidic solutions and in alkaline solutions.

indicator	acid	alkali
bromophenol blue	yellow	blue
Congo red	violet	red

A few drops of each indicator are added to separate samples of a solution of pH 2.

What are the colours of the indicators in this solution?

	in a solution of pH 2	
	bromophenol blue is	Congo red is
A	blue	red
B	blue	violet
C	yellow	red
D	yellow	violet

- 22 Which property of elements increases across a period of the Periodic Table?

- A** metallic character
- B** number of electron shells
- C** number of outer shell electrons
- D** tendency to form positive ions

- 23 Which element is a transition metal?

	colour of chloride	melting point of element / °C
A	white	113
B	white	1495
C	yellow	113
D	yellow	1495

24 Fluorine is at the top of Group VII in the Periodic Table.

Which row shows the properties of fluorine?

	colour	state at room temperature	reaction with aqueous potassium iodide
A	brown	gas	no reaction
B	brown	liquid	iodine displaced
C	yellow	gas	iodine displaced
D	yellow	liquid	no reaction

25 Group I metals are also known as the Alkali Metals.

Which statement about the metals in Group I is **not** correct?

- A** In their reactions they lose electrons.
- B** Their atoms all have one electron in their outer shell.
- C** They form +1 ions in their reactions with non-metals.
- D** They form covalent compounds by sharing electrons.

26 Which element is a metal?

	charge on element ion	electrical conductivity
A	negative	low
B	positive	high
C	negative	high
D	positive	low

27 Which property makes aluminium ideal for making food containers?

- A** conducts electricity
- B** conducts heat
- C** mechanical strength
- D** resistance to corrosion

28 Which substance is **not** involved in the extraction of iron from hematite?

- A carbon
- B carbon monoxide
- C calcium carbonate
- D nitrogen

29 Pure metals conduct electricity and can be hammered into different shapes.

Why are metals sometimes used as alloys?

- A Alloys are cheaper than the metals they are made from.
- B Alloys are easier to hammer into different shapes.
- C Alloys are harder and keep their shape better.
- D Alloys conduct electricity better.

30 Below are some metals in decreasing order of reactivity.

magnesium

zinc

iron

copper

Titanium reacts with acid and cannot be extracted from its ore by heating with carbon.

Where should titanium be placed in this list?

- A below copper
- B between iron and copper
- C between magnesium and zinc
- D between zinc and iron

31 Water has been contaminated with sea-water.

Which substances can be removed by chlorination and filtration?

- A bacteria, sand and sodium chloride
- B bacteria and sand only
- C bacteria and sodium chloride only
- D sand and sodium chloride only

32 Iron rusts when it reacts with1.....

Rusting can be prevented by covering the iron with a more reactive metal, such as2.....

Which words correctly complete gaps 1 and 2?

	1	2
A	oxygen	copper
B	oxygen	magnesium
C	oxygen and water	copper
D	oxygen and water	magnesium

33 Nitrogen, phosphorus and potassium are essential elements for plant growth.

Which mixture provides all three essential elements?

	mixture	formula
A	ammonium phosphate + potassium chloride	$(\text{NH}_4)_3\text{PO}_4$ + KCl
B	ammonium phosphate + ammonium nitrate	$(\text{NH}_4)_3\text{PO}_4$ + NH_4NO_3
C	ammonium phosphate + ammonium chloride	$(\text{NH}_4)_3\text{PO}_4$ + NH_4Cl
D	ammonium nitrate + potassium chloride	NH_4NO_3 + KCl

34 Which information about carbon dioxide and methane is correct?

		carbon dioxide	methane
A	formed when vegetation decomposes	✓	✗
B	greenhouse gas	✓	✓
C	present in unpolluted air	✗	✗
D	produced during respiration	✗	✓

key
✓ = true
✗ = false

35 The list shows four methods that were suggested for the formation of carbon dioxide.

- 1 action of an alkali on a carbonate
- 2 action of heat on a carbonate
- 3 complete combustion of methane
- 4 reaction of a carbonate with oxygen

Which methods would result in the production of carbon dioxide?

- A** 1 and 2 **B** 1 and 4 **C** 2 and 3 **D** 3 and 4

36 Organic compounds may have names ending in -ane, -ene, -ol or -oic acid.

How many of these endings indicate the compounds contain double bonds in their molecules?

- A** 1 **B** 2 **C** 3 **D** 4

37 The table shows the boiling points of four members of the homologous series of alcohols.

compound		boiling point /°C
name	formula	
methanol	CH ₃ OH	65
ethanol	C ₂ H ₅ OH	78
propanol	C ₃ H ₇ OH	X
butanol	C ₄ H ₉ OH	117

What is the value of X?

- A** 55°C **B** 82°C **C** 98°C **D** 115°C

- 38 The table shows some fractions that are obtained from petroleum by fractional distillation, together with some of their uses.

fraction	use
refinery gas	cooking
gasoline	fuel for cars
1	making chemicals
2	jet fuel
3	fuel for ships
bitumen	making roads

Which row correctly identifies fractions 1, 2 and 3?

