

**CAMBRIDGE INTERNATIONAL EXAMINATIONS**

**GCE Advanced Level**

## **MARK SCHEME for the May/June 2013 series**

### **9701 CHEMISTRY**

**9701/42**

Paper 4 (A2 Structured Questions), maximum raw mark 100

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

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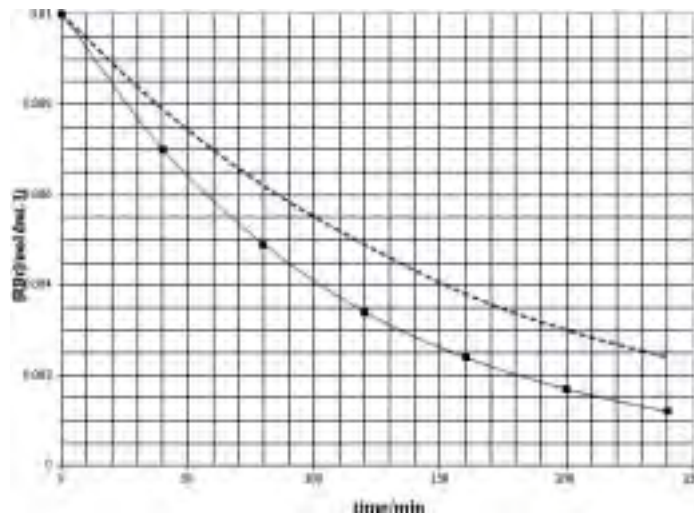
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(ii) nucleophilic substitution [1]

[2]

(b) (i)



plotting of all points (plotted to within  $\frac{1}{2}$  small square) [1]  
good line of best fit [1]

(ii)  $t_{1/2} = 118 \text{ min}$  or  $79 \text{ min} (\pm 5 \text{ min})$

or

construction lines for two half-lives **and** mention that half-life is constant

or

calculate the ratio of two rates at two different concentrations [1]

(iii) either ratio of initial rates (slopes)

or

ratio of  $t_{1/2}$

or

ratio of times for  $[\text{RBr}]$  to fall to the same level: all should be = 1.5 [1]

therefore reaction is first order w.r.t.  $[\text{OH}^-]$  [1]

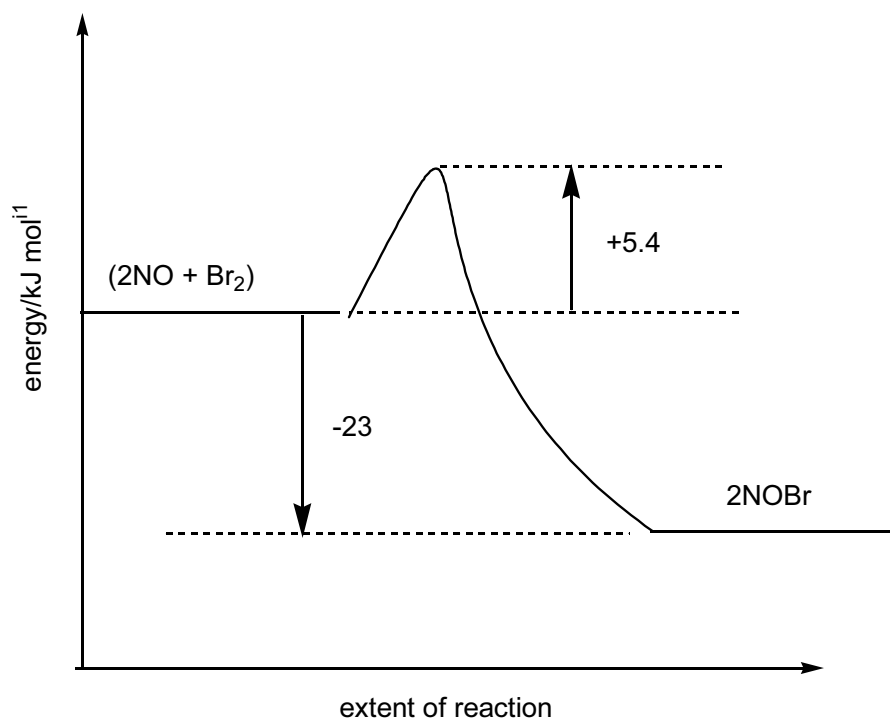
(iv)  $\text{rate} = k[\text{RBr}][\text{OH}^-]$  [1]

$$\text{initial rate} = 0.01 / 185 = 5.4 \times 10^{-5} \text{ (mol dm}^{-3} \text{ min}^{-1}\text{)} [1]$$

$$k = 5.4 \times 10^{-5} / (0.01 \times 0.1) = 0.054 \text{ (mol}^{-1} \text{ dm}^3 \text{ min}^{-1}\text{)} [1]$$

