## MARK SCHEME for the May/June 2013 series

## 0620 CHEMISTRY

0620/62

Paper 6 (Alternative to Practical), maximum raw mark 60

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the May/June 2013 series for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level components and some Ordinary Level components.



	Page 2	Mark Scheme	Syllabus	Paper
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1	(a) pestle a	nd / or mortar (1) filter / funnel (1)		[2]
	<b>(b) (i)</b> labe	elled arrow at liquid in mortar (1)		
	(ii) labe	lled arrow at liquid in either tube or liquid in funnel o	or any combination	n (1) [2]
	( <b>c) (i)</b> top	line labelled (1)		[1]
	(ii) thre	e (1)		[1]
2	<b>(a)</b> black (1)	)		[1]
	<b>(b) (i)</b> cop	per / Cu (1)		
	(ii) wate	er / H <sub>2</sub> O (1) <b>accept:</b> steam		[2]
	(c) boiling p	oint / freezing point (1)		
	100 °C / <b>note:</b> do	0 °C (1) o not accept a chemical test		[2]

	Page 3			Mark Scheme	Syllabus	Paper
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3	(a)	table of results for Experiment 1				
		final and initial volumes and difference completed correctly 26.00, 0.0 and 26.0 (1)				
		to 1	deci	mal place (1) <b>accept:</b> volumes to 2 d.p. (e.g. 26.00	)	[2]
	(b)	tabl	e of r	results for Experiment 2		
		final and initial volumes and difference completed correctly 19.0 and 32.0 (1) 13.0 (1) [i <b>ignore:</b> decimal places, <b>accept:</b> 19, 32,13, <b>allow:</b> ecf on final and initial volumes				
	(c)	(i)		urless <b>not:</b> clear to purple / pink (1) ept: colour change either way round		[1]
		(ii) not an acid / alkali reaction or potassium manganate is coloured or pink / acts as indicator / there is already a colour change / owtte (1)				ink / acts as an [1]
	(d)	(i)	expe	eriment 1 (1) allow: ecf on (a) and (b)		[1]
		<ul> <li>(ii) experiment 1 is twice the volume of experiment 2 / experiment 2 is half the volume experiment 1 (1) note: must be a quantitative comparison, do not allow quotes of figure from table allow: ecf (e.g. 13 times as much as experiment 2)</li> </ul>				
		(iii)	solu	tion B / experiment 1 more concentrated / stronger (	(1) or converse	
				ble / twice (1) bre: reference to reactivity		[2]
	(e)	half	value	e from table result for experiment 2 (6.5) (1) allow:	ecf	
		cm <sup>3</sup>	<sup>3</sup> (1)			
	half volume of <b>C</b> used (1)				[3]	
	(f)	(f) oxidation (1) reduction (1)				
		or: electrons are lost (1) gained (1) transferred (2)[2]accept: oxidation numbers increase (1) decrease (1)accept: hydrogen / $H_2$ / H lost (1) gained (1)accept: oxygen / $O_2$ / O gained (1) lost (1)				
	(g)	advantage easy to use / quick / convenient (1) ignore: large volumes				
		disa	advan	ntage not accurate / owtte (1)		[2]

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4	(a)	colourles	es (1) <b>ignore:</b> clear, <b>not:</b> white		[1]
	(b)	white (1)	precipitate (1)		
		dissolves	s / clears (1)		[3]
	(c)	white pre	ecipitate (1) insoluble / does not dissolve (1)		[2]
	(d)	no chang	ge / colourless solution / no reaction (1)		[1]
	(e)	white (1)	precipitate (1)		[2]
	(g)	carbon d	lioxide / CO <sub>2</sub> (1)		[1]
	(h)	calcium / <b>note:</b> Ca	$Ca^{2+}$ (1) <b>accept:</b> any Group 2 metals carbonate / $Ca^{2+}$ (1) <b>accept:</b> any Group 2 metals carbonate / $Ca^{2+}$	CO <sub>3</sub> <sup>2<sup>-</sup></sup> (1)	[2]
5	(a)	thermom	eter diagrams completed correctly (3) –1 each inco	rrect	
		23, 29, 3	5, 41, 39, 35, 31 <b>ignore:</b> decimal places		[3]
	(b)	points pl	otted correctly (3), -1 each incorrect		
			secting straight lines (1) nes extending beyond intercept but must be just two	lines and no curve	[4] s
	(c)	16 (cm <sup>3</sup> )	±0.5 (1) any indication (1)		[2]
	(d)	23(°C)(	1)		[1]
	(e)	good ins	ulator or reference to minimising heat losses (1)		[1]
	(f)	reaction	produced heat or energy (1) <b>accept:</b> reaction is exc	othermic	
			finished / reactant(s) used up / KOH used up / neutr I used up/ neutralised	alised (1)	[2]
	(g)	exothern	nic (1)		[1]

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6 **note:** all methods can gain the first three marks but only methods that would give usable results can gain the last three marks

known / same mass / amount of metal (1)

known / same volume / amount of acid (1)

test both **A** and **B** (1)

a method of collecting results (1)

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time or run side by side (1)
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comparison of results (1)

max 6

[6]