	1	2	3
A	diesel oil	fuel oil	lubricating fraction
B	fuel oil	diesel oil	kerosene
C	kerosene	naphtha	diesel oil
D	naphtha	kerosene	fuel oil

- 39 Which columns describe the hydrocarbons ethane and ethene?

	1	2	3	4
state at room temperature	gas	gas	liquid	liquid
reaction with oxygen	burns	burns	burns	burns
reaction with aqueous bromine	no reaction	decolourises bromine	no reaction	decolourises bromine

- A** 1 (ethane) and 2 (ethene)
B 1 (ethane) and 4 (ethene)
C 2 (ethene) and 3 (ethane)
D 3 (ethane) and 4 (ethene)
- 40 Which of the statements about ethanol are correct?
- Ethanol can be formed by an addition reaction.
 - Ethanol can be formed by fermentation.
 - When ethanol burns in air, it forms carbon dioxide and water.
- A** 1, 2 and 3 **B** 1 and 2 **C** 1 and 3 **D** 2 and 3

DATA SHEET
The Periodic Table of the Elements

		Group																																																																																		
		I	II	III	IV	V	VI	VII	0																																																																											
7	Li Lithium 3	9	Be Beryllium 4	1 H Hydrogen 1							4	He Helium 2																																																																								
23	Na Sodium 11	24	Mg Magnesium 12	55	Mn Manganese 25	56	Fe Iron 26	57	Co Cobalt 27	58	Ni Nickel 28	59	Cu Copper 29	60	Zn Zinc 30	65	Ga Gallium 31	70	Ga Gallium 31	73	Ge Germanium 32	74	As Arsenic 33	75	Se Selenium 34	76	Br Bromine 35	77	Kr Krypton 36	84																																																						
39	K Potassium 19	40	Ca Calcium 20	51	V Vanadium 23	48	Ti Titanium 22	45	Sc Scandium 21	46	Ti Titanium 22	47	V Vanadium 23	48	Cr Chromium 24	49	Mn Manganese 25	50	Fe Iron 26	51	Co Cobalt 27	52	Ni Nickel 28	53	Cu Copper 29	54	Zn Zinc 30	115	In Indium 49	116	Sn Tin 50	117	Sb Antimony 51	118	Te Tellurium 52	119	Sn Tin 50	120	Pb Lead 82	121	Tl Thallium 81	122	Pb Lead 82	123	Bi Bismuth 83	124	Po Polonium 84	125	At Astatine 85	126	Rn Radon 86	86																																
85	Rb Rubidium 37	86	Sr Strontium 38	91	Zr Zirconium 40	92	Nb Niobium 41	93	Mo Molybdenum 42	94	Tc Technetium 43	95	Ru Ruthenium 44	96	Rh Rhodium 45	97	Pd Palladium 46	98	Ag Silver 47	99	Cd Cadmium 48	100	In Indium 49	101	Sn Tin 50	102	Sb Antimony 51	103	Te Tellurium 52	104	I Iodine 53	105	Xe Xenon 54	133	Cs Caesium 55	134	Ba Barium 56	135	La Lanthanum 57	136	Ce Cerium 58	137	Pr Praseodymium 59	138	Nd Neodymium 60	139	Pm Promethium 61	140	Ce Cerium 58	141	Pr Praseodymium 59	142	Nd Neodymium 60	143	Pm Promethium 61	144	Sm Samarium 62	145	Eu Europium 63	146	Gd Gadolinium 64	147	Tb Terbium 65	148	Dy Dysprosium 66	149	Ho Holmium 67	150	Er Erbium 68	151	Tm Thulium 69	152	Yb Ytterbium 70	153	Lu Lutetium 71	175	Lu Lutetium 71	176	Yb Ytterbium 70	177	No Nobelium 102	178	Lr Lawrencium 103	103
133	Cs Caesium 55	134	Ba Barium 56	181	Ta Tantalum 73	182	Hf Hafnium 72	183	La Lanthanum 57	184	W Tungsten 74	185	Re Rhenium 75	186	Os Osmium 76	187	Ir Iridium 77	188	Pt Platinum 78	189	Au Gold 79	190	Hg Mercury 80	201	Tl Thallium 81	202	Pb Lead 82	203	Bi Bismuth 83	204	Po Polonium 84	205	At Astatine 85	206	Rn Radon 86	226	Ra Radium 88	227	Ac Actinium 89	228	Th Thorium 90	229	Pa Protactinium 91	230	U Uranium 92	231	Np Neptunium 93	232	Pu Plutonium 94	233	Am Americium 95	234	Cm Curium 96	235	Bk Berkelium 97	236	Cf Californium 98	237	Es Einsteinium 99	238	Fm Fermium 100	239	Md Mendelevium 101	240	No Nobelium 102	241	Lr Lawrencium 103																	

*58-71 Lanthanoid series
†90-103 Actinoid series

a
X
b

a = relative atomic mass
X = atomic symbol
b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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