[8 max 7]

(c)

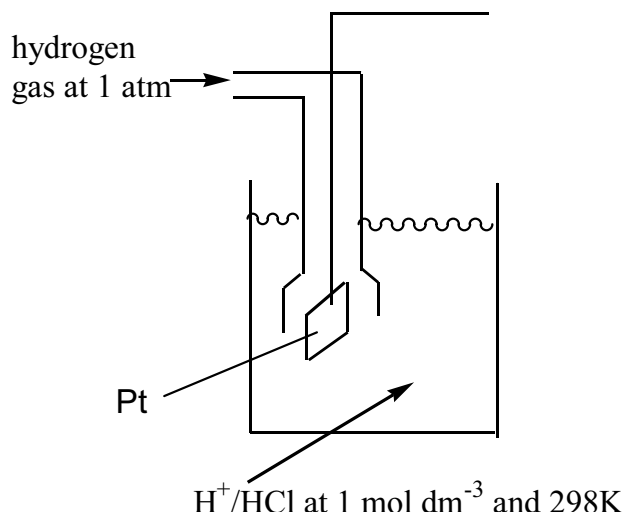


four marking points: one activation "hump"  
2NOBr (not just NOBr)  
 $\Delta H$  labelled correctly (arrow down, or double headed, or just a line)  
 $E_a$  labelled correctly (arrow up, or double headed, or just a line)  
 all four points [2]  
 three or two points [1]  
 [2]

[Total: 11]

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2 (a) (i)



$\text{H}_2(\text{g})$  going in (i.e. not being produced) [1]  
 platinum electrode in contact with solution, with  $\text{H}_2$  bubbling over it [1]  
 $\text{H}^+$  or  $\text{HCl}$  or  $\text{H}_2\text{SO}_4$  [1]  
 solution at  $1 \text{ mol dm}^{-3}$  (or  $0.5 \text{ M}$  if  $\text{H}_2\text{SO}_4$ ) and  $T=298 \text{ K}$ ,  $p=1 \text{ atm}$  [1]

(ii)  $E^\circ = 1.33 - (-0.41) = 1.74 \text{ V}$  [1]



(iii) Colour would change from orange [1]  
 to green [1]  
**[8]**

(b) there are two ways of calculating the ratio:  
 $\text{pK}_a = -\log_{10}(\text{K}_a) = -\log_{10}(1.79 \times 10^{-5}) = 4.747$  (4.75) or  $[\text{H}^+] = 10^{-5.5} = 3.16 \times 10^{-6}$  [1]

$$\log_{10}([\text{B}] / [\text{A}]) = \text{pH} - \text{pK}_a = 0.753$$
 (0.75) or  $[\text{salt}] / [\text{acid}] = \text{K}_a / [\text{H}^+]$  [1]

$$\therefore [\text{B}] / [\text{A}] = 10^{0.753} = 5.66$$

$$\text{or } = 1.79 \times 10^{-5} / 3.16 \times 10^{-6} = 5.66$$

$$(\text{or } [\text{A}] / [\text{B}] = 0.177) \quad [1]$$

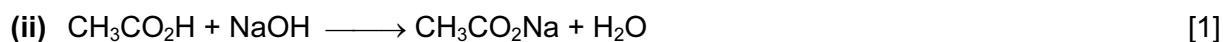
(correct ratio = [3] marks)

