

MARK SCHEME for the May/June 2010 question paper
for the guidance of teachers

0620 CHEMISTRY

0620/31

Paper 31 (Extended Theory), maximum raw mark 80

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes must be read in conjunction with the question papers and the report on the examination.

- CIE will not enter into discussions or correspondence in connection with these mark schemes.

CIE is publishing the mark schemes for the May/June 2010 question papers for most IGCSE, GCE Advanced Level and Advanced Subsidiary Level syllabuses and some Ordinary Level syllabuses.



Page 2	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – May/June 2010	0620	31

- 1 (i) sulfur [1]
- (ii) iodine [1]
- (iii) copper **ignore** (II) [1]
- (iv) calcium [1]
- (v) helium [1]
not name of a compound
accept correct symbols
- 2 (i) chloromethane [1]
cond biggest molecular mass / biggest mass of one mole / its molecules
move slowest / heaviest molecule / highest density [1]
accept atomic mass if correct numerical value given
ignore it is the heaviest (gas) / biggest molecule
accept particles or molecules
not atoms
- (ii) carbon dioxide / calcium carbonate [1]
not methane [1]
water [1]
sodium chloride / brine / seawater [1]
- (iii) chlorine [1]
not chlorine water
cond light / UV / heat / high temperature if numerical value given about
200°C / lead tetraethyl [1]
not warm
- (iv) oxygen and nitrogen (in air) [1]
not from fuel, negates mark 1 [1]
(react) at high temperatures / lightning / in engine
not combustion or exhaust, negates mark 2
- (v) $2\text{O}_3 \rightarrow 3\text{O}_2$ [2]
not balanced = [1]
- 3 (a) (i) bubbles / effervescence / hydrogen / gas pushes up / lifts metal [1]
- (ii) does not react with acid / zinc and iron react with acid [1]
not just unreactive
- (b) (i) with copper / first experiment [1]
- (ii) copper acts as a catalyst [1]
- (c) (i) smaller gradient [1]
not rate is slower
- (ii) same final volume of hydrogen / same level (on graph) [1]

Page 3	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – May/June 2010	0620	31

- (d) temperature / heat [1]
 increase temperature – reaction faster particles have more energy / particles move faster / particles collide more frequently / more particles have enough energy to react
not more excited
accept arguments for a decrease in temperature [1]
- powdered
 greater surface area
 greater collision rate / more particles exposed (to acid)
 any **two** [2]
not concentration / light / catalyst / pressure
- 4 (a) (i) ethanol [1]
 $\text{CH}_3\text{-CH}_2\text{-OH}$ [1]
- propanoic acid [1]
 $\text{CH}_3\text{-CH}_2\text{-COOH}$ [1]
 independent marking, no ecf
accept C_2H_5
not – HO
- (ii) type of compound – salt / sodium carboxylate / alkanoate [1]
not soap / sodium stearate etc
 use – soap / cleaning / detergent [1]
- (iii) terylene / PET / Dacron / diolen / mylar / crimplene [1]
- (b) (i) polyamide / amide / peptide / polypeptide [1]
- (ii) correct amide linkage NHCO then CONH [1]
cond to mark 1, 2 monomers (different shading in box) [1]
cond continuation (to **ONE** correct linkage) [1]
- OR** nylon 6
 only one linkage – NHCO [1]
cond only one monomer [1]
cond continuation (to correct linkage) [1]
- (iii) use locating agent [1]
 measure distance travelled by sample / travelled by solvent front [1]
cond this is $R_f = 0.5$ [1]
 for mark 3, either mark 1 or mark 2 must be awarded
- accept** run a chromatogram of glycine [1]
 compare with sample
 same position [1] max [2]

Page 4	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – May/June 2010	0620	31

- 5 (a) (i) macromolecular / giant covalent / giant atomic [1]
all atoms held in position / in tetrahedral structure / to four other carbon
atoms / all strong bonds [1]
- (ii) jewellery / drilling / cutting / engraving / cutting edges in scalpels [1]
mark first use offered
- (iii) layer structure / sheets [1]
molecules / ions in layers = [0]
layers can slide (over each other) [1]
- (iv) lubricant / pencils / electrodes [1]
mark first use offered
- (b) (i) 4e between carbon and oxygens [1]
2 non-bonding pairs on both oxygens [1]
cond correct coding – only scored if marks 1 and 2 awarded [1]
ignore O₂ in atom
- (ii) 4O around each Si [1]
2Si around each O [1]
must refer to diagram **not** valencies **or** electron distributions
- (iii) SiO₂ has higher mp or bp
SiO₂ is a solid, CO₂ is a gas (at rtp)
(when both are solids) then SiO₂ is harder
has higher density
SiO₂ insoluble, CO₂ soluble [2]
any **two**, comparison needed
- 6 (a) rates equal [1]
concentrations do not change / macroscopic properties remain constant [1]
accept amounts do not change
- (b) endothermic [1]
cond favoured by high temperatures [1]
- (c) (i) move to left [1]
cond bigger volume / more moles etc [1]
do not insist on “gas”
- (ii) less yellow solid / more brown liquid [1]
accept yellow to brown / less solid more liquid / goes brown

Page 5	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – May/June 2010	0620	31

- 7 (a) a transition element has more than one oxidation state or valency [1]
accept different oxidation states
- (b) by removing oxygen concentration of O₂ decreases [1]
prevents the back reaction / equilibrium shifts to right [1]
- (c) oxidation number reduced (from (+) 4 to 0) [1]
accept accepts electrons **or** accepts four electrons
if number given must be 4
- (d) low density / lightweight / light [1]
propellers / fittings on ships / inert anodes in electrolysis / hip replacements /
ship building / chemical plants / cathodic protection / diving equipment [1]
- (e) (i) percentage of oxygen = 31.6% [1]
- (ii) calculate the number of moles of atoms for each element
number of moles of Ti = 31.6/48 = 0.66
number of moles of O = 31.6/16 = 1.98 **accept** 2 [1]
both correct for one mark
- (iii) the simplest whole number ratio for moles of atoms:
Fe : Ti : O
1 : 1 : 3 [1]
- (iv) formula is FeTiO₃ **accept** TiFeO₃ [1]
must be whole numbers from (iii) or cancelled numbers from (iii)
mark **ecf** throughout

Page 6	Mark Scheme: Teachers' version	Syllabus	Paper
	IGCSE – May/June 2010	0620	31

- 8 (a) same general formula
 same chemical properties
 same functional group
 physical properties vary in predictable way
 common methods of preparation
 consecutive members differ by CH₂
 any **two** [2]
mark first two
ignore others unless it contradicts a point which has been awarded a mark
- (b) (i) $2\text{HCOOH} + \text{CaCO}_3 \rightarrow \text{Ca}(\text{HCOO})_2 + \text{CO}_2 + \text{H}_2\text{O}$ [2]
not balanced = [1]
- (ii) zinc + methanoic acid → zinc methanoate + hydrogen [2]
 [1] for each product
- (iii) protected by oxide layer [1]
- (c) butanoic acid [1]
 CH₃-CH₂-CH₂-COOH / C₄H₈O₂ / C₃H₇COOH / C₄H₇OOH [1]
 C₂H₄O [1]
 mark **ecf** to molecular formula