since  $B + A = 100$ ,  $\therefore (100-A) / A = 5.66 \Rightarrow$

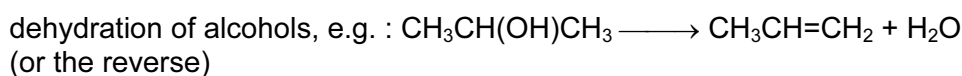
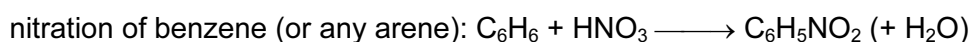
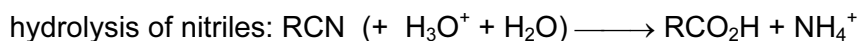
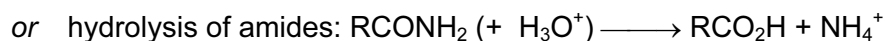
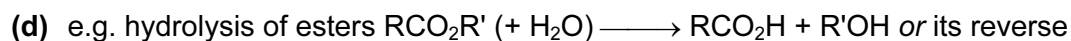
$$\frac{\text{vol of acid} = 15 \text{ cm}^3}{\text{vol of salt} = 85 \text{ cm}^3} \quad [1]$$

**[4]**

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[2]



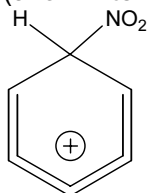
[3]

[Total: 17]



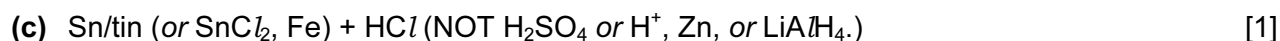
[3]

(b) (allow intermediate from methylbenzene)



[1]

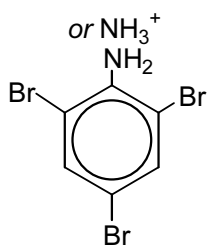
[1]



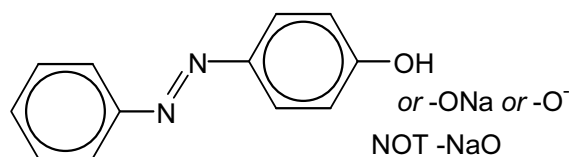
[1]

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(d) (i)



**A**



**B**

[1] + [1]

(ii)  $\text{NaNO}_2 + \text{HCl}$  or  $\text{H}_2\text{SO}_4$  or  $\text{H}^+$  or  $\text{HNO}_2$

[1]

$T \leq 10^\circ\text{C}$

[1]

**[4 max 3]**

(e) (i) amide

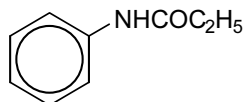
[1]

(ii)  $M_r = 108 + 11 + 14 + 16 = 149$

$\%N = (14 \times 100) / 149 = 9.4\%$

[1]

(iii)



[1]

**[3]**

**[Total: 11]**

4. (a) (i) Many electrons of similar energy in a valence-shell orbital

or

successive ionisation energies rise steadily (no big jumps)

or

ability to form bonds with ligands can stabilise very low or very high oxidation states

or

4s + 3d orbitals/shells/energy levels have similar / same energies

[1]

(ii)  $\text{VO}_2^+$ : +5

$\text{CrF}_6^{2-}$ : +4

$\text{MnO}_4^{2-}$ : +6

[3 × 1]

**[4]**

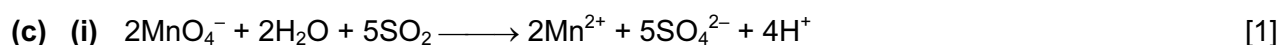
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(b)

- (colour due to) absorption of light/photons/frequencies/wavelengths  
or  
colour seen is complement of colour absorbed.
- d-orbitals/d-subshell split (by ligand field)
- (when photon is absorbed), electron is promoted or moves (from lower) to higher (d-)orbital
- energy difference/gap or  $\Delta E$  or splitting corresponds to photon/frequency/wavelength in visible region
- in s-block elements the energy gap is too large (to be able to absorb visible light)

[any four 4 × 1]

**[4]**



(ii) solution will go from purple [1]

to colourless [1]

**[3]**

(d) (pale) blue solution [1]

gives a (pale) blue ppt. [1]

which re-dissolves, or forms a solution, which is dark/deep blue or purple [1]

**[3]**

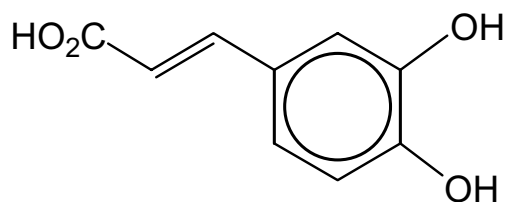
**[Total: 14]**





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(vi)



skeletal or structural [1]  
**[9 max 8]**

(c)  $M_r(\mathbf{E}) = 180$ , so  $0.1 \text{ g} = 1/1800$  ( $5.56 \times 10^{-4}$ ) mol [1]

3 mol NaOH react with 1 mol of **E**, so  $n(\text{NaOH}) = 3/1800 = 1/600 \text{ mol} = 1.67 \times 10^{-3} \text{ mol}$  [1]

volume of 0.1M NaOH =  $1000/(600 \times 0.1) = 16.7 \text{ cm}^3$  [1]  
**[3]**

**[Total: 17]**

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6 (a)

substance	protein synthesis	formation of DNA
cysteine	✓	
cytosine		✓
glutamine	✓	
guanine		✓

[3]

[3]

(b) (i) Hydrogen bonding [1]

Between bases *or* between A, T, C and G (all four needed) [1]

(ii) Bonds are (relatively) weak *or* easily broken [1]

This enables strands to separate *or* DNA to unzip/unwind/unravel. [1]

[4]

(c) changes / mutations in DNA

- by the addition / insertion / deletion / substitution / replacement of a base
- adds / deletes / replaces an amino acid *or* changes the amino acid sequence
- this causes a loss of function *or* changes the shape / tertiary structure of the protein

any three points [3]

[3]

[Total: 10]

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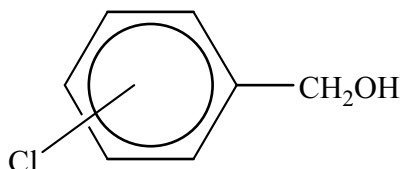
7. (a) (i)  $\frac{43.3}{3.35} = \frac{100}{1.1 \times n}$

$$n = \frac{100 \times 3.35}{43.3 \times 1.1} = 7.03 = 7 \text{ (calculation must be shown)} \quad [1]$$

(ii) The M and M+2 peaks are in the ratio 3 : 1 hence the halogen is chlorine/Cl [1]

(iii) L contains 7 hydrogen atoms *or* there are 3 types/environments of proton/H [1]

(iv) The multiplet with 4 hydrogens *or* peaks at  $\delta$  7.3 suggests a benzene ring  
 The singlet with 2 hydrogens *or* peak at  $\delta$  4.7 suggests a  $-\text{CH}_2-$  group  
 The singlet with 1 hydrogen *or* peak at  $\delta$  2.3 suggests an  $-\text{OH}$  group  
*or* reaction with Na suggests an OH group  
 OH must be an alcohol, not a phenol (due to its  $\delta$  value)  
 Since L also contains 7 carbon atoms and chlorine, this accounts for 126 of the 142 mass, the remaining atom must be oxygen  
 Thus L is



(allow the 2-, 3- or 4- isomer)

[6]  
**[9 max 7]**

(b) (i) we expect propene to have a  $\text{CH}_3$  peak *or* a peak at  $m/e$  15  
*or* cyclopropane would have fewer peaks [1]

(ii) cyclopropane would have 1 peak (ignore splitting)  
 propene would have 2 (*or* 3, *or* 4) peaks (ignore splitting)  
*or* propene would have peaks in the  $\delta$  4.5-6.0 (alkene) region  
 no splitting of cyclopropane peak  
 (any two points) [2]

[3]  
**[Total: 10]**

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- 8 (a) (i)  $\text{CH}_2 = \text{CH}-\text{CO}_2\text{H}$  or  $\text{CH}_2 = \text{CH}-\text{CO}_2\text{R}$  or  $\text{CH}_2 = \text{CH}-\text{COCl}$  [2]
- (ii) addition (polymerisation) [1]
- (iii)  $\text{C}(\text{CH}_2\text{OH})_4$  [1]
- (iv) water [1]
- [5]
- (b) (water is bonded to the polymer by) hydrogen bonding [1]  
hydrogen bonds are weak or easily broken [1]  
[2]
- (c) (i) cross-linking causes no reduction in the number of –OH groups  
or cross-linking molecules also have –OH groups [1]
- (ii) property e.g. becomes harder / more rigid / less flexible / stronger / higher melting  
point. [1]  
because the chains are more strongly / tightly held [1]  
[3]
- [Total: